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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

OCHKIN, A.V.

5(0) SOV/153-2-2-31/31

AUTHORS: Mashovets, V. P., Pomomareva, A. M.

TITLE: Chronicle. All-Union Competition for the Best Students-paper Concerning Chemistry and Chemical Technology for the Scholastic Year 1957-1958 (Khronika. Vsesoyuznyy konkurs na luchshuyu studencheskuyu rabotu po khimii i khimicheskoy tekhnologii za 1957-1958 uchebnyy god)

PERIODICAL: Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya tekhnologiya. 1959, Vol 2, Nr 2, pp 303-304 (USSR)

ABSTRACT: The Ministerstvo vysshego obrazovaniya SSSR (Ministry for University-education of the USSR) carried out the competition mentioned in the title, within the framework of the Studencheskiye nauchnyye obshchestva (Scientific Student Societies) covering 37 subjects of science, technology, arts, and culture. The Leningradskiy tekhnologicheskiy institut imeni Lensoveta (Leningrad Technological Institute Imeni Lensovet) was entrusted with the subject "Chemistry and Chemical Technology". A commission was formed consisting of Professor V. B. Aleshkovskiy, V. P. Mashovets (Chairman), I. P. Mukhlenov, A. A. Petrov, B. A. Peray-Koshits, Docent P. A. Yablonskiy, and Candidate of Chemical Sciences A. M. Pomomareva (Secretary). The following persons acted as critics: The Professors A. P. Alabyshev, A. M. Ginetling, I. S. Ioffe, N. I. Khyaginichev, L. Ya. Kremnev, A. N. Kusov, A. N. Malkov, I. N. Nazlenitskiy, K. P. Nishchenko, Th. V. Moraschevskiy, with the collaborators, N. N. Nogaits, Th. K. Novodranov, V. V. Perekalin, A. L. Rotinyan, A. V. Shtalkin, A. V. Stepanskin, and T. A. Favorskaya with collaborators, A. N. Khaletskiy; Docents: A. Ye. Akin, L. N. Batuner, M. I. Gil'dengershel', O. P. Ginzburg, I. A. D'yakonov, S. G. Zhaveronok, S. N. Zhilov, Ye. S. Roskin, P. N. Sokolov, N. P. Starostenko, N. N. Sychev, A. P. Tsuchchenko; Chief scientific researcher B. P. Ioffe; Candidates of Sciences: G. A. Bel'chenko, N. K. Ryazayeva, O. N. Sethina, B. P. Yur'yev; Engineers: Kostyreva, Sazyushova, and Yarmolinskiy. The paper "Synthesis and Self-oxidation of the p-Bi-secondary Butyl-benzene" by V. S. Zavgorodniy, Fifth-year student of the Voronezhskiy gosudarstvennyy universitet (Voronezh State University) was awarded a medal for being the best. The second candidate for the medal is the Fifth-year-student of the Kievskiy gosudarstvennyy universitet (Kiyev State University) K. F. Lyashev. He submitted the paper "Kinetics of the Non-stationary Catalytic Decomposition-process of Hydrogen-peroxide on Platinum". The third medal was awarded to the Fourth-year-students of the Ivanovskiy khimiko-tekhnologicheskiy institut (Ivanovo Chemical-technological Institute): D. V. Nekova, A. I. Satsikova, T. T. Simagina, and E. N. Satyagina for the paper: "Method of Continuous Regeneration of Zinc-chloride From Waste Water of the Kineshma Fibre Factory". Besides these three papers, the commission selected further 6 papers which deserve publication owing to their maturity and originality. The papers are: "Utilization of Phosphorous Gypsum for the Production of Local Construction-binding Materials" by the Fourth-year-students of the Ivanovo Institute (see above): A. V. Tochilova and A. A. Fedyeva; "Study of the Influence of the Dispersion of Polymer Particles, When Being Disintegrated, on the Molecular Weight" by the Third-year-student of the Moskovskiy

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Chronicle. All-Union Competition for the Best Students-paper Concerning Chemistry and Chemical Technology for the Scholastic Year 1957-1958 SOV/153-2-2-31/31

tekhnologicheskiy institut legkoy promyshlennosti (Moscow Technological Institute for Light Industry) V. N. Goroditlov; "Study of the Cathodical Polarization at the Precipitation of Chromium From Sulphide-solutions" by the Fifth-year-student of the Ural'skiy politekhnicheskiy institut (Ural Polytechnical Institute) V. G. Petropavlovskiy; "Gold Extraction From Watery Cyanide-solutions" by the Fifth-year-students of the Moskovskiy khimiko-tekhnologicheskiy institut imeni D. I. Mendeleyeva (Moscow Chemical-technological Institute imeni D. I. Mendeleyev) A. V. Ochkin, V. A. Borisov, and N. Mrak; "Some Investigations of the Vulcanizates of Rubbers Containing Carboxyl" by the Fourth-year-students of the Yaroslavskiy tekhnologicheskiy institut (Yaroslavl' Technological Institute) G. I. Komarova and T. A. Shchadricheva; "Investigation of the Cathodic and Anodic Processes at Gold-plating" by the Fifth-year-student of the Leningradskiy tekhnologicheskiy institut im. Lensoveta (Leningrad Technological Institute imeni Lensovet) R. A. Nosova; "Spectral Determination of Molybdenum and Tungsten in Tri-hetero-polyacids" by the Third-year-student of the Kishinevskiy gosudarstvennyy universitet (Kishinev State University) V. A. Dagarev; "Capture of Dichlorine-ethane by Bone-fat in Foam-condition" by the Fourth-year-students of the Kazanskiy khimiko-tekhnologicheskiy institut (Kazan' Chemical-technological Institute) L. I. Tashina, N. A. Nurutdinov, and T. G. Sirazuev. Taken collectively, the competition has shown a high standard of the scientific research work in the circles of the Studencheskoye Nauchnoye obshchestvo (Scientific-student.-societies) of many universities.

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68114

SOV/78-5-1-23/45

AUTHORS: Zvyagintsev, O. Ye., Zakharov-Nartsissov, O. I., Ochkin, A. V.

TITLE: Solvation and Polymerization of Cyanoauric (I) Acid in Aqueous Solutions

PERIODICAL: Zhurnal neorganicheskoy khimii, 1960, Vol 5, Nr 1, pp 131-138 (USSR)

ABSTRACT: This article deals with the conditions of extraction of cyanoauric (I) acid and its salts from aqueous solutions by means of organic solvents. The authors investigated: 1) the dependence of $\frac{[H^+]}{\alpha}$ on the equilibrium concentration $[H^+]$ of the hydrogen ions in the extraction of $HAu(CN)_2$ by means of various alcohols (α = distribution coefficient of $HAu(CN)_2$ among aqueous and organic phase) for which they derived equation (6); 2) the dependence of the alcohol content of the aqueous phase upon the sulfuric acid concentration (Table 1). It was found that the solubility of alcohols in the aqueous phase decreases with rising concentration of H_2SO_4. 3) Furthermore, the authors studied the dependence of α upon the equilibrium concentration of $AuH(CN)_2$ in the organic phase (Table 2). It was found that

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68114

Solvation and Polymerization of Cyanoauric (I) Acid in Aqueous Solutions SOV/78-5-1-23/45

solvates of the form $HAu(CN)_2 \cdot xSol$ are formed by reaction of $HAu(CN)_2$ with the aliphatic alcohols under consideration. These solvates are present in the organic and aqueous phase alike. The distribution coefficient decreases with rising concentration of cyanoauric (I) acid in the aqueous phase, which is explained by the formation of polymers of the form $[HAu(CN)_2]_n$. Such dimers are present in the aqueous phase, and are not extracted by a 1:1 mixture of n-amyl alcohol or cyclohexanone and benzene. No polymers were detected in strongly dilute solutions. There are 1 figure, 2 tables, and 6 references, 5 of which are Soviet.

ASSOCIATION: Moskovskiy ordena Lenina khimiko-tekhnologicheskiy institut im. D. I. Mendeleyeva (Moscow "Order of Lenin" Institute of Chemical Technology imeni D. I. Mendeleyev)

SUBMITTED: July 9, 1959

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

ZAKHAROV-NARTSISSOV, O.I.; OCHKIN, A.V.

Extraction of nitric acid with triheptylamine. Zhur.neorg.khim. 6 no.8:1936-1939 Ag '61. (MIRA 14:8)
(Nitric acid) (Amines)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

ZVYAGINTSEV, O.Ye., ZAKHAROV-NARTSISSOV, O.I.; OCHKIN, A.V.

Extraction of gold hydrocyanic acid with n-trioctylamine. Zhur.neorg. khim. 6 no.8:1978-1979 Ag '61. (MIRA 14:8)
(Hydrocyanic acid) (Gold compounds) (Trioctylamine)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

ZVYAGINTSEV, O.Ye.; ZAKHAROV-NARTSISSOV, O.I.; OCHKIN, A.V.

Extraction of gold from cyanide solutions obtained in ore treatment with n-trioctylamine. Zhur. prikl. khim. 34 no. 12:2601-2605 D '61. (MIRA 15:1)

(Gold ores) (Cyanide process)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

ZAKHAROV-NARTSISSOV, O.I.; OCHKIN, A.V.

Extraction of certain strong monobasic acids with trioctylamine. Zhur.neorg.khim. 7 no.3:665-670 Mr '62. (MIRA 15:3)
(Acids, Inorganic) (Trioctylamine)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

FROLOV, Yu.G.; OCHKIN, A.V.

Problems involved in the mechanism of extraction. Zhur.neorg.-khim. 7 no.6:1486-1489 Je '62. (MIRA 15:6)
(Extraction (Chemistry))

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

SUDARIKOV, B.N.; FROLOV, Yu.G.; IL'ICHEV, V.A.; PUSHKOV, A.A.; ZAKHAROV-NARTSISSOV, O.I.; OCHKIN, A.V.

Physicochemical properties of some n-aliphatic amines. Trudy MKHTI no.43:21-28 '63. (MIRA 17:10)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

SUDARIKOV, B.N.; ZAKHAROV-NARTSISSOV, S.I.; OCHKIN, A.V.

Oxidation of tetravalent uranium in sulfuric acid solutions by atmospheric oxygen. Trudy MKHTI no.43:78-81 '63. (MIRA 17:10)

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85360

S/120/60/000/005/034/051
E032/E314

21.5200

AUTHORS: Ovechkin, V.V., Pirkin, I.A., Polyakov, A.S. and Ochkin, D.V.

TITLE: Method of Calibrating a Scintillation Gamma-spectrometer 19

PERIODICAL: Pribory i tekhnika eksperimenta, 1960, No. 5, pp. 126 - 128

TEXT: The conversion of the areas under the photopeaks in scintillation γ-spectra to the total intensities of γ-rays, for a medium-sized NaI(Tl) crystal and energies $E_\gamma \gtrsim 300$ keV, can only be carried out if the γ-ray spectrometer is calibrated in a preliminary experiment. This calibration is usually carried out with the aid of standard γ-ray sources with energies close to the energy of the γ-rays under investigation. However, such standard γ-ray sources are not always available. Calculated data suitable for calibration purposes and applicable to the many practical cases, are largely not available either. An absolute calibration curve covering a wide energy interval can be obtained for a scintillation γ-spectrometer with the aid of γ-sources whose intensities are

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E032/E314

Method of Calibrating a Scintillation Gamma-spectrometer

not standardised, provided each of them has a number of γ-lines with known intensity ratios, including soft γ-quanta with energies $\lesssim$ 100 keV (Ref. 1). By definition, the efficiency (relative aperture) of a γ-spectrometer for γ-rays of given energy is given by $\varepsilon = S/N = f(E_\gamma)$ where S is the counting rate in the photopeak and N is the total intensity of γ-rays of the given energy emitted into an angle of 4π. The ratio of efficiencies for hard and soft γ-rays emitted by a given source is then:

$$\varepsilon_i/\varepsilon_o = (S_i/S_o)(N_o/N_i) \qquad (1)$$

where the subscripts i and o refer to hard and soft rays, respectively. Since the soft γ-rays are absorbed in the surface layer of the NaI(Tl) crystal (for example, for E_γ = 90 keV, $\mu = 7.5\ cm^{-1}$), it follows that $\varepsilon_o = S_o/N_o \simeq \omega_o$

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Method of Calibrating a Scintillation Gamma-spectrometer

where ω_o is the relative solid angle subtended by the crystal at the source. The magnitude of S_o must of course be corrected for the absorption of soft quanta in the crystal envelope and in the source, as well as for the fraction of K x-rays of iodine which escape from the crystal (Ref. 2). Thus, the solid angle ω_o can be calculated from:

$$\omega_o = 1/2 \left\{1 - R/\sqrt{R^2 + a^2}\right\} \qquad (2)$$

where R is the distance from the source to the crystal (diameter 2a), S_i/S_o can be measured directly and N_i/N_o can be obtained from published data. Substituting the values for the various quantities in Eq. (1) for a number of γ-sources, one can obtain the calibration function $\varepsilon = f(E_\gamma)$ for a given geometry. The authors have measured this dependence for a NaI(Tl) crystal, 40 mm in diameter and 50 mm long, placed at a

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Method of Calibrating a Scintillation Gamma-spectrometer

distance R = 5.3 cm from the source ($\omega_o = 3.2 \times 10^{-2}$). The following γ-sources were employed:

Hg^{203}(E_o = 71 keV, E_1 = 279 keV, α_K = 0.159 (Ref.3), $N_o:N_1$ = 0.14)

Cs^{137}(E_o = 32 keV, E_1 = 661 keV, α_K = 0.11 (Ref.4), $N_o:N_1$ = 0.072)

$Ce^{144} + Pr^{144}$(E_o = 80, E_1 = 134, E_2 = 700, E_3 = 1490, E_4 = 2180 keV; $N_o:N_1:N_2:N_3:N_4$ = 7.2:15.3:3.56:0.56:1.44 (Ref. 5);

Se^{75}(E'_o = 140, E_1 = 270, E_2 = 400 keV; $N_o:N_1:N_2$ = 123:141:22.3 (Ref. 6).

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E032/E314

Method of Calibrating a Scintillation Gamma-spectrometer

In the first two cases the ratio $N_o:N_1$ was calculated from α_K taking into account the fluorescence emitted from the K-shell (Ref. 7). As can be seen from Fig. 1 all the experimental points except for the 134 keV γ-rays from Ce^{144} lie on the continuous curve and agree with the four calculated points E_γ = 280, 661, 1330 and 2620 keV (full circles) which were obtained by combining the data taken from Refs. 8 and 9 and applying them to our geometry. The experimentally determined function $\varepsilon = f(E_\gamma)$ was confirmed by control measurements using the following γ-ray sources Ra^{226} + daughter products (E_o' = 610, E_1 = 350, E_2 = 770, E_3 = 1120, E_4 = 1760, E_5 = 2200 keV, $N_o':N_1:N_2:N_3:N_4:N_5$ = 100:62.5:18.7:45.3:54.3:21.5 (Refs. 10,11)); I^{131}(E_o' = 640, E_1 = 364, E_2 = 720 keV, $N_o':N_1:N_2$ = 11.6:100:2.4 (Ref. 12).

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E032/E314

Method of Calibrating a Scintillation Gamma-spectrometer

Since the intensity of the γ-rays in the soft part of the γ-spectrum for these isotopes is not well known, ε was normalised so that $\varepsilon_{610} = 7.3 \times 10^{-3}$ and $\varepsilon_{640} = 7.0 \times 10^{-3}$. In addition, it was assumed that $\varepsilon_{140} = 2.7 \times 10^{-2}$ in the case of Se^{75}. The γ-ray spectrum was measured with the aid of a 100-channel kicksorter (AI-100 (AI-100)). A typical γ-spectrum (Ra^{226} in equilibrium with its decay products) is shown in Fig. 2. There are 2 figures and 12 references: 1 Swedish, 1 Italian, 4 English and 6 Soviet.

Acknowledgments are expressed to L.T. Polyakova for assistance in the measurements.

SUBMITTED: July 17, 1959

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

84397

S/056/60/0[illegible]/048
B004/B070

24.6720

AUTHORS: Nakhutin, I. Ye., Ovechkin, V. V., Ochkin, D. V., Polyakov, A. S., Khoduleva, Z. K.

TITLE: Preparation of the Radioactive Isotope Kr^{35} and Investigation of Its Gamma Radiation 79

PERIODICAL: Zhurnal eksperimental'noy i teoreticheskoy fiziki, 1960, Vol. 39, No. 4(10), pp. 991-992

TEXT: Kr^{35} was obtained by dissolving neutron-irradiated uranium in nitric acid and by separating chromatographically by active carbon at 77°K the gases liberated from moisture, nitrogen oxides, and radioactive iodine. For the measurement of emission, Kr^{35} was filled in a plexiglass cylinder with an aluminum foil bottom. The yield was determined from the ratio $k_\gamma = N_\gamma/N_\beta$, where N_γ, N_β are, respectively, the numbers of 517 kev gamma quanta and of β particles emitted per unit time in the solid angle 4π. The beta radiation was measured by an СИ-2Б (SI-2B) counter, and the gamma radiation by a NaI(Tl) scintillator

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

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Preparation of the Radioactive Isotope Kr^{35} and Investigation of Its Gamma Radiation

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B004/B070

and АИ-100 (AI-100)[28] analyzer. The gamma yield of Kr^{35} was found to be (0.41±0.06)% per decay. This value is significantly lower than that given by H. Zeldes et al. (Ref. 1). The authors checked the data by measurements on I^{131} and Cs^{134} whose gamma quantum yield is exactly known. There are 6 references: 3 Soviet and 3 US.

SUBMITTED: May 23, 1960

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APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001237730003-5

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to: Ochkin, D.V.

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