"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

L 13628-66

ACC NR: AP6003316

protection methods for metals using sodium nitrite and dicyclohexylammonium nitrite, and such corrosion inhibitors for both ferrous and nonferrous metals as salts of nitro- and dinitrobenzoic acids and certain amines. I. L. Rozenfel'd is associated with the Institute of Physical Chemistry, Academy of Sciences USSR. Orig. art. has: 10 tables. [ATD PRESS: 4169-F]

SUB CODE: 11 / SUBM DATE: 05May65 / ORIG REF: 039 / OTH REF: 050

Card 2/2

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

L 28529-66 EWP(j)/EWT(m)/EWP(t)/ETI IJP(c) RM/JD/WB/GD

ACC NR: AT6013803 (A) SOURCE CODE: UR/0000/65/000/000/0284/0295

AUTHOR: Rozenfel'd, I. L.; Persiantseva, V. P.; Reyzin, B. L.; Shustova, Z. F.; Gavrish, N. M.

ORG: none

TITLE: Investigation of certain nitrobenzoic amine salts as corrosion inhibitors for ferrous and nonferrous metals

SOURCE: Korrosiya metallov i splavov (Corrosion of metals and alloys), no. 2. Moscow, Izd-vo Metallurgiya, 1965, 284-295

TOPIC TAGS: amine salt, corrosion inhibitor, ferrous metal, nonferrous metal

ABSTRACT: The article presents the results of an investigation of the protective properties of certain inhibitors (nitro- and dinitrobenzoates) synthesized at the authors' laboratory; these properties were tested in natural as well as accelerated conditions involving cyclic and continuous exposure to moisture, with the aid of a specially developed device (Persiantseva, V. P., Rozenfel'd, I. L. Zavodskaya laboratoriya, 1958, 24, 7, 282). (The tests under natural conditions simulated the conditions under which metal products are stored in unheated warehouses and lasted for 21 months.) The inhibitors investigated were: hexamethyleneimine meta-nitrobenzoate, hexamethyleneimine ortho-nitrobenzoate, hexamethyleneimine 3,5-dinitrobenzoate, and piperidine 3,5-dinitrobenzoate. The coating of metal surface with

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an inhibitor was accomplished through adsorption from vapor phase or by washing the specimens in alcohol solutions of the inhibitors with subsequent drying at room temperature. Protective properties were determined according to the time elapsed until the appearance of first signs of corrosion and according to corrosion rate (as determined by gravimetric method). Findings: When applied in the form of alcohol solutions, all the four tested chemicals proved to be effective inhibitors of atmospheric corrosion under conditions simulating storage of metals in unheated warehouses, in industrial districts (where the atmosphere is more contaminated), for not only ferrous metals but also the most widely used nonferrous metals, (Cu and its alloys, Ag, Sn, Al and its alloys, Ni and Cr coatings, and Zn and Cd coatings passivated in a $K_2Cr_2O_7$ solution). These findings should represent a major advance considering that previously the only other known volatile inhibitors used in industry protected only ferrous metals. Orig. art. has: 7 tables and 1 figure.

SUB CODE: 11, 07 / SUBM DATE: 19Jul65/ ORIG REF: 004/ OTH REF: 002

Card 2/2

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

L 28537-66 EWP(j)/EWT(m)/T/EWP(t)/ETI IJP(c) RM/WW/JD/WB/GD

ACC NR: AT6013804 SOURCE CODE: UR/0000/65/000/000/0296/0304

AUTHOR: Rozenfel'd, I. L.; Persianteeva, V. P.; Gulyayeva, I. P.

ORG: none

TITLE: Protective properties of inorganic inhibitors in the presence of extraneous ions

SOURCE: Korroziya metallov i splavov (Corrosion of metals and alloys), no. 2. Moscow, Izd-vo Metallurgiya, 1965, 296-304

TOPIC TAGS: corrosion inhibitor, ion, sodium compound, sulfate, chloride, electrolyte

ABSTRACT: Inorganic inhibitors, which chiefly include sodium and potassium salts of acids whose anions contain atoms of elements V, IV and VII of the periodic table, are widely used to protect metals against corrosion in neutral media. Yet despite the numerous studies of these inhibitors, the interaction between inhibiting and aggressive media still has not been adequately investigated and hence the effectiveness of the inhibitors in various electrolytes cannot be predicted. To fill this gap, the authors investigated the corrosion rate of Fe as a function of inhibitor concentration for a fixed concentration of aggressive ions (Cl^-, SO_4^{2-}). In addition the protective concentrations of inhibitors for various concentrations of aggressive ions were investigated; this made it possible to derive a mathematical relation for

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calculating the required amount of inhibitor for real systems. Thus the dependence of the protective concentration of such compounds as Na_2WO_4, Na_3PO_4, Na_2MoO_4, NaH_2PO_4 on the content of sulfate ions (SO_4^{2-}) in the solution is expressed by the equation: $Y = 0.17X + 0.003$, where Y is molal inhibitor concentration, and X is molal concentration of aggressive ion. For the aggressive ion Cl^- the corresponding relation is: $Y = 0.38X + 0.005$. The accompanying electrochemical tests pertained to electrode impedance which, as was anticipated, differed depending on whether the inhibitor forms phase layers or adsorption layers at the surface of the protected metal. The corrosion rate of metal as a function of inhibitor concentration in a solution of 30 mg/liter NaCl and 70 mg/liter Na_2SO_4 was found to decrease in all cases -- except monosubstituted phosphate and sodium vanadate -- with increasing inhibitor concentration (Fig. 1). Thus, the investigated inhibitors may be arranged in the following series of increasing protective capacity:

$$Na_3PO_4 \approx Na_2HPO_4 > Na_2WO_4 > Na_2MoO_4 > NaH_2PO_4 > Na_2VO_4.$$

It is found that these inhibitors markedly alter impedance when the current applied is small, do not affect capacitance characteristics in the presence of high frequencies but somewhat reduce ohmic resistance at these frequencies. Such changes in impedance indicate that inhibitors of this type (XO_4^{n-}) alter the polarization characteristics of the system. Orig. art. has: 9 figures.

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L 28537-65

ACC NR: AT6013804

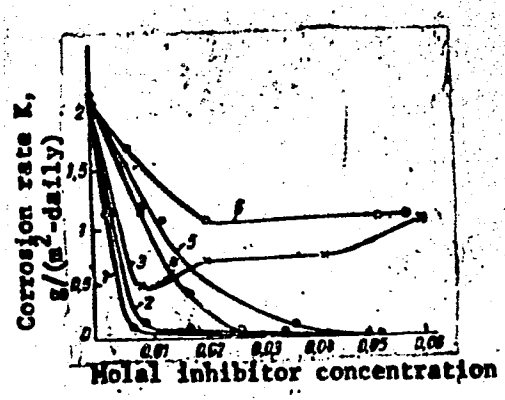

Fig. 1. Metal corrosion rate as a function of inhibitor concentration in solution containing 30 mg/liter NaCl and 70 mg/liter Na_2SO_4:

1 - trisodium phosphate; 2 - sodium tungstate; 3 - monosubstituted sodium phosphate; 4 - disubstituted sodium phosphate; 5 - sodium molybdate; 6 - sodium vanadate

SUB CODE: 11, 07 SUBM DATE: 19Jul65/ ORIG REF: 005/ OTH REF: 006

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

L 28536-66 EWP(j)/EWT(m)/T/EWP(t)/ETI IJP(c) RM/WW/JD/WB/GD

ACC NR: AT6013805 SOURCE CODE: UR/0000/65/000/000/0305/0314

AUTHOR: Rozenfel'd, I. L.; Parsiantseva, V. P.

ORG: none

53
B+1

TITLE: Mechanism of the inhibition of metal corrosion by ammonia

SOURCE: Korroziya metallov i splavov (Corrosion of metals and alloys), no. 2. Moscow, Izd-vo Metallurgiya, 1965, 305-314

TOPIC TAGS: corrosion inhibitor, ammonia, steel, surface ionization, hydrolysis

ABSTRACT: Recently, to improve the effectiveness of the corrosion protection of metals in storage and transport, ammonia salts and other compounds capable of releasing ammonia during hydrolysis and dissociation have begun to be added to the standard inhibitors (e.g. $NaNO_3$). It is generally assumed that the mechanism of protective effect may then manifest itself in one of the following phenomena: a) NH_3 displaces the pH of the medium to the region of low corrosive activity; b) NH_3 (or NH_4OH) binds the aggressive component of the medium (e.g. CO_2); c) the mixing of NH_3 salts with $NaNO_3$ leads to the formation of the highly volatile NH_4NO_3. Since these assumptions lacked experimental proof, the authors decided to investigate the protective capacity of NH_3 as a function of its content in the atmosphere, as well as to investigate the effect of NH_3 on the electrochemical behavior of steel. The in-

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vestigation pertained to NH_3 of 100% relative humidity at 25°C, on varying its concentration by placing in the reaction vessel NH_4OH of various concentrations corresponding to ammonia vapor pressures of from 66.66 to 666.6 n/m^2 (0.5 to 5 mm Hg), with the corrosion being determined as a function of time until the appearance of first signs of corrosion on the steel specimens. These experiments showed that even insignificant concentrations of NH_3 in the atmosphere (266.64 n/m^2 or 2 mm Hg) can effectively protect steel against corrosion. An attendant investigation of the electrochemical behavior of steel in NH_3 atmosphere showed that the presence of NH_3 displaces the steel's potential by 500 mv in the positive direction. The mechanism of the protection of steel by NH_3 is due to the hydration of NH_3 and the concomitant formation of NH_3 ions. The prior adsorption of hydroxyl ions on the (positively charged) surface of steel makes possible the adsorption of NH_3 ions on the metal surface, which leads to a sharp change in the ionization rate of Fe. These findings are also highly significant to the determination of the mechanism of the protective effect of other amine-class compounds which, in solutions, form cations similar to the NH_3 ion. Orig. art. has: 7 figures and 4 formulas.

SUB CODE: 11,11,10720/ SUBM DATE: 19Jul65/ ORIG REF: 005/ OTH REF: 003

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APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFEL'D, I. L.; PERSIANTSEVA, V. P.

New priciple of anticorrosive protection by inhibitors. Dokl. AN SSSR 156 no. 1:162-165 My '64. (MIRA 17:5

1. Institut fizicheskoy khimii AN SSSR. Predstavleno akademikom A. N. Frumkinym.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ACCESSION NR: AP4035821 S/0020/64/156/001/0162/0165

AUTHOR: Rozenfel'd, I. L.; Persiantseva, V. P.

TITLE: On a new principle of anticorrosive protection by inhibitors

SOURCE: AN SSSR. Doklady*, v. 156, no. 1, 1964, 162-165

TOPIC TAGS: anticorrosive protection, corrosion inhibitor, anode reaction decrease, cathode reaction increase, stationary potential, iron stationary potential, hexamethyleneimine 3,5 dinitrobenzoate, organic cation adsorption, surface ion adsorption, passivation, dinitrobenzoic acid salts, dicyclohexylamine nitrite

ABSTRACT: This consists essentially in obtaining full passivation potential not by decreasing the reaction rate at the anode but by sharply increasing that at the cathode; it requires inhibitor reduction with sufficient speed in a neutral medium. Organic salts are capable of shifting the stationary potential of iron by 300-400 millivolts towards the positive side without significantly changing the rate of metal ionization at the anode; this applies in particular to carboxyl and secondary NO_2 groups attached to the benzene ring. The newly synthesized inhibitor hexamethyleneimine 3,5-dinitrobenzoate proved easily reducible and highly effective.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

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Another requirement for protection, besides electrochemical reduction, is the adsorption of organic cations by the metal surface. If this adsorption reaches a significant level the specific current will easily attain the critical value required for the passive iron state. This requirement is met by organic salts of dinitrobenzoic acid containing easily absorbed organic cations and reducing the nitrite group with sufficient speed. This new principle is applicable with electrolytes containing a low activator concentration (chlorine ions). The chemical reactions involved are presented and discussed for 2 inhibitors representing the 2 kinds of protection: dicyclohexylamine nitrite and the above-mentioned compound. The first shifts the stationary potential towards positive due to considerable checking of anode reaction; the second achieves the same by accelerating cathode depolarization. In dilute electrolytes this latter will also decelerate anode ionization reaction. A new addition to inhibitor classification is thus required, i.e. the class of compounds affording metal protection by accelerating the process at the cathode. Orig. art. has: 3 figures and 2 formulas.

ASSOCIATION: Institut fizicheskoy khimii Akademii nauk SSSR (Institute of Physical Chemistry, Academy of Sciences, SSSR)

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ACCESSION NR: AP4035821

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SUB CODE: MM, OC NO REF SOV: 002 OTHER: 000

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, V.P.

KUDRYAVTSEV, N.T.; PERSIANTSEVA, V.P. and KALB, V.M.

"High-luster Copper Plating From Acid Electrolytes." Nauchno-tekhnicheskoye obshchestvo mashinostroitel'noy promyshlennosti, Kiyevskoye oblastnoye pravleniye. Zashchitno-dekorativnyye i spetsial'nyye pokrytiya metallov (Protective, Decorative, and Special Coatings for Metals) Kiyev, Mashgiz, 1959. 291 pp. (pp 87)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

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S/080/61/034/009/010/016
D204/D305

AUTHORS: Rozenfel'd, I.L., Persiantsyeva, V.P., Terent'yev, P.B. and Polteva, M.N.

TITLE: Investigating the influence of chemical composition and structure of organic compounds on their ability to retard the corrosion process

PERIODICAL: Zhurnal prikladnoy khimii, v. 34, no. 9, 1961, 2047 - 2056

TEXT: This is report I from the series of papers on investigating the mechanism of protection of metals against corrosion by volatile inhibitors. The results of an investigation of the dependence of protective properties of various classes of compounds on their structure and the presence of the functional groups OH, NO_2, NH_2 and complex organic radicals, are reported. In order to carry out these investigations, accelerated methods were developed for testing the protective properties of the compounds, for determining

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the pressures of the saturated vapors of volatile inhibitors and the electrochemical behavior of metals under thin films of electrolytes in an atmosphere of volatile inhibitors. The investigation of the protective properties of volatile inhibitors was carried out by imitating corrosion under natural conditions whereby alternate condensation and drying of electrolytes on metal surfaces takes place. The study was carried out in an atmosphere of 100 % relative humidity with 5 cycles of condensation of moisture on the specimens per day. Organic nitrous bases and their salts with weak organic and inorganic acids, complex esters of acids, and inorganic ammonium salts were studied. The protective properties of the compounds were considered to be satisfactory, if no observable corrosion products had formed after 10 days of accelerated tests. It was found that the protective properties of amine salts are determined not only by the radical and the functional group, and thus by the composition of the compound, but also by their structure, on which their adsorptive ability evidently depends. Complex esters of acids and weak aromatic amines cannot be

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Investigating the influence of ...

used as volatile inhibitors, since the former retard corrosion of steel only slightly and the latter not at all. The protective properties of volatile inhibitors are independent of the hydrogen ion concentration established in the moisture film after the latter is saturated with inhibitor vapors. There are 1 figure, 7 tables and 7 references: 4 Soviet-bloc and 3 non-Soviet-bloc. The references to the English-language publications read as follows: H.R. Backer, Ind. Eng. Ch., 46, 12, 2592, 1954; A. Wachter, T. Sky, N. Stillman, Corrosion, 7, 9, 284, 1951; W.D. Harki, D. Florence, J. Phys. Chem. 6, 847, 1938.

SUBMITTED: July 18, 1960

Card 3/3

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFEL'D, I.L.; RUBINSHTEYN, F.I.; YAKUBOVICH, S.V.; PERSIANTSEVA, V.P.; Prinimali uchastiye: GILLER, R.S.; KURSKAYA, A.G.

Studying chrome acid guanidine as a corrosion inhibitor for oil paints. Lakokras.mat.i ikh prin. no.3:15-21 '62. (MIRA 15:7)
(Protective coatings)
(Guanidine)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFEL'D, I.L.; POLTEVA, M.N.; PERSIANTSEVA, V.P. (Moskva)

Physicochemical properties of dicyclohexylamine nitrite, a volatile corrosion inhibitor. Zhur. fiz. khim. 35 no.7:1474-1477 Jl '61. (MIRA 14:7)

1. AN SSSR, Institut fizicheskoy khimii.
(Dicyclohexylamine)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

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D204/D305

AUTHORS: Rozenfeld, I.L., Polteva, M.N., Persiatseva, V.P., and Terent'yev, P.B.

TITLE: Pressure of saturated volatile inhibitor vapors

PERIODICAL: Zhurnal prikladnoy khimii, v. 34, no. 9, 1961, 2056 - 2061

TEXT: This is report II of a series of papers on investigating the mechanism of protection of metals against corrosion by volatile inhibitors. One of the important characteristics of volatile inhibitors is their saturated vapor pressure. Compounds having high vapor pressure are most effective. For the successful application of such inhibitors, the temperature dependence of the pressure of the saturated vapor must also be known. The inclination of the straigth line obtained by plotting negative logarithm of pressure of saturated vapors against 1/T enables the changes of pressure with temperature to be determined, and the temperature range in which an in-

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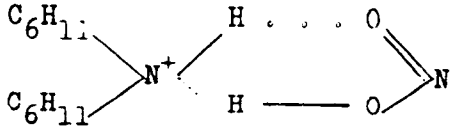
hibitor is effective to be defined. By means of the Knudsen method, the temperature dependence of the pressure of saturated vapors of the volatile inhibitors dicyclohexylamine nitrate and morpholene cinnamate was investigated. On the basis of this dependence, the value of the latent heat of sublimation for di-cyclohexylamine nitrate was calculated (25 Kcal/mol). From a comparison of the value of the latent heat of sublimation and the dipole moment, it is proposed that the structure of di-cyclohexylamine in the vapors is as follows:



There are 3 figures, 1 table and 10 references: 3 Soviet-bloc and 7 non-Soviet-bloc. The references to the 4 most recent English-language publications read as follows: E.G. Stroud, W.H.I. Vernon, J. Applied Chem., 2, 166, 1952; A. Wachter, T. Sky, N. Stilliman,

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

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D204/D305

Pressure of saturated volatile ...

Corrosion, 7, 9, 284, 1951; E.G. Stroud, W.H.I. Vernon, U.K. Pat. 691109, 1951; H. Patzelt, Corrosion, 9, 1, 19, 1953.

SUBMITTED: July 18, 1960

Card 3/3

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

S/081/62/000/023/053/120
B124/B101

AUTHORS: Rozenfel'd, I. L., Rubinshteyn, F. I., Yakubovich, S. V., Persiantseva, V. P.

TITLE: Study of guanidine chromate as a corrosion inhibitor in oil paints

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 23, 1962, 411, abstract 23I334 (Lakokrasochn. materialy i ikh primeneniye, no. 3, 1962, 15-21)

TEXT: A new way to increase the protective qualities of pigmented coatings by means of modifying inert fillers and film-forming materials with corrosion inhibitors (CI) is suggested. The effect of organic CI on the properties of the oil paints was examined. It has been shown that guanidine chromate (GC) has strong passivating properties and that its effect on the oil coating is to inhibit metal ionization by anodic reaction. Conditions for obtaining corrosion-resisting oil paints are determined, with GC used as the CI. [Abstracter's note: Complete translation.]

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ROZENFELD, I. L. [Rozenfel'd, I.L.]; PERSIANTEVA, V.P. [Persiantseva, V.P.]; TERENTIEV, P.B. [Terent'yev, P.B.]; POLTEVA, M.N.; KUZNETOVA, M.M. [Kuznetsova, M.M.]

Studies on the influence of chemical composition, structure and certain physicochemical properties of the organic compounds upon their capacity of braking the corrosion process. Analele chimie 17 no.3:175-196 Jl-S '62.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFELD, I.L., PERSIANTSEVA, V.P., TERENTIYEV, P.B.

"Mechanism of metal protection from corrosion with the aid of volatile inhibition."

Report submitted to the Second Intl. Congress on Corrosion of Metals New York City 11-15 March 1963

INSTITUTE OF PHYSICAL CHEMISTRY, MOSCOW

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, V.P.; ROZENFEL'D, I.L.

Laboratory methods of investigating volatile inhibitors. Trudy Inst.fiz.khim. no.7:41-50 '59. (MIRA 13:5)
(Inhibition (Chemistry)--Testing)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

S/137/62/000/004/143/201
A060/A101

AUTHORS: Persiantseva, V. P., Rozenfel'd, I. L., Novitskaya, M. A., Akimova, T. I., Labutin, A. L.

TITLE: The action mechanism of volatile inhibitors

PERIODICAL: Referativnyy zhurnal, Metallurgiya, no. 4, 1962, 102 - 103, abstract 4I625 ("Vestn. tekhn. i ekon. inform. N.-i. in-t tekhn.-ekon. issled. Gos. kom-ta Sov. Min. SSSR i khimii", 1961, no. 2, 68 - 76)

TEXT: An investigation was carried out upon the action mechanism of volatile inhibitors by studying the adsorption processes and the action of adsorption layers upon the process kinetics of electrochemical reactions. The protective properties of a large number of compounds were preliminarily studied by the methods of accelerated tests, extended tests and testing on paper. All the tests were carried out at 100% relative humidity and at a temperature of 30°C, and also in a warehouse location. The volatile inhibitor is adsorbed by the metallic surface in the form of molecules or ions, which form as result of hydrolysis in the water film of the electrolyte (complex organic cations, hydroxyl groups, or acid

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The action mechanism of volatile inhibitors

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A060/A101

radicals). These absorbed groups in some cases slow down the rate of only the anodic reaction, and in others - of both the anodic and the cathodic one. The characteristics determining the effectiveness of volatile inhibitors are: vapor pressure, absorptive power, and the strength of the bond between the inhibitor or protective group and the metallic surface, and also the rate of inhibitor retardation of the electrochemical reactions causing the corrosion process. There are 11 references.

V. Tarisova

[Abstracter's note: Complete translation]

Card 2/2

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFEL'D, I.L.; PERSIANTSEVA, V.P.; KUZNETSOVA, M.M.; POLTEVA, M.N.; TERENT'YEV, P.B.

Electrochemical behavior of metals in the atmosphere of volatile inhibitors. Zhur.prikl.khim. 34 no.10:2239-2244 O '61. (MIRA 14:11)

(Metals) (Electrochemistry) (Inhibition (Chemistry))

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

5(4)

AUTHORS: Rozenfel'd, I. L., Persiantseva, V. P. SOV/20-122-2-26/42

TITLE: The Influence of the Adsorption of Volatile Inhibitors on the Electrochemical Behavior of Iron (Vliyaniye adsorbtsii letuchikh ingibitorov na elektrokhimicheskoye povedeniye zheleza)

PERIODICAL: Doklady Akademii nauk SSSR, 1958, Vol 122, Nr 2, pp 260-263 (USSR)

ABSTRACT: This paper deals with the investigation of the electrochemical kinetics in thin electrolyte layers according to the method of I. L. Rozenfel'd (Ref 3). This method was improved in such a manner that it was possible to apply electrolyte films to the electrode surface after it had adsorbed the inhibitor from the gaseous phase without violating the hermeticity of the space. The volatile inhibitors investigated in this paper are benzylamine (C_7H_9N) and morpholine, which protect iron well from corrosion. A diagram shows the dependence of the potential of iron on the time of the previous soaking of the electrode in an atmosphere saturated by benzyl-

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amine. The adsorption of the inhibitor moves the steady potential towards the positive side. If the time of the soaking of the electrode in the atmosphere saturated by the inhibitor vapors increases, the observed effect increases, too. After a certain time, this effect acquires a constant value. The surface, obviously, is then totally saturated with the inhibitor and the potential is displaced by more than 200 mV. If there is a thin layer of electrolyte on the surface of the metal, the adsorption of the same inhibitor from the gaseous phase causes by far less intense effects. In this case, the potential of the iron is displaced only by 50-60 mV. The adsorption of the inhibitor from the gaseous phase, therefore causes more noticeable variations of the potential and, possibly, also a higher passivation degree of the electrode. The adsorption connection of the inhibitor with the metal surface and the stability of the passive state may in a certain degree be characterized by data concerning the time dependence of the potential of the metal after the saturation of the electrolyte film attached to this surface. If the time of the previous soaking of the electrode in the atmosphere com-

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

The Influence of the Adsorption of Volatile Inhibitors on the Electrochemical Behavior of Iron SOV/20-122-2-24/42

posed by the inhibitor vapors increases, the potential of the metal becomes more and more stable, and no activating influence of the electrolyte is observed. Also the adsorption of morpholine has a similar influence upon the electrochemical behavior of the iron electrode. According to the above-discussed experiments and considerations, the volatile inhibitors also are adsorbed by the metal surface, and they change their electrochemical properties. The greatest displacement of the steady potential is observed after a previous adsorption of an inhibitor by the electrode from the gaseous phase. In this case, the maximum saturation of the surface with the inhibitor and a marked stability of the passive state are observed. There are 4 figures and 5 references, 3 of which are Soviet.

ASSOCIATION: Institut fizicheskoy khimii Akademii nauk SSSR (Institute of Physical Chemistry, AS USSR)

Card 3/4

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, V. P., POLTEVA, M. V., TERENTYEV, P. B., ROZENFELD, I. L., Moskva:

"A Study On The Mechanism Of Metals Protection Against Corrosion By Volatile Inhibitors".

report submitted for the European Symposium on Corrosion Inhibitors, Ferrara Italy, 29 Sep-1 Oct 1960.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, V.P.; TITOV, P.S.

Electrolytic brass plating from pyrophosphate electrolytes. Nauch. dokl. vys. shkoly; khim. i khim. tekh. no.3:584-587 '58. (MIRA 11:10)

1. Predstavlena kafedroy elektrokhimii i korrozii Moskovskogo instituta tsvetnykh metallov i zolota imeni M.I. Kalinina.
(Brass) (Electroplating)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, V.P.

PHASE I BOOK EXPLOITATION SOV/4271

Akademiya nauk SSSR. Institut fizicheskoy khimii

Issledovaniya po korrozii metallov. [Vyp.] 5: Novyye metody i pribory dlya korrozionnykh ispytaniy (Investigations on Corrosion of Metals [No. 5]: New Methods and Instruments for Corrosion Testing) Moscow, Izd-vo AN SSSR, 1959. 176 p. (Series: Its: Trudy, Vyp. 7) Errata slip inserted. 5,000 copies printed.

Resp. Ed.: N. D. Tomashov, Doctor of Chemistry, Professor; Ed. of Publishing House: Z. G. Yegorov; Tech. Eds: G. A. Astaf'yeva and Ye. V. Zelenzova; Editorial Board: N. D. Tomashov, A. V. Byalobzheskiy, Candidate of Chemistry, and P. V. Shchigolev, Candidate of Chemistry.

PURPOSE: This collection of articles is intended for scientific workers at research institutes and technical personnel of plant laboratories.

COVERAGE: The articles included in this collection deal basically with methods of corrosion investigation which have not yet been published in Soviet periodical literature but are of definite interest for studying corrosion processes. A wide range of problems is covered. In addition to the methods discussed the articles provide some experimental data which make possible full utilization of each individual method. No personalities are mentioned. References accompany each article.

Klark, G.B., M.I. Mikhaylovskaya, Ye. N. Mikhaylovskiy, and N.D. Tomashov. Electrochemical Method for Investigating Atmospheric Corrosion of Metals 11

Rozenfel'd, I.L., T.I. Pavlutskaya, K.A. Shigalova, and T.I. Akimova. Methods of Electrochemical and Corrosion Investigations in Thin Layers of Electrolytes 22

Persiantseva, V.P., and I.L. Rozenfel'd. Laboratory Methods for Investigating Volatile Inhibitors 41

Mirolyubov, Ye.N., M.M. Kurtepov, and N.D. Tomashov. A Method for Obtaining Anodic Polarization Curves by Means of Cathodic Polarization 51

Bozin, N.A., and M.M. Kurtepov. Electrochemical Method for the Rapid Evaluation of the Corrosion Resistance of Metals 54

Glukhova, A.I. Investigation by Means of Profilograms of Changes in the Microgeometry of Stainless Steel Surfaces During Corrosion 60

Tomashov, N.D., V.N. Modestova, and G.K. Blinchevskiy. Methods for Investigating the Corrosion and Electrochemical Behavior of Metals Under Stress 64

Tomashov, N.D., and N.I. Isayev. Use of the Resistance-Capacitance Method for Investigating the Behavior of Protective Films During the Corrosion of Metals Under Stress 78

Card 3/6

(18

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFEL'D, I.L., prof.; PERSIANTSEVA, V.P., kand.tekhn.nauk

Corrosion inhibitors. Khim. nauka i prom. 3 no.4:500-505 '58.
(MIRA 11:10)

(Corrosion and anticorrosives)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

AUTHORS: Persiantseva, V. P., Titov, P. S. SOV 156-[illegible]

TITLE: The Electrolytic Production of Brass From Pyrophosphate Electrolytes (Elektroliticheskoye latunirovaniye iz pirofosfatnykh elektrolitov)

PERIODICAL: Nauchnyye doklady vysshey shkoly, Khimiya i khimicheskaya tekhnologiya, 195 , Nr 3, pp. 5[illegible] - 5[illegible]7 (USSR)

ABSTRACT: The conditions for the production of brass by electrolysis from solutions containing copper pyrophosphate and zinc pyrophosphate were investigated. It was found that in the dissolution of copper pyrophosphate and zinc pyrophosphate in solutions of sodium pyrophosphate complex ions are formed which have the following composition: $[MeP_2O_7]^{2-}$ and $[Me(P_2O_7)_2]^{6-}$.

The dependence of the quality of the electrolytic brass surface on such factors as the concentration of copper and zinc solutions, on the current density, the temperature, the mixture of the electrolytes during electrolysis, as well as on the acidity of the solution was investigated. It was found that the following conditions are necessary for the production of a good brass surface: the concentration of zinc and copper in the electrolyte

Card 1/2

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

The Electrolytic Production of Brass From Pyrophosphate Electrolytes SOV/156-[illegible]-[illegible]-[illegible]/52

must not be lower than 0,5 - 0,45 mole/l; the concentration of the free phosphate solution should be 0,15 - 0,[illegible] N; the pH value of the electrolysis solution should be 8 - 9,4; the temperature of the solution should be 35 - 40° C, and the current density should be 0,7 - 1 V(?). The chemical analysis, and the microscopic and radiographic investigations of a brass surface so produced showed that this surface had mono-phase structure. There are 2 figures, 3 tables, and 4 references, which are Soviet.

ASSOCIATION: **Kafedra** Elektrokhimii i korrozii Moskovskogo instituta tsvetnykh metallov i zolota im.M.I.Kalinina (Chair of Electrochemistry and Corrosion at the Moscow Institute of Nonferrous Metals and Gold imeni M.I.Kalinin)

SUBMITTED: January 17, 1958

Card 2/2

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

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SOV/81-59-9-31685

Translation from: Referativnyy zhurnal. Khimiya, 1959, Nr 9, p 308 (USSR)

AUTHORS: Turkovskaya, A.V., Persiantseva, V.P.

TITLE: The Effect of Some Factors on the Rate of Atmospheric Corrosion of Magnesium and Its Alloys

PERIODICAL: Sb. nauchn. tr. Nauchno-tekhn. o-vo tsvetn. metallurgii, Mosk. in-t tsvetn. met. i zolota, 1958, Nr 29, pp 172 - 178

ABSTRACT: The author studied the effect of moisture, CO_2, SO_2 and the presence of solid dust particles on the rate of atmospheric corrosion (RAC) of Mg and the Mg-alloy of the composition (in %): Al 6.5, Mn 0.3, Zn 1, which has been passivated in an acidic bichromatic bath containing NH_4Cl. The study has shown that RAC of these materials increases sharply at an increase in the relative humidity to 90 - 95%. An intensification of RAC of Mg was noted in an atmosphere polluted by CO_2 and SO_2, in which case the joint presence of both gases leads to higher corrosion losses than in an atmosphere containing these gases separately. Particles of coal dust promote the corrosion in SO_2 and almost do not affect the corrosion rate in an atmosphere polluted by CO_2. A. Shatalov

Card 1/1

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31862

S/081/62/000/003/042/90
B156/B101

18.8310

AUTHORS: Persiantseva, V.P., Rozenfel'd, I. L., Novitskaya, M.A., Akimova, T.I., Labutin, A.L.

TITLE: Mechanism by which volatile inhibitors work

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 3, 1962, 327-328, abstract 31211 (Vestn. tekhn. i ekon. inform. N.-i in-t tekhn.-ekon. issled. Gos. kom-ta Sov. Min. SSSR i khimii, no. 2, 1961, 68-76

TEXT: Research into the protective properties of a large number of compounds used as volatile corrosion inhibitors (VCI) has revealed a number of VCI which are effective at protecting steel and nonferrous metals from corrosion (a table is included). Study of the electrochemical behavior of steel in the presence of VCI has shown that a potential shift characteristic of adsorption of VCI by the metal surface takes place. The effects of four VCI are examined in detail; these are benzyl amine, morpholine, dicyclohexyl amine nitrite, and cyclohexyl amine carbonate. It has been found that VCI is adsorbed in the form of molecules or ions

Card 1/2

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFEL'D, I.L.; PERSIANTSEVA, V.P.; TERENT'YEV, P.B.; POLTEVA, M.N.

Effect of the chemical composition and structure of organic compounds on their capacity to inhibit corrosion processes. Zhur.prikl.khim. 34 no.9:2047-2056 S '61. (MIRA 14:9)
(Corrosion and anticorrosives)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ROZENFEL'D, I.L.; POLTEVA, M.N.; PERSIANTSEVA, V.P.; TERENT'YEV, P.B.

Saturated vapor pressure of volatile inhibitors. Zhur.prikl.khim. 34 no.9:2056-2061 S '61. (MIRA 14:9)
(Inhibition (Chemistry)) (Vapor pressure)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, V.P.; KUDRYAVTSEV, N.T.; KAL', V.M.

Bright copperizing of shaped hardware using acid electolytes. Metalloved. i obr. met. no.9:42-46 S '58. (MIRA 11:10)

1.Moskovskiy khimiko-tekhnologicheskiy institut imeni Mendeleyeva.
(Protective coatings) (Electrolytes)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

Persiantseva, V.P.

AUTHORS: Persiantseva, V.P., Shnayder, V.A.

32-11-9/60

TITLE: Colorimetric Determination of the 2.6-, 2.7-Naphtalene Disulphide Acid Content in a Nickel Electrolyte (Kolorimetricheskoye opredeleniye 2.6-, 2.7-naftalindisul'fokisloty v nikelevom elektrolite)

PERIODICAL: Zavodskaya Laboratoriya, 1957, Vol. 23, Nr 11, pp. 1298-1298 (USSR)

ABSTRACT: The acid mentioned is characterized by the fact that it causes the "phenomenon of mould" in the nickel electrolyte. For the purpose of determining it, it is recommended to carry out separation of nickel from the electrolyte chromographically by the application of the cationite "SDV-3". For this purpose an adsorber was mounted in a cylindrical vessel (filter) of 15 mm diameter up to a height of 200 mm. The electrolyte was made to pass through the adsorber with a velocity of 100 ml per 20 min. After nickel adsorption was completed the filter was rinsed out while the same velocity was maintained for 10 minutes. The analysis itself was then carried out: 5 ml of the electrolyte was dissolved in distilled water of up to 20-25 ml total volume and sent through the adsorption filters. The filters were rinsed out with distilled water; the solution together with the rinsing water was introduced into a platinum vessel, where 10 ml

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

Colorimetric Determination of the 2.6-, 2.7-Naphtalene Disulphide Acid Content in a Nickel Electrolyte 32-11-9/60

of the 30% NaOH solution was added. The whole content was vaporized until the sediment became dry, which was then melted at 320 ± 2°. To the cooled melt 35 ml of nitrating solution; 10 g nitrate of ammonium in 100 ml hydrochloric acid having the specific weight of 1.84-1.82 was added. After the strong reaction hereby caused had ended, distilled water up to a volume of 100 ml was added. The solution, which in this case was of yellow color, was colorimetrized by a blue light filter. According to the previously prepared gauge curve of determination, the naphtalene sulphide acid content (volume) was determined. The method described is distinguished by its high accuracy. There is 1 table.

ASSOCIATION: Chemical-Technological Institute imeni D.I.Mendeleyev (Khimiko-tekhnologicheskiy institut im. D.I.Mendeleyeva)

AVAILABLE: Library of Congress

Card 2/2

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

VIDOVIC, Milan, dr.; PERSIC, Davor, Mr Ph.

Gastric carcinoma and serum alkaline phosphatase. Lijecn. vjesn. 84 no.8:773-779 '62.

1. Iz Kirurske klinike Medicinskog faktulteta u Zagrebu.
(BLOOD ALKALINE PHOSPHATASE) (STOMACH NEOPLASMS)
(ENZYME TESTS)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

YUGOSLAVIA

Dr Milan VIDOVIC and Mr Ph [illegible] PERSIC, Surgical Clinic [illegible] Faculty (Kirurska klinika Medicinskog fakulteta) University [illegible]

"Gastric Carcinoma and Serum Alkaline Phosphatase."

Zagreb, Lijecnicki Vjesnik, Vol 84, No 8, Aug 1962, pp [illegible]

Abstract [English summary modified]: Study of 142 patients with gastric cancer, correlating hepatic metastases with increased alkaline phosphatase levels. Authors suggest that Shinowara et al. pegged upper limit of norm too high, should be 6. Provided all other possible causes of an increase in alkaline phosphatase level are excluded, detection of it is a strong indication of presence of hepatic metastases. There are occasional false negatives, and false positives cannot be ruled out. Diagram, 3 tables, 31 Western and 3 Yugoslav references.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

RETLHIN, R.; BAZEVIC, D.; BUCAN, N.; PERSIC, N.; RADOSEVIC, Z.

On the attitude toward psychiatric patients. Neuropsihijatrija 9 no.4:273-285 '61.

1. Iz Neurolosko-psihijatrijske klinike Medicinskog fakulteta u Zagrebu (Predstojnik: Prof. dr R. Lopasic)

(MENTAL DISORDERS)

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, N.; VIREK, Z.; PALMOVIC, R.; ARKO, K.; BETRIANI, M.

On some modifications of electric shock therapy. Neuropsihijatrija 9 no.2/3:166-177 '61.

1. Iz Neurolosko-psihijatrijske klinike (Predstojnik: Prof. dr R. Lopasic) i Kirurske klinike Medicinskog fakulteta u Zagrebu (Predstojnik: Prof. dr D. Juzbasic).
(SHOCK THERAPY ELECTRIC)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, Nikola, dr.

Psychiatry in the Soviet Union. Lijecn. vjesn. 85 no.7:747-758 '63.

1. Iz Neurolosko-psihijatrijske klinike Medicinskog fakulteta u Zagrebu.

(PSYCHIATRY)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, Nikola, dr.; KOPORCIC, Petar, dr.

Polyclinical and dispensary psychiatric services in Croatia. Lijecn. vjesn. 83 no.5:445-458 '61.

1. Iz Neurolosko-psihijatrijske klinike Med. fakulteta Sveucilista u Zagrebu u Zagrebu i Zavoda za socijalno osiguranje NR Hrvatske u Zagrebu.

(PSYCHIATRY)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, Nikola

On the possibility of the use of electric stimulation in insulin hypoglycemia. Rad Jugosl. akad. znan. umj.[Med]323:65-149 '61.

(SHOCK THERAPY INSULIN) (ELECTROTHERAPY)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

ASPERGER, Zdravko, dr.; PERSIC, Tomislav, dr.

Accidental hypothermia. Lijecn. vjesn. 85 no.4:423-430 '63.

1. Iz Interne poliklinike Medicinskog fakulteta u Zagrebu.
(HYPOTHERMIA)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, Zvonko, d-r

A contribution to the treatment of complicated crural fractures.
Voj.san.pregl., Beogr. 17 no.9:927-930 S '60.

1. Vojnomedicinska Akademija u Beogradu, Hirurska klinika
(LEG fract & disloc)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PAPO, Isidor, generalmajor sanitetske sluzbe profesor dr; PERSIC, Zvonko, dr

Pectus excavatum. Voj.san.pregl., Beogr. 17 no.10:1040-1045 O '60.

1. Vojnomedicinska Akademija u Beogradu, Klinika za hirurske bolesti
(THORAX abnorm)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, Zvonko, dr

Severe injury of theleg. Voj.san.pregl., Beogr. 18 no.1:72-74 Ja '61.

1. Vojnomedicinska Akademija u Beogradu, Hirurska klinika.
(LEG wds & inj)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PIRSIC, Zvonko, dr.

Stercoral fistula in postoperative scar after appendectomy as a complication of sigmoid diverticulitis. Vojnosanit. pregl. 18 no.9: 802-804 S '61.

1. Vojnomedicinska akademija u Beogradu, Klinika za hirurske bolesti.

(APPENDECTOMY compl) (DIVERTICULITIS compl)
(INTESTINAL FISTULA etiol)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIDSKAYA, L.

Derivation of a bounded solution to a certain functional equation. Vest. AN Kazakh. SSR 20 no.1:77-80 Ja '64.
(MIRA 17:3)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIDSKIY, K.P., akademik

Abstract holomorphic functions. Vest. AN Kazakh. SSR 14 no.1: 91-97 Ja '58. (MIRA 11:2)

1.AN KazSSR.

(Functions)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIDSKIY, A. S. (Engineer)

"Argon arc welding of flanged pieces of siifons made from thin-foilage steel 1Kh18N9T δ = 0.2-0.25-0-3 mm with use of linings and heating and with subsequent rolling of seams by steel rollers".

Report presented at the regular conference of the Moscow city administration NTO Mashprom, April 1963.
(Reported in Avtomaticheskaya Svarka, No. , August 1963, pp 93-95, M. M. Popekhin)

JPRS24,651 19 May 64

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIDSKIY, K.P., akademik.

Roots of transcendental equations. Vest. AN Kazakh. SSR 13 no.6: 79-85 Je '57. (MLRA 10:9)

1. AN KazSSR.
(Functions, Transcendental)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PRESIDSKIY, K.P., akademik.

Geometrization of sets. Vest.AN Kazakh.SSR 13 no.9:91-93 S '57. (MIRA 10:10)

1. AN KazSSR.

(Aggregates)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIDSKIY, K.P., akademik.

Note on the geometry of linear normed spaces. Vest. AN Kazakh. SSR 13 no.4:91-95 Ap '57. (MLRA 10:6)

1. Akademiya nauk Kazakhskoy SSR.
(Geometry, Differential)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERTSOVSKIY, Lazar' Moiseyevich; PRUSAKOV, Mendel' Borisovich; VISLOUKH, L.A., inzhener, redaktor; KHITROV, P.A., tekhnicheskiy redaktor

[Manual for traction substation attendants] Posobie dezhurnomu tiagovoi podstantsii. Moskva, Gos.transp.zhel-dor. izd-vo, 1956. 175 p. (MLRA 10:1)

(Electric railroads--Substations)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIDSKIY, V.Ya., kand.med.nauk

Relation of cocoa butter content on the nature of the effect of drinking cocoa on gastric secretion and evacuation [with summary in English]. Terap.arkh. 31 no.3:42-48 Mr '59.

(MIRA 12:4)

1. Iz gospital'noy terapevticheskoy kliniki (zav. - deystvitel'nyy chlen AMN SSSR prof. V.N. Ivanov) Kiyevskogo meditsinskogo instituta.

(COCOA, effects,
on stomach secretion & evacuation, eff. of cocoa butter content (Rus))
(GASTRIC JUICE,
secretion, eff. of cocoa drinking, relation of cocoa butter content (Rus))
(STOMACH, physiol.
motor funct., eff. of cocoa drinking, relation to cocoa butter content (Rus))

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERUMOVA, N.D., doktor meditsinskikh nauk

Discussion of Professor S.A.Reinberg's article "Criticism of certain basic aspects in radiobiology." Vest.rent. i rad. 31 no.4:51-54 Jl-Ag '56. (MLRA 9:10)

(RADIOTHERAPY
basic radiobiol. aspects)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

FARUNOV, K.I.; KOVALENKO, V.A.: YEFANOV, I.I., retsenzent; PARTSEVSKIY, V.N., redaktor; BERLOV, A.P., tekhnicheskiy redaktor

[Over-all organization of work in drifting crews; the experience of the Tashtagol mine] Kompleksnaia organizatsiia truda v gornoprokhodcheskikh brigadakh; iz opyta Rashtagol'skogo rudnika. Moskva, Gos. nauchno-tekhn. izd-vo lit-ry po chernoi i tsvetnoi metallurgii, 1956. 25 p. (MIRA 10:1)

1. Glavnyy inzhener Glavnogo upravleniya gosudarstvennoy metallurgicheskoy promyshlennosti TSentra i Vostoka Ministerstva chernoi metallurgii (for Yefanov)
(Mining engineering)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANOVA, I.V.

PERSIANOVA, I.V.: "Investigation of the adiabatic compressibility of binary fluid mixtures". Moscow, 1955. Min Higher Education USSR. Moscow Order of Lenin Chemicotechnological Inst imeni D.I. Mendeleyev. (Dissertations for the Degree of Candidate of Chemical Sciences).

SO: Knizhnaya letopis' No 45, 5 November 1955. Moscow.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANINOV, L.S.

Intra-arterial blood transfusion in obstetric and gynecological practice. Akush. gin., Moskva no.4:47-52 July-Aug 1952. (CLML 23:2)

1. Professor. 2. Of the Obstetric-Gynecological Clinic (Director -- Prof. L. S. Persianinov), Minsk Medical Institute.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

[illegible]

Intrear[illegible] transfusi[illegible] [illegible]tric [illegible]
[illegible]
[illegible], [illegible]
[illegible]

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANINOV, L.S.

[A.V.Vishnevskii's method of local anesthesia in obstetrical and gynecological operations] Mestnaia anestezia po A.B. Vishnevskomu pri akusherskikh i ginekologicheskikh operatsiiakh. Moskva, Medgiz, 1955. 97 p. (MLRA 9:1)
(Local anesthesia)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANINOV, L.S., professor

Intra-arterial blood transfusion in acute blood loss in labor. Akush.i gin. no.2:94-102 Mr-Ap '55. (MLRA 8:7)

1. Iz kafedry akusherstva i ginekologii (zav. -prof. L.S.Persianinov) Minskogo meditsinskogo instituta.

(BLOOD TRANSFUSION, in various diseases,
hemorrh. in labor, intra-arterial)
(LABOR, complications,
hemorrh., intra-arterial blood transfusion)
(UTERUS, hemorrhage,
in labor, intra-arterial blood transfusion)
(HEMORRHAGE,
uterus, intra-arterial blood transfusion)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANOV, R.A. (Leningrad); SHCHUKIN, B.K. (Leningrad)

Increasing the rapid action of impulse-selector devices. Izv.AN SSSR Otd.tekh.nauk no.7:17-24 Jl '56. (MLRA 9:9)
(Telemetering)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

1ST AND 2ND ORDERS

PROCESSES AND PROPERTIES INDEX

3RD AND 4TH ORDERS

BC

PERSIANSHEVA, N.

a-3

[illegible] of certain [illegible] and [illegible] in solution. V. [illegible] and N. [illegible] (J. Russ. Phys. Chem. Soc., [illegible]) See this vol., 319.

COMMON ELEMENTS

COMMON VARIABLES INDEX

MATERIALS INDEX

ASM-SLA METALLURGICAL LITERATURE CLASSIFICATION

AUTHOR INDEX

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, N.

Tishchenko, D., Persiantseva, N., Foliadov, V.- "New type of terpene transformations. Part 1. Synthesis of nitriles of the homoterpene series." (p. 1829)

SO: Journal of General Chemistry, (Zhurnal Obshchei Khimii), 1952, Vol. 22, No. 10

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

CA

15

Resin acids of the coniferous trees. I. Abietic acid V. N. Krestinskii, N. I. Persiantseva and A. A. Novak. *J. Applied Chem.* (U. S. S. R.) 12, 1407-20(in French 1420-1)(1939).—Two modifications of abietic acid were distinguished: I, m. 171-3°, $[\alpha]_D$ −80.9°, and II, m. 190-2°, $[\alpha]_D$ −14.58°. Both modifications form monoclinic crystals with the following characteristics: the element of symmetry—double axes L_4, a general form—dihedron (*111*), particular forms—pinacoids (*001*) and (*100*) and dihedrons ($0\bar{1}1$) and (*011*). The crystals of I had the form of triangles and those of II of elongated prisms. The I can be prepd. by the thermal isomerization of the primary acids, by the distn. of rosin *in vacuo* (1-3 mm.) or by the treatment of the primary acid of galipot or rosin (colophony) with HCl, SO_2 or AcOH, followed by recrystn. The II can be prepd. by the distn. *in vacuo* of the oil obtained by the decompn. with H_2SO_4 of the sulfate soap, the waste in the sulfate prepn. of cellulose, or from colophony (previously heated to 300° for 5-6 hrs.) also by distn. *in vacuo*. Methylation of I and II yielded 20-51% esters, yellow and colorless, resp., $b_{0.4}$ 202-208°. I and II did not condense with maleic anhydride at all, but their Me esters yielded a condensation product m. 214-15°. The structural formula of abietic acids with location of the conjugated ethylene bonds in the 2nd ring is not acceptable, but the formula contg. these bonds in the positions 9, 14 and 7, 8 or 5, 13 can be considered, since such a structure can explain the behavior of acids toward maleic anhydride.

A. A. Podgorny

ASM-SLA METALLURGICAL LITERATURE CLASSIFICATION

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIANTSEVA, V.P.; ROZENFEL'D, I.L.

Laboratory method for studying volatile inhibitors. Zav. lab. 24 no. 7:832-836 '58. (MIRA 11:7)
(Corrosion and anticorrosives)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

SOV/129-58-9-10/16

AUTHORS: Persiantseva, V. P., Kudryavtsev, N. T. and Kalb, V.M.

TITLE: Bright Copper Profiled Components Produced from Acidic Electrolytes (Blestyashcheye medneniye profilirovannykh izdeliy iz kislykh elektrolitov)

PERIODICAL: Metallovedeniye i Obrabotka Metallov, 1958, Nr 9, pp 42-46 (USSR)

ABSTRACT: Various authors have pointed out the possibility of producing a bright copper plating from cyanide and acidic solutions (Ref 1). Much attention to this problem is paid in an earlier paper of the team of the authors of this paper (Ref 2) in which it was shown that the most intensive component of an acidic electrolyte bringing about brightness is thiourea. However, if a certain amount of thiourea is present, the precipitates will be very brittle and unserviceable for layer thicknesses exceeding 5μ (Ref 3). V.G. Solokhina, N. T. Kudryavtsev and Lapatukhin, V.S. (Ref 2) proposed an electrolyte containing 250 g/litre $CuSO_4$; 5 H_2O; 50 g/litre H_2SO_4; 0.005 g/litre thiourea and 0.5 g/litre sodium salt, 2.6 (and 2.7) naphthalene diacid which reduces the brittleness of the copper

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deposits resulting from the introduction of thiourea. This electrolyte was tested in the printing industry for copper cylinders intended for deep engraving. A current density of up to 10 A/dm^2 is permissible for depositing copper onto a rotating cylinder; during the process the electrolyte was agitated by means of compressed air. The temperature of the electrolyte should be 15 to 22°C and the produced coatings have a hardness up to 250 kg/mm^2. As a result of investigations, the authors developed a regime of depositing the coatings on profiled components and a technique of purifying the electrolyte from the accumulating decomposition products. The data obtained in laboratory investigations were verified under shop conditions. The laboratory investigations were carried out in electrolytic baths of 1 and 6 litre capacity incorporating a rotating cathode. Low carbon steels with a total surface area of 0.2 dm^2 were coated and as anodes sheets of electrolytic copper were used. Depending on the current density, the duration of the

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

Bright Copper Profiled Components Produced from Acidic Electrolytes SOV/129-58-9-10/16

coating was varied in such a way as to obtain coatings of a thickness of about 20μ. The pilot plant tests were carried out in a bath of 200 litre capacity, maintaining a bath temperature of 12 to 20°C. The concentrations of the sulphuric acid, the sodium salt and the thiourea were respectively 90, 0.5 and 0.005 g/litre; the current density was 3-10 A/dm^2, the temperature 12 to 20°C. As the brightness of the copper deposits decreased, thiourea was added. Naphthalene sulfonic acid was introduced at the rate of 0.04 to 0.06 g per Ah. The data given in the Table, p 43, show that, up to a certain current density, the copper deposits are bright for the electrolytes used in the experiments ($CuSO_4 \cdot 5H_2O$ contents of 146, 200 and 250 g/litre). With increasing concentraion of the copper sulphate, the permissible current density for which bright deposits are obtained will increase somewhat. However, in a solution with an increased concentration of copper sulphate, a rapid saturation of the near anode layer takes place

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SOV/129-58-9-10/16

Bright Copper Profiled Components Produced from Acidic Electrolytes

involving separation of copper sulphate crystals. Therefore, an electrolyte with an average copper sulphate concentration of 200 g/litre was used for which the optimum current density is 326 A/dm^2. Works experiments are also described. The 15 to 20μ thick copper layer is produced in 20 to 25 mins. If the specified technological regime is adhered to, the brightness of the copper coating produced from acidic copper electrolytes containing additions of brightening reagents will be at least as high as for mechanically polished coatings. The brightness of nickel deposited on such a copper layer will be higher than that of nickel deposited on a chemically polished or mechanically polished copper layer. At present in the Krasnyy Oktyabr' Works copper coating is effected only in accordance with this regime and this enabled improving the productivity and reducing the number of copper

Card 4/5

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SOV/129-58-9-10/16

Bright Copper Profiled Components Produced from Acidic Electrolytes

plating baths from 16 to 4.
There are 2 figures, 1 table and 3 references, 2 of which are Soviet, 1 English.

ASSOCIATION: Moskovskiy khimiko-tekhnologicheskiy institut imeni Mendeleyeva (Moscow Chemical-Technological Institute imeni Mendeleyev)

1. Copper plating--Test results 2. Electrolytes--Properties

Card 5/5

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERS[illegible], [illegible] P.

PERS[illegible]TSEVA, [illegible] P. -- [illegible] Processes [illegible]
[illegible]
[illegible] Technical [illegible]

[illegible]: [illegible], [illegible]

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

S/076/61/035/007/006/019
B127/B102

AUTHORS: Rozenfel'd, I. L., Polteva, M. N., and Persiantseva, V. P.

TITLE: Physicochemical properties of the volatile corrosion inhibitor dicyclohexylamine nitrite

PERIODICAL: Zhurnal fizicheskoy khimii, v. 35, no. 7, 1961, 1474-1477

TEXT: The authors studied the physicochemical properties of the widely used dicyclohexylamine nitrite with the empirical formula ($C_{12}H_{24}N_2O_2$) and the lattice constants a = 8.16 ± 0.04Å, b = 8.56 ± 0.04Å, c = 19.82 ± 0.04Å; $P2_12_12_1$, z = 4. In order to determine the vapor pressure of the compound, the effusion method by Knudsen was applied. Table II presents the vapor pressures obtained for different degrees of purity. The heat of sublimation was calculated from the Clausius-Clapeyron equation: σ = 4.575 A = 25.3 kcal/mole. For further clarification, the dipole moment of this compound in dilute benzene solution was measured to be ~4D. It is assumed therefrom that not the salt form exists in dilute solution but a molecular compound of

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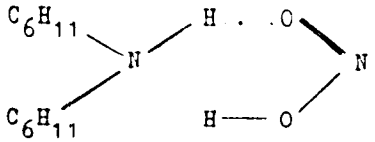
the following structure:



The solid phase possibly consists of ions forming molecular complexes on evaporation. The high value of the heat of sublimation (~ 25 kcal/mole) suggests that a break of the crystal lattice destroys the weak Van der Waals forces and is also accompanied by a displacement of groups. Discrepancies between the values obtained and values from earlier papers are due to more accurate investigation methods at low temperatures: 35.2-65.8°C and to the high purity of the dicyclohexylamine nitrate (10 crystallizations from alcohol). Ya. K. Syrkin, M. Ye. Dyatkina and Ye. A. Shott-L'vova assisted in the measurements. There are 1 figure, 2 tables, and 9 references: 5 Soviet and 4 non-Soviet. The two references to English-language publications read as follows: Ref. 3: A. Wachter, Mod. Packaging, 22, 147, 1948; Ref. 8: A. Wachter, et. al.: Corrosion, 7, 284, 1951.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

Physicochemical properties ... S/076/61/035/007/006/019
B127/B102

ASSOCIATION: Akademiya nauk SSSR Institut fizicheskoy khimii (Academy of Sciences, USSR Physicochemical Institute)

SUBMITTED: October 12, 1959

Table 2: Vapor pressure of crystalline dicyclohexylamine nitrite in 10^4 mm Hg.
Legend: 1) After 3 recrystallizations; 2) after 7 recrystallizations; 3) after 8 recrystallizations; 4) after 10 recrystallizations.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, DAVORIN

Determination of alkaline phosphatase in serum. Marijana Fišer Herman and Davorin Persić. *Acta Pharm. Jugosl.* 4, 62-72(1954)(German summary).—A no. of detns. of alk. phosphatase in serum were done by the method by King (*C.A.* 28, 7285[3]) and then by the method by Shinowara (*C.A.* 35, 2581[4]). It is concluded that the K. method has many advantages over the method of S. 23 references. V. Mihajlov

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

FISER-HERMAN, Marijana; PERSIC, Davorin

The determination of alkaline phosphatase in the blood. Acta pharm. jugosl. 4 no.2:62-72 1954.

1. Zavod za klinicku kemiju, Farmaceutski fakultet Zagreb; primljeno 3.IV.1954.

(PHOSPHATASES
alkaline, in blood, determ.)
(BLOOD
phosphatases, alkaline, determ.)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, Ivan

Cardiac arrest in surgery. Zdrav. vest., Ljubljana 24 no. 9-10:347-349 1955.

1. Kirurgicni Oddelek splosne bolnisnice v vipavi - sef primarij dr. I. Persic.

(CARDIAC ARREST,
in surg., resuscitation with epinephrine (Sl))
(EPINEPHRINE, ther. use
resuscitation in cardiac arrest during surg. (Sl))

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, IVAN

PERSIC, Ivan

Surgical diseases in children and their mortality. Zdrav. vest., Ljubljana 23 no.9-10:223-230 1954.

1. Kirurgicni oddelek Splosne bolnisnice v Vipavi, sef dr. Ivan Persic.

(SURGERY, OPERATIVE, in inf. & child mortal.)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC,N.; HRCKO,N.

Simultaneous application of electric shock and insulin in the treatment of mental disorders. Neuropsihijatrija 7 no.1-2:64-73 '59.

1. Iz Neurolosko-psihijatrijske klinike Med. fakulteta u Zagrebu, Predstojnik:prof. dr. R. Lopasic.
(SHOCK THERAPY ELECTRIC)
(SHOCK THERAPY INSULIN)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, N.; ROGINA, V.; BOZOVIC, M.

Effect of largactil on insulin hypoglycemia. Neuropsihijatrija 3 no.2:92-99 1955.

1. Neurolosko-psihijatrijska klinika Medicinskog fakulteta. Pred: Prof. dr. R. Lopasic). Bolnica za zivcane i dusevne bolesti u Vrapcu (Direktor: Dr. J. Glazer) i Zavod za patofiziologiju Medicinskog fakulteta. (Pred: Dr. P. Sokolik).

(INSULIN, effects,
hypoglycemia in dogs, eff. of chloropromazine on)
(CHLOROPROMAZINE, effects,
on exper. hypoglycemia induced with insulin in dogs)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, N.; ROGINA, V.; TURNER, V.

Prolonged insulin coma. Neuropsihijatrija 3 no.3-4:193-215 1955.

1. Neurolosko-psihijatrijska klinika Medicinskog fakulteta u Zagrebu. Pred.: Prof. dr. R. Lopasis. Bolnica za zivcane i dusevne, bolesti u Vrapcu, Direktor: Prim. dr. J. Glaser. Institut za medicinska istrazivanja u Zagrebu (Pred: Prof. dr. V. Vouk).

(HYPERINSULINISM, experimental, histopathol. & ther. (Ser))

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

TOMASEGOVIC, Z.; JANKOVIC, Z.; PETKOVIC, V.; STANIC, M.; BETLHEIM, S.; BLAZEVIC, D.; PERSIC, N.; ZORINC, S.; TEODOROVIC, B.; VRANCIC, J.; VODOPIJA, I.; ANTONIAZZO, Z.; CULIC, R.; GALINOVIC-WEIBGLASS, M.; [redacted] MRAVUNAC, B.; KOEHLER-KUBELKA, N.; CEZNER, M.; KOHN, V.; TEKAVCIC, B.; EMILI, H.; SMERDEL, S.; SOOS, E.; VUKSANOVIC, V.; JANJATOVIC, M.; [redacted] GRUENWALD, P.; SKRABALO, Z.; CREPINKO, I.; HAUPTMANN, E.; VIDACEK, S.; HORVAT, A.; MIOCKA, O.; IVANCEVIC, D.; PERGER, A.; KRSNJAVI, B.; PRAZIC, M.; SALAJ, B.; SUROTIC, R.; RADOSEVIC, Z.; KELER-BACOKA, M.; HAHN, A.; MATKOVIC, B.; RADONIC, M.

Reveiw of periodicals; medicine. Bul sc Youg 9 no.4/5:145-147 Ag-O '64.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, Nikola, dr.; BETLHEIM, Stjepan, dr.; BLAZEVIC, Duska, dr.; BECK-DVORZAK, Maja, dr.; BUCAN, Neda, dr.; CIVIDINI, Eugenija, dr.; RADOSEVIC, Zlata, prof.

Attitude of the milieu toward the mentally ill. Lijecn. vjesn. 87 no.4:385-395 Ap '65.

1. Iz Neurolosko-psihijatrijske klinike Medicinskog fakulteta u Zagrebu.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

PERSIC, N.; GRCEVIC, N.

Mental disorders as dominant symptoms in brain tumor. Neuropsihijatrija 2 no.1-2:72-87 1954.

1. Iz Neuropsihijatrijske klinike Medicinskog fakulteta u Zagrebu (Predstojnik: prof. dr R.Lopasic)

(BRAIN, neoplasms, causing ment.disord.)

(MENTAL DISORDERS, etiology and pathogenesis, brain tumor)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8

[illegible]

PERSIC, Dr Nikola, Neurological-Psychiatric Clinic (Neurolosko-Psihijatrijska Klinika), Faculty of Medicine (Medicinski Fakultet), Zagreb.

"Psychiatry in the Soviet Union."

Zagreb, Lijecnicki Vjesnik, Vol 85, No 7, July 196[illegible], pp 74[illegible]-75[illegible].

Abstract: [Author's English summary modified] The author deals with the history of psychiatry in the USSR (with particular reference to I.P. PAVLOV), the status and organizational principles of Soviet psychiatric services (open psychiatric departments at psychiatric colonies for purposes of rehabilitation, organizational sections at psychiatric hospitals, psychiatric sections in general hospitals, psychiatric sections in psychiatric hospitals to treat mental disturbances in children, psychiatric dispensaries in the more sparsely inhabited areas, evening and day psychiatric hospitals, numerous psychiatric institutes), and medical studies and the system of specialized training (five to six months in a large psychiatric hospital on the basis of a specific program, three years of work in such a hospital, and five months of institutional specialization).

Seven statistical tables on institutions, staff, and patients. Six Soviet references.

1/1

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001240120012-8"

PERSIC, Nikola, dr.; [illegible], dr.; [illegible], Sergije, dr.; MALETIC, Aleksandar, dr.

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