"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

SOV/78-3-11-4/23

The Determination of the Solubility Products of Silver Citrate at Different Ionic Strength of the Solution

The thermodynamic value of the solubility product of silver citrate was determined for solutions with an ionic strength of 0,002 - 0,103. In the case of an ionic strength of 0,002 of the solution the solubility of $Ag_3C_6H_5O_7$ amounts to $3,4.10^{-4}$g-mol/l and the solubility product $k^c = (3,3 \pm 0,1).10^{-13}$.
In the case of an ionic strength of 0,103 of the solution the solubility amounts to $5,5.10^{-4}$g-mol/l and $k^c = (2,4 \pm 0,3).10^{-12}$.
These results show that the solubility product changes by almost the tenfold with the change of the ionic strength of the solution of 0,002 - 0,103.
Figure 1 shows the dependence of the negative logarithm of the solubility products of silver nitrate (pK) on the ionic strength of the solution.
There are 1 figure, 2 tables, and 4 references, 1 of which is Soviet.

SUBMITTED: April 10, 1957

Card 2/2

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

MORACHEVSKAYA, M.D.; ZLOBIN, V.S.; PTITSYN, B.V.

Adsorption of strontium on hydroxylapatite crystals [with summary in English]. Biokhimiia 23 no.4:564-567 Jl-Ag '58. (MIRA 12:3)

1. The Technological Institute of Food Industry, High School Ministry of the U.S.S.R., Leningrad.
(GASTRIC JUICE,
radiostrontium, adsorption with hydroxyl-apatite crystals (Rus))
(STRONTIUM, radioactive,
in gastric juice, adsorption with hydroxyl-apatite crystals (Rus))

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

DASHKEVICH, Leonid Borisovich, kand.khim.nauk; PTITSYN, B.V., prof., nauchnyy red.; VOROB'YEV, G.S., red.izd-va; GURDZHIYEVA, A.M., tekhn.red.

[Progress in radiochemistry] Uspekhi radiatsionnoi khimii. Leningrad. Ob-vo po rasprostraneniiu polit. i nauchn.znanii RSFSR, Leningr.otd-nie, 1959. 31 p. (MIRA 12:9)
(Radiochemistry)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

SOV/78-4-2-20/40

5(2)

AUTHORS: Sheronov, L. N., Ptitsyn, B. V.

TITLE: On a Citrate Complex of Zirconium (O kompleksnom tsitrate tsirkoniya)

PERIODICAL: Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 2, pp 367-371 (USSR)

ABSTRACT: Upon the interaction of 1 mole zirconium carbonate with 1.5 moles citric acid, which is semi-saturated with KOH, a complex potassium zirconyl citrate of the composition $K[ZrO(C_6H_5O_7)].2.5H_2O$ is formed. The complex $K_3C_6H_5O_7.ZrC_6H_4O_7.9.5H_2O$ described by Mandl (Ref 1) is not formed under these production conditions. In order to determine the coordination formula of the compound obtained the molecular electric conductivity and the molecular weight of this compound were determined. The molecular weight of potassium zirconyl citrate in aqueous solution shows values from 175-195. The molecular weight and the molecular electric conductivity show that, in aqueous solution, the compound dissociates into two ions:

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SOV/78-4-2-20/40

On a Citrate Complex of Zirconium

$$K\left[ZrO(C_6H_5O_7)\right] \rightleftharpoons K^+ + \left[ZrO(C_6H_5O_7)\right]^-$$

The aqueous solution of the compound is weakly acid. For the potassium zirconyl citrate complex produced by Mandl the following formula was suggested: $K\left[ZrO(C_6H_5O_7)\right] \cdot K_2HC_6H_5O_7 \cdot 8.5H_2O$

and the following coordination formula:

$$K_3 \left[ZrO \begin{matrix} C_6H_5O_7 \\ HC_6H_5O_7 \end{matrix}\right] 8.5H_2O$$

There are 2 tables and 6 references, 2 of which are Soviet.

SUBMITTED: December 12, 1957

Card 2/2

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

5(4),21(1) SOV/78-4-4-10/44

AUTHORS: Tekster, Ye. N., Vinogradova, L. I., Ptitsyn, B. V.

TITLE: The Determination of the Stability Constants of the Complex Oxalates of Magnesium and Uranyl Using an Oxalate-silver Electrode (Opredeleniye konstant nestoykosti kompleksnykh oksalatov magniya i uranila s pomoshch'yu oksalatno-serebryanogo elektroda)

PERIODICAL: Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 4, pp 764-768 (USSR)

ABSTRACT: The stability constants of the complexes $K_2[Mg(C_2O_4)_2]$ and $K_6[(UO_2)_2(C_2O_4)_5]$ were determined using an oxalate-silver electrode. Solutions of various concentrations of both complexes were saturated with silver oxalate at 25°, and the potential of the oxalate-silver electrode was measured in these solutions in order to determine the equilibrium activity of the $C_2O_4^{2-}$ ion. The results of these measurements are given in a table. The integral stability constant for the magnesium complex $K_2[Mg(C_2O_4)_2]$ was calculated:

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$K_2[Mg(C_2O_4)_2]$

I (ion strength)	$K_{integral}$
$0.89 \cdot 10^{-1}$	$5.7 \cdot 10^{-5}$
$0.49 \cdot 10^{-1}$	$2.9 \cdot 10^{-5}$
$0.31 \cdot 10^{-1}$	$2.9 \cdot 10^{-5}$

The measured results required for the calculations are summarized in a table. The dissociation of the complex ion $[(UO_2)_2(C_2O_4)_5]^{6-}$ occurs according to the following scheme:

$$[(UO_2)_2(C_2O_4)_5]^{6-} \rightleftharpoons C_2O_4^{2-} + 2[UO_2(C_2O_4)_2]^{2-} \quad (1)$$

$$2[UO_2(C_2O_4)_2]^{2-} \rightleftharpoons C_2O_4^{2-} + [(UO_2)_2(C_2O_4)_3]^{2-} \quad (2)$$

$$[(UO_2)_2(C_2O_4)_3]^{2-} \rightleftharpoons C_2O_4^{2-} + 2UO_2C_2O_4 \quad (3)$$

$$2UO_2C_2O_4 \rightleftharpoons 2UO_2^{2+} + 2C_2O_4^{2-} \quad (5)$$

together: $[(UO_2)_2(C_2O_4)_5]^{6-} \rightleftharpoons 5C_2O_4^{2-} + 2UO_2^{2+}$ (6)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

The Determination of the Stability Constants of the Complex Oxalates of Magnesium and Uranyl Using an Oxalate-silver Electrode SOV/78-4-4-10/44

It is assumed that the complex ion $[(UO_2)_2(C_2O_4)_4]^{4-}$ exists in the solution. The stability constants $K_1 \cdot K_2$ and K_2 for $K_6[(UO_2)_2(C_2O_4)_5]$ were calculated as follows:

$K_6[(UO_2)_2(C_2O_4)_5]$

I (ion strength)	$K_1 \cdot K_2$	K_2
$0.69 \cdot 10^{-1}$	$3.8 \cdot 10^{-5}$	"
$0.22 \cdot 10^{-1}$	"	$4.8 \cdot 10^{-2}$
$0.08 \cdot 10^{-1}$	"	3.0

The data required for the calculations are given in a table. A further table gives the results of the calculation of K_2. There are 4 tables and 7 references, 3 of which are Soviet.

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APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

The Determination of the Stability Constants of the Complex Oxalates of Magnesium and Uranyl Using an Oxalate-silver Electrode SOV/78-4-4-10/44

ASSOCIATION: Kafedra obshchey i analiticheskoy khimii Leningradskogo tekhnologicheskogo instituta pishchevoy promyshlennosti (Chair of General and Analytical Chemistry of the Leningrad Technological Institute of the Foodstuffs Industry) and Kafedra tekhnologii iskusstvennykh radioaktivnykh veshchestv Leningradskogo tekhnologicheskogo instituta im. Lensoveta (Chair of the Technology of Artificial Radioactive Materials of the Leningrad Technological Institute imeni Lensovet)

SUBMITTED: December 30, 1957

Card 4/4

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

SOV/78-4-10-13/40

5(2)

AUTHORS: Ptitsyn, B. V., Tekster, Ye. N.

TITLE: Determination of the Instability Constant of Oxalate Complexes of Uranyl by Means of the Method of the Shifted Equilibrium

PERIODICAL: Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 10, pp 2248-2254 (USSR)

ABSTRACT: On description of the solubility method used in the determination of instability constants it is disregarded that in several cases a displacement of the dissociation equilibrium of the complex ion may take place brought about by reaction of a ion situated in the inner sphere of the complex with a ion in the outer sphere. In such cases the equilibrium can be approached from two sides and the thermodynamic equilibration be proved. In the following the equilibrium of the reaction

$$\left[UO_2(C_2O_4)_y\right]^{2-2y} + 2xAg^+ \rightleftharpoons \left[UO_2(C_2O_4)_{y-x}\right]^{2-2(x-y)} + xAg_2C_2O_4$$

is investigated. The approach to the equilibrium from the left of the equation is denoted as first modification, the approach from the right as second modification of the method of the shifted equilibrium. The equilibria of the following systems

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SOV/78-4-10-13/40

Determination of the Instability Constant of Oxalate Complexes of Uranyl by Means of the Method of the Shifted Equilibrium

were investigated: $K_2\left[UO_2(C_2O_4)_2\right] + AgNO_3$ at an initial molar ratio between silver and uranyl ion of 2 and 4; and $K_6\left[(UO_2)_2(C_2O_4)_5\right] + AgNO_3$ at an ionic ratio of 1, 3, and 5. The equilibrium distribution of silver ions between the solid and liquid phase at 25° was determined. The results are illustrated in the coordinate system $p\left[Ag^+\right]$, $f(pc_{UO_2^{2+}})$ in figure 1.

Tables 1 and 2 give the computed values of the instability constants K_I and K_{II}. Figures 2 and 3 show the dependence of the constants on $\sqrt{J}$ (J = ionic strength). For K_I $8.3.10^{-6}$ was found by means of the first modification, $3.6.10^{-5}$ by means of the second modification; for K_{II} by means of both modifications $1.0.10^{-6}$ in accordance. The value of K_I obtained by means of the second modification is considered the more likely one. The reasons for it shall be given in a later paper. There are 3 figures, 2 tables, and 5 references, 3 of which are Soviet.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

SOV/78-4-10-13/40

Determination of the Instability Constant of Oxalate Complexes of Uranyl by Means of the Method of the Shifted Equilibrium

ASSOCIATION: Leningradskiy tekhnologicheskiy institut im. Lensoveta (Leningrad Technological Institute imeni Lensovet) Tekhnologicheskiy institut pishchevoy promyshlennosti (Technological Institute of Food Industry)

SUBMITTED: July 11, 1958

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APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

MORACHEVSKAYA, M.D.; PTITSYN, B.V.

Determination of iodides in the presence of bromides and chlorides by means of radioactive iodine. Zav.lab. 26 no.3:269-271 '60. (MIRA 13:6)

1. Leningradskiy tekhnologicheskiy institut pishchevoy promyshlennosti.
(Iodides) (Iodine--Isotopes)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

S/079/60/030/007/021/039/XX
B001/B066

AUTHORS: Alekseyeva, I. P., Pesotskaya, V. M., and Ptitsyn, B. V.

TITLE: Oxidation Potential of Permanganate

PERIODICAL: Zhurnal obshchey khimii, 1960, Vol. 30, No. 7, pp. 2104-2108

TEXT: The purpose of this work was a more exact characterization of the oxidizing effect of permanganate. Its oxidizing effect depends on the acidity of the solution, but research workers disagree in this respect. The investigation of the dependence of the potential of the permanganate solution on pH is of interest since the reduced form which is in equilibrium with the permanganate ion may be determined from the potential as a function of the percentual content. In an acid solution, the latter may be in equilibrium with some of its reduction products, except compounds of bivalent manganese, as potassium permanganate oxidizes the $Mn^{\cdot\cdot}$ at any percentual content. Therefore, the universal equation

$$E = 1510 + \frac{59.1}{5} \lg \frac{[MnO_4^{\prime}]\,[H^{\circ}]^8}{[Mn^{\cdot\cdot}]}$$

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

cannot be used to characterize the oxidation properties of permanganate. To study these states of equilibrium different methods were applied (Refs. 3-5,6,7). B. V. Ptitsyn and V. F. Petrov (Ref. 8) indicated that also in the case of bichromate solution the oxidation potential may be affected by the nature of the acid. This might also apply to permanganate. The present paper deals with the following problems: 1) Quantitative determination of the oxidation potential of the permanganate solution as dependent on the pH of the medium. 2) The question as to whether the nature of the acid influences the potential. 3) A more exact characterization of the states of equilibrium established in the permanganate solution under the conditions of the experiments to be performed. The authors investigated the dependence of a potassium permanganate solution (0.01 mole) in solutions of $HClO_4$, HNO_3, H_2SO_4, CH_3COOH. The effect depends on the nature of the acid at low pH only. An attempt was made to interpret the nature of the reduced forms which are in equilibrium with the permanganate ion, on the basis of the functions of the permanganate solution potential on pH in solutions of different acids. Table 1 gives the data of the potential dependence of $KMnO_4$ dissolved in $HClO_3$, and Table 2 the values

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

Oxidation Potential of Permanganate

S/079/60/030/007/021/039/XX
B001/B066

$\Delta E/\Delta pH$ for various redox systems. The diagram illustrates the oxidation potential of potassium permanganate as dependent on pH in the presence of different acids. For the oxidation potential of permanganate in $HClO_4$ and HNO_3, E was found to be 1600 mv, in H_2SO_4: E = 1650 mv. There are 1 figure, 2 tables, and 9 references: 4 Soviet, 4 US, and 1 Spanish.

ASSOCIATION: Leningradskiy tekhnologicheskiy institut pishchevoy promyshlennosti (Leningrad Technological Institute of the Alimentary Industry)

SUBMITTED: July 6, 1959

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APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.; SHEKONOV, L.N.

Complex niobium oxalate. Izv. Sib. otd. AN SSSR no.9:44-46 '61. (MIRA 14:10)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR, Novosibirsk.

(Niobium compounds)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.; SHERONOV, L.N.

Complex zirconium oxalate. Izv. Sib. otd. AN SSSR no.10:80-83 '61. (MIRA 14:12)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR, Novosibirsk.
(Zirconium oxalates)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.; VINOGRADOVA, L.I.; VASIL'YEVA, L.L.; Prinimala uchastiye: LUKINYKH, N.L.

Use of a silver citrate electrode for the determination of instability constants of complex citrates. Zhur.neorg.khim. 7 no.5:1009-1011 My '62. (MIRA 15:7)
(Citrates) (Silver compounds) (Electromotive force)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

ALTYNOV, V.O.; PTITSYN, B.V.

Theory of the silver chloride electrode and the determination of complex chloride instability constants. Zhur.neorg.khim. 7 no.9: 2103-2109 S '62. (MIRA 15:9)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.
(Electrodes) (Silver chloride) (Chlorides)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

NOVOSELOV, R.I.; MUZYKANTOVA, Z.A.; PTITSYN, B.V.

Oxidation potential as a characteristic of the reaction velocity in the kinetic methods of analysis. Zhur.neorg.khim. 8 no.1:135-140 Ja '63. (MIRA 16:5)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.
(Oxidation) (Chemical reaction, Rate of)
(Chemistry, Analytical)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

ALTYNOV, V.I.; PTITSYN, B.V.

Theory of a lead chloride electrode. Zhur. neorg. khim. 9 no.10: 2407-2410 O '64. (MIRA 17:12)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

NIVOSELOV, R.I.; ZEMSKOV, S.V.; PTITSYN, B.V.

Oxidation of tetrammine platinum chloride by iodine. Dokl. AN SSSR 158 no.5:1133-1135 O '64. (MIRA 17:10)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR. 2. Chlen-korrespondent AN SSSR (for Ptitsyn).

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

ZEMSKOV, S.V.; PTITSYN, B.V. [deceased]

Oxidation of cis- and trans-isomers of bivalent platinum by sodium vanadate. Zhur. neorg. khim. 10 no.6:1502-1503 Je '65. (MIRA 18:6)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

NIKOLAYEVA, N.M.; PTITSYN, B.V. [deceased]; PASTUKHOVA, Ye.D.

Hydrolysis of potassium chloroplatinate. Zhur. neorg. khim. 10 no.5:1058-1061 My '65. (MIRA 18:6)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

NIKOLAYEVA, N.M.; PETTSEN, B.V. [deceased]; GORBACHEVA, I.I.

Hydrolysis of potassium chloroplatinite. Zhur. neorg. khim. 10 no.5:1051-1057 My '65. (MIRA 18:6)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V. [deceased]; SHERONOV, L.N.

Certain number of niobium complex compounds of relative stability. Izv. SO AN SSSR no.3: Ser. khim. nauk no.1:68-71 '65. (MIRA 18:8)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR, Novosibirsk.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

TEMSKOV, S.V.; PTITSYN, B.V.

Oxidation of the platinite series by sodium vanadate. Dokl. AN SSSR 160 no.2:343-345 Ja '65. (MIRA 18:2)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.
2. Chlen-korrespondent AN SSSR (for Ptitsyn).

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

BELYAYEV, A.V.; KAZAKOV, V.P.; PTITSYN, B.V.

Certain features of the behavior of complex compounds of Rh (III) in solution as linked with the compensation effect. Dokl. AN SSSR 160 no.1:149-150 Ja '65. (MIRA 18:2)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.
2. Chlen-korrespondent AN SSSR (for Ptitsyn).

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

NOVOSELOV, R.I.; MUZYKANTOVA, Z.A.; PTITSYN, B.V.

Complexes unstable with time. Zhur. neorg. khim. 9 no.11:2590-2593 N '64 (MIRA 18:1)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PESHCHEVITSKIY, B.I.; PTITSYN, B.V.; LESKOVA, N.M.

Hydrolysis of chloroplatinite ion. Izv. Sib. otd. AN SSSR no. 11: 143-145 '62. (MIRA 17:9)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR, Novosibirsk.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

ZVYAGINTSEV, O.Ye., red.; PTITSYN, B.V., red.; SHUL'MAN, V.M., red.; PESHCHEVITSKIY, B.I., red.; ZAYTSEVA, I.P., red.; OVCHINNIKOVA, T.K., tekhn. red.

[Problems in the analysis of noble metals; proceedings of the Fifth All-Union Conference on the Analysis of Noble Metals] Voprosy analiza blagorodnykh metallov; trudy Vsesoiuznogo soveshchaniia po analizu blagorodnykh metallov. 5th. Novosibirsk, Izd-vo Sibirskogo otd-niia AN SSSR, 1963. 100 p. (MIRA 17:4)

1. Vsesoyuznoye soveshchaniye po analizu blagorodnykh metallov. 5th.

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.; SHERONOV, L.N.

Method for determining the relative stability of complex compounds. Izv. SO AN SSSR no.3 Ser. khim. nauk no.1:3-8 '63. (MIRA 16:8)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR, Novosibirsk.

(Complex compounds)

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BELYAYEV, A.V.; PTITSYN, B.V. [deceased]

Effect of bivalent mercury on K_3RhCl_6 aquation reaction.
Zhur.ob.khim. 35 no.12:2124-2127 D '65.
(MIRA 19:1)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR. Submitted November 10, 1964.

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PTITSYN, B.V. [deceased]; VINOGRADOVA, L.I.; MAKSIMYUK, Ye.A.

Use of silver oxalate electrode for determining the instability constants of an iron oxalate complex. Zhur.neorg.khim. 10 no.8: 1929-1930 Ag '65. (MIRA 19:1)

1. 1-y Leningradskiy meditsinskiy institut imeni I.P.Pavlova, kafedra neorganicheskoy khimii, i Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR.

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CHERNYAYEV, I.I.; ZEMSKOV, S.V.; PTITSYN, B.V. [deceased]

Oxidation-reduction properties of nitrite complexes of platinum. Zhur.neorg.khim. 10 no.11:2404-2410 N '65. (MIRA 18:12)

1. Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR. Submitted May 5, 1964.

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PTITSYN, B.V. [deceased]; VINOGRADOVA, L.I.; MAKSIMYUK, Ye.A.

Oxidation of Cr^{3+} and Fe^{3+} complex oxalates by potassium permanganates. Zhur.neorg.khim. 10 no.11:2493-2495 N '65. (MIRA 18:12)

1. Submitted April 11, 1964.

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PTITSYN, B.V. [deceased]; VINOGRADOVA, L.I.; MAKSIMYUK, Ye.A.

Potentiometric titration of complex ions with ammonium vanadate. Zhur.neorg.khim. 10 no.11:2496-2498 N '65. (MIRA 18:12)

1. Kafedra neorganicheskoy khimii I Leningradskogo meditsinskogo instituta imeni I.P.Pavlova i Institut neorganicheskoy khimii Sibirskogo otdeleniya AN SSSR. Submitted April 11, 1964.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, D., inzh.; MAL'TSEV, V., inzh.

The TO-1 automatically controlled continuous maintenance line. Avt.transp. 38 no.7:24-25 Jl '60. (MIRA 13:7)
(Motor vehicles--Maintenance and repair)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, G.A.

PHASE I BOOK EXPLOITATION 405

Ptitsyn, Gennadiy Anatol'yevich and Kokichev, Valentin Nikolayevich

Zuboreznyye stanki; spravochnoye posobiye (Gear-cutting Machines; a Handbook) 2d ed., enl. and rev. Moscow, Mashgiz, 1957. 448 p. 22,000 copies printed.

Reviewer: Fedotenok, A.A., Candidate of Technical Sciences, Docent; Ed.: Shavlyuga, N.I., Candidate of Technical Sciences, Docent; Ed. of Publishing House: Borodulina, I.A. Tech. Ed.: Speranskaya, O.V.; Managing Ed. of Mashgiz, Leningrad Branch: Bol'shakov, S.A.

PURPOSE: The book is intended for technologists working in the field of gear-cutting, as well as for foremen and set-up men for gear-cutting machines. The book is recommended as a textbook for technical schools.

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Gear-cutting Machines; a Handbook 405

COVERAGE: This textbook describes methods and formulas for setting-up of gear-cutting machines. Dimensions necessary for the design of accessories, standard designs of mounting attachments and mandrels, as well as basic data for the selection of gear-cutting tools are given. The publication of the first edition of this textbook on gear-cutting tools indicated that such a book is indispensable to technologists and foremen employed in the field of gear-cutting as well as to designers planning jigs and fixtures of gear-cutting machine tools. This book found wide application in plants and planning organizations. In revising the material for the second edition, the authors preserved its basic contents and character of presentation. Only outdated material was excluded and the remaining material was reviewed and corrected. Separate chapters were enlarged to include: data on new models of foreign and domestic gear-cutting machines; extent of technological precision of machine tools for cutting spiral bevel gears using the "single cutter"method; data on adjusting machine tools for gear hobbing; recommendations on the

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selection of cutting speeds; requirements which precision gear-cutting machines should incorporate in accordance with latest data on foreign techniques, etc. There are 18 Soviet references. No personalities are mentioned.

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PTITSYN, Gennadiy Anatol'yevich; KOKICHEV, Valentin Nikolayevich; FEDOTENOK, A.A., kand.tekhn.nauk, dotsent, retsenzent; SHAVLYUGA, N.I., kand. tekhn.nauk, dotsent, red.; BORODULINA, I.A., red.izdatel'stva; SPERANSKAYA, O.V., tekhn.red.

[Gear-cutting machines; a handbook] Zuboreznye stanki; spravochnoe posobie. Izd.2-oe, dop.i perer. Moskva, Gos.nauchno-tekhn.izd-vo mashinostroit.lit-ry, 1957. 448 p. (MIRA 11:1)
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PTITSYN, Gennadiy Antonovich; KOKICHEV, Valentin Nikolayevich; IVANOV, A.F., nauchnyy red.; SOKOLOV, K.M., inzh., retsenzent; KLIMINA, Ye.V., red. izd-va; KOROVENKO, Yu.N., tekhn. red.

[Calculation and reconditioning of used gears] Raschet i izgotovlenie zubchatykh peredach v remontnom dele; spravochnoe posobie. Leningrad, Gos. soiuznoe izd-vo sudostroit. promyshl., 1961. 518 p. (MIRA 14:8)

(Gearing—Maintenance and repair)

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PTITSYN, G.V.

PTITSYN, G.V.

It was a great event. Elek. i tepl. tiaga no.11:21 N '57. (MLRA 10:11)

1. Starshiy inzh. zavoda "Dinamo."
(Locomotives)

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CHAYKOVSKIY, E.F.; PTITSYN, G.V.

Use of the effect of positive surface ionization in studying the cathode sputtering of metals. Zhur.tekh.fiz. 34 no.12:2194-2196 D '64. (MIRA 18:2)

1. Vsesoyuznyy nauchno-issledovatel'skiy institut minokristallov, Khar'kov.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

CHAYKOVSKIY, E.F.; PTITSYN, G.V.

Positive surface ionization of potassium on textured platinum bands. Zhur. tekh. fiz. 35 no.3:528-534 Mr '65. (MIRA 18:6)

1. Vsesoyuznyy nauchno-issledovatel'skiy institut monokristallov.

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CHAYKOVSKIY, E.F.; PYATIGORSKIY, G.M.; PTITSYN, G.V.

Temperature hysteresis of positive surface ionization and the work function of a homogeneous emitter. Zhur. tekh. fiz. 35 no.6:1132-1138 Je '65. (MIRA 18:7)

1. Vsesoyuznyy nauchno-issledovatel'skiy institut monokristallov, Khar'kov.

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CHAYKOVSKIY, E.F.; PTITSYN, G.V.

Positive surface ionization of Na on textured platinum foils. Zhur. tekh. fiz. 35 no.6:1158-1159 Je '65. (MIRA 18:7)

1. Vsesoyuznyy nauchno-issledovatel'skiy institut monokristallov, Khar'kov.

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AUTHOR: Chaykovskiy, E.F.; Pyatigorskiy, G.M.; Ptitsyn, G.V.

TITLE: Thresholds and optimum conditions for positive surface ionization

SOURCE: Zhurnal tekhnicheskoy fiziki, v. 35, no. 8, 1965, 1493-1500

TOPIC TAGS: surface ionization, heat adsorption, positive ionization, emission threshold, alkali metal, tungsten, nickel, platinum, molybdenum

ABSTRACT: The theory of surface ionization is discussed and a method is devised for deriving the heat of adsorption of the ion from the envelope of the family of surface ionization curves (ion current versus surface temperature for constant incident atom flux) obtained by varying the incident atom flux. This method is applicable regardless of whether the ionization potential is greater or less than the work function. It is suggested that the point at which the ionization curve is tangent to the envelope be regarded as the threshold point. Experimental surface ionization curves were obtained for K on W, Ni, and Pt, for Rb on W and Ni, for Cs on Mo, and for Na on Pt with an apparatus that has been described elsewhere by two of the authors (ZhTF 35, 1158, 1965). The emitter and collector were oriented polycrystalline strips separated by approximately 0.35 mm. The W and Mo surfaces

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consisted mainly of (100) faces, and the Ni and Pt surfaces of (111) faces. The alkali metal vapor pressure was varied from 10^{-6} to 10^{-3} mm Hg; this caused a shift of 350 to 450° in the threshold temperature. The experimental data were processed by the method developed in the first part of the paper, and the following values for the heat of adsorption of the ions were found: for K on W, Ni, and Pt - 2.3, 2.4, and 2.6 eV, respectively; for Rb on W and Ni - 1.8 and 2.1 eV; for Cs on Mo - 1.9 eV; and for Na on Pt - 3.1 eV. The probable error in each case was 0.1 eV. These findings are in satisfactory agreement with data in the literature, although the latter were obtained with unoriented polycrystalline emitters. It is suggested that this agreement may be due to the predominance of the electrostatic image force in determining the heat of adsorption. "In conclusion, the authors consider it their pleasant duty to express their gratitude to Doctor of Physico-mathematical Sciences Professor N.I.Ionov and to Doctor of Physico-mathematical Sciences E.Ya.Zandberg for very valuable suggestions." Orig. art. has: 36 formulas and 1 figure. [15]

ASSOCIATION: Vsesoyuznyy nauchno-issledovatel'skiy institut monokristallov, Khar'kov (All-Union Scientific Research Institute of Single Crystals)

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EWA(m)-2 Pab-10/Pr-4/Pad/Pt-7/Pu-4 LJP(c) JD/WW/HW/JG
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AUTHOR: Chaykovskiy, E.F.; Pyatigorskiy, G.M.; Ptitsyn, G.V.

21

TITLE: Temperature hysteresis of positive surface ionization and the work function of a uniform emitter

SOURCE: Zhurnal tekhnicheskoy fiziki, v.35, no.6, 1965, 1132-1138

TOPIC TAGS: surface ionization, alkali metal, hysteresis, work function, platinum, molybdenum, nickel, rubidium, potassium, cesium

ABSTRACT: The authors calculate the critical threshold temperature T_c above which hysteresis of the threshold temperature for positive surface ionization on a uniform emitter does not occur. By using the known linear dependences of the ionic heat of adsorption and the emitter work function on the degree of surface coverage when the coverage is small, they obtain the relation $e(F - V)/kT_c = 2 + \log(g_a/g_i)$, where F is the work function, V is the ionization potential, and g_a and g_i are the statistical weights of the electron in the atom and ion, respectively. Corrections to this formula were calculated with

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ACCESSION NR: AP5015639

the nonlocalized adsorption model of E.F.Chaykovskiy and G.M.Pyatigorskiy (DAN SSSR 153,401,1963) and the localized adsorption model of V.M.Gavrilyuk (Ukr.fiz.zh.4,734,1959; Kinetika i kataliz 2,No.4,497, 1961) and found to be small (the calculations with the nonlocalized adsorption model are given). It is proposed that the correction be evaluated empirically for a given adsorption mechanism by employing an emitter with a known work function and that the empirical correction be used to derive the work functions of other emitters from surface ionization threshold hysteresis measurements. Surface ionization threshold temperature hysteresis measurements were made with Rb on W, Rb on Ni, K on Pt and Cs on Mo by the vapor method of E.Ya. Zandberg and N.I.Ionov (UFN 67,No.4,581,1959) using equipment that has been described elsewhere (E.F.Chaykovskiy, Yu.B.Skrynnik and G.M. Pyatigorskiy, PTE,1965). The 14 x 4 x 0.1 mm electrodes were made of well-oriented foil and were mounted with a 0.4 mm gap between them. Both electrodes were made of the same metal to avoid contamination of the emitter by sputtering. The surface ionization current was measured as a function of temperature (with both increasing and decreasing temperatures) for different values of the alkali metal vapor

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pressure. The threshold temperature at that vapor pressure above which threshold temperature hysteresis no longer occurs was determined by graphical interpolation. The data for Rb on W were used to obtain the correction term, and this correction was employed to derive the work functions of the other emitters from the remaining data. The work functions for the (111) face of Ni, the (111) face of Pt and the (100) face of Mo were thus found to be 4.57, 4.82 and 4.26 eV, respectively. The limit of error of these work functions is estimated to be 0.05 eV. "In conclusion, the authors consider it their pleasant duty to express their gratitude to Professor N.I.Ionov, Doctor of Physico-mathematical Sciences, and E.Ya.Zandberg, Doctor of Physico-mathematical Sciences, for very valuable discussions." Orig. art.has: 19 formulas and 3 figures.

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ASSOCIATION: Vsesoyuznyy nauchno-issledovatel'skiy institut monokristallov, Khar'kov (All-Union Scientific Research Institute for Single Crystals

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L 54770-65 EWT(1)/EPF(c)/EPA(w)-2/T/EWA(m)-2 Pab-10/Pr-4 IJP(c)
ACCESSION NR: AP5015645 UR/0057/65/035/006/1158/1159

AUTHOR: Chaykovskiy, E.F.; Ptitsyn, G.V.

TITLE: Positive surface ionization of sodium on oriented platinum strips

SOURCE: Zhurnal tekhnicheskoy fiziki, v.35, no.6, 1965, 1158-1159

TOPIC TAGS: surface ionization, work function, platinum, sodium

ABSTRACT: The authors have previously investigated the surface ionization of K on oriented Pt strips and obtained the value 4.7 ± 0.2 eV for the work function of the (111) face of Pt (ZhTF 35,No.3,1965). In the present paper they report results of similar measurements with the same apparatus of the surface ionization of Na on Pt. The measurements were undertaken to provide data for a case in which the ionization potential exceeds the work function. The measurements were made by the vapor method with 0.4 x 1.2 x 0.01 cm Pt electrodes separated by an 0.04 cm gap in a vacuum of 10^{-6} mm Hg. The elec-

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ACCESSION NR: AP5015645

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trodes were cleaned by prolonged heating at 1800°K, and x-ray investigations showed the surface to consist of outcroppings of (111) planes. The temperature dependence of the ion current was determined for Na vapor pressures from 2 x 10-5 to 6 x 10-3 mm Hg and from these data the value of 4.77 ± 0.07 eV was derived for the work function of the (111) face of Pt. This result is in good agreement with the value obtained similarly from measurements with K (cited above) and with that obtained by the authors (ZhTF 35,1122,1965) from measurements of the temperature above which surface ionization threshold temperature hysteresis does not occur. The small scattering of the experimental points showed that the contrast of the investigated platinum surface was low. "In conclusion, the authors consider it their pleasant duty to express their gratitude to Professor N.I.Ionov, Doctor of Physics-mathematical Sciences, and E.Ya.Zandberg, Doctor of Physico-mathematical Sciences, for very valuable discussions." Orig.art.has: 1 figure.

Card 2/3

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L 54770-65

ACCESSION NR: AP5015645

ASSOCIATION: Vsesoyuznyy nauchno-issledovatel'skiy institut monokristallov, Khar'kov (All-Union Scientific Research Institute for Single Crystals)

SUBMITTED: 21Nov64 ENCL: 00 SUB CODE: EM, IC

NR REF SOV: 004 OTHER: 001

Card 3/3 MB

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, G.V.

Transient processes in traction motors with mixed excitation and starting in stages. Trudy MEI no.29:114-125 '57.
(MIRA 13:3)

(Electric motors, Direct current) (Transients (Electricity))

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

MINOV, Dmitriy Konstantinovich; SHTERTSER, O.N., kand.tekhn.nauk, dotsent, retsenzent; KOSAREV, G.V., kand.tekhn.nauk, dotsent, retsenzent; PTITSYN, G.V., red.; BORUNOV, N.I., tekhn.red.

[Mechanical aspects of electric rolling stock] Mekhanicheskaia chast' elektricheskogo podvizhnogo sostava; ustroistvo, teoriia, proektirovanie. Moskva, Gos.energ.izd-vo, 1959. 381 p. (MIRA 12:9)

1. Leningradskiy politekhnicheskiy institut (for Shtertser).
2. Moskovskiy energeticheskiy institut im. Molotova (for Kosarev).
(Electric railroads--Rolling stock) (Streetcars)
(Trolleybuses)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, G. V. ~~Doc~~ Cand Tech Sci -- (diss) "Automatic systems of control for streetcars with ~~engines of~~ compound ~~excitation~~ motors." Mos, 1957. 20 pp 20 cm. (Min of Higher Education USSR. Moscow Order of Lenin Power Engineering Inst im V.M. Molotov), 100 copies (KL, 21-57, 103)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, I.F.

Silver water and its use in pharmacy. Sov.zdrav.Kir. no.1:53-55 Ja-F '63. (MIRA 16:3)

1. Iz apteki No.58, Sokulukskiy rayon Kirgizskoy SSR.
(SILVER--THERAPEUTIC USE)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

L 39549-66 EWT(d) GD/JXT(CZ)

ACC NR: AT6008686 SOURCE CODE: UR/3063/64/020/002/0040/0047

AUTHOR: Bartenev, L. S. (Engineer); Glebovich, G. V. (Candidate of technical sciences); Ptitsyn, K. N. (Engineer)

ORG: none*

TITLE: Peculiarities in the development of a superhigh-speed oscillograph

SOURCE: *Gorkiy. Politekhnicheskiy institut. Trudy, v. 20, no. 2, 1964. Radiotekhnika, elektronika i energetika (Radio engineering, electronics and power engineering), 40-47

TOPIC TAGS: oscillograph, electron beam oscillograph, superhighspeed oscillograph

ABSTRACT: The following difficulties were encountered in developing an electron-tube oscillograph with a direct beam-sweep time of 4—5 nsec: (1) The

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L 39549-66

ACC NR: AT6008686

nearest-to-linear midsegment of a steep voltage pulse (obtained from a nonlinear ferrite system) had to be used for sweep-voltage shaping (1 nsec, 500 v); (2) A diaphragm connected to the second anode had to be mounted behind the deflecting system; its window was so proportioned that the beam passed it only within the screen size, thus eliminating the undesirable stray lighting of the screen; (3) Test impulse distortion was minimized by using RK-3, RK-6, or RK-50-11-13-type Soviet-made cable and TW or coaxial deflecting system; (4) For calibrating the pulse duration, a telescopic coaxial delay line is recommended. The minimum distortion-permissible duration of the test impulse can be determined from: $\Delta t_o \geqslant 100\, \Delta t/a$, where Δt is the combined delay caused by the tube and cable and "a" is the specified percentage error of reproduction. The stability of the oscillograph operation largely depends on the stability of bias voltage on the slave stages. A laboratory hookup (oscillograms and circuit diagram shown) permitted recording 10^{-10} -sec processes. Orig. art. has: 3 figures and 10 formulas.

SUB CODE: 09 / SUBM DATE: none / ORIG REF: 006

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APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

KAZITSYNA, L.A.; KUPLETSKAYA, N.B.; PTITSYNA, V.A.; REUTOV, O.A.

Double diazonium salts of monovalent copper. Izv.AN SSSR.Otd. khim.nauk no.3:562-563 Mr '63. (MIRA 16:4)

1. Moskovskiy gosudarstvennyy universitet im. M.V.Lomonosova.
(Diazonium compounds) (Copper salts)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSA, S. S.

"The Treatment of Purulent Corneal Ulcers with Penicillin," Vest, Oftalmol., 28, No. 2, 1949. Eye Department, Ryazhan' Oblast Clinical Hospital imeni Semashko, -c1949-.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSA, S. B.

33584. Mestnoye Primeneniye Penitsillina Pri Gnoynykh Yazvakh Rogovoy Obolochki. Sbornik Nauch. Rabot (Ryaz. Obl. Otd. Zdravookhraneniya), Vyp. 2, 1949, c. 32-43 Bibliogr: 8 Nazv.

SO: Letopis'nykh Statey, Vol. 45, Moskva, 1949

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSA, S. B.

"A Rare Case of a Disease of the Recklinghausen Canals," Vest. Oftalmol., 27, No. 3, 1948; Eye Dept., Ryazan Oblast Clinical Hosp. imeni Senashko, -c1948-.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

FEDOROV, B. P.; PTITSINA, N. V.

Anthracene Derivatives

Investigation of anthracene derivatives. 10. Relative activity of mesocarbonic atoms of anthracene and of 9, 10-dichloroanthracene in reactions with oxidizing agents. B. P. Fedorov, N. V. Ptitsina. Izv. AN SSSR. Otd. khim nauk, No.1, 1952.

Monthly List of Russian Accessions, Library of Congress, September 1952. UNCLASSIFIED.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

FEDOROV, B.P.; PTITSINA, N.V.

Anthracene derivatives. I. Relative activity of ms-carbon atoms of anthracene and 9,10-dichloroanthracene in reactions with oxidizing agents. Bull. Acad. Sci. U.S.S.R., Div. Chem. Sci. '52, 135-47 [Engl. translation].
(CA 47 no.19:9950 '53)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.

USSR

Determination of dissociation constants of individual complexes by the method of shifting the equilibrium. II. Pritsyn and L. I. Vinogradova. Zhur. Obshchei Khimii 25, 217-20(1955); J. Gen. Chem. (U.S.S.R.) 25, 201-6(1955) (Engl. translation).—Consts. of successive steps of dissocn. of acido-complexes were detd. through the displacement of the equil. by reaction with various equivs. of ions (H^+ and Ag^+) that form insol. or undissocd. compds. Thus, if C moles of $[MA_4]^{---}$ are mixed with xg.-ions of Ag^+ to react thus: $[MA_4]^{---} + xAg^+ + xH_2O \rightleftharpoons [MA_{4-x}(H_2O)_x]^{3-x} + xAgA\downarrow$ (where A is the anion and $x = 1$ to 6), then $K_x = (xC - [Ag^+])K^x_{AgA}/[Ag^+]^{x+1}$, where K_{AgA} is the soly. product of the ppt. Malcolm M. Anderson

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B. V.

USSR/Inorganic Chemistry - Complex Compounds C.

Abs Jour : Referat Zhur - Khimiya, No 2, 1957, 4100

Author : Vinogradova, L.I., Ptitsyn, B.V.

Title : Determination of Instability Constants of Trioxalatoferriate of Potassium by the Method of Displaced Equilibrium

Orig Pub : Zh. neorgan. khimii, 1956, 1, No 3, 427-431

Abstract : Determination of instability constants (K) of trioxalatoferriate of potassium (I) is based on utilization of the previously described method (RZhKhim, 1955, 45708) of study of the equilibrium of the complex under investigation with ions that displace the equilibrium of secondary dissociation of complex particle due to formation of little soluble or little dissociated compounds. To study the stability of I Ag^+ and H^+ are utilized as such ions. Determined were the values of the thermodynamic constants: $K_1 = 2.3 \cdot 10^{-5}$; K(total) = 2.1 .

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.

C.

USSR/Inorganic Chemistry - Complex Compounds

Abs Jour : Referat Zhur - Khimiya, No 2, 1957, 4101

Author : Vinogradova, L.I., Ptitsyn, B.V.

Title : Determination of Instability Constants of Trioxalatoferriate of Potassium by Interaction of Iron Salt with Silver Oxalate

Orig Pub : Zh. neorgan. khimii, 1956, 1, No 3, 432-437

Abstract : From comparison of stability constants (K) of ferrioxalate complexes determined on the basis of a study of equilibrium of the system $Fe(NO_3)_3$ (I) - $Ag_2C_2O_4$ (II).

with the values of K determined by the method of displaced equilibrium (see preceding abstract), it follows that as a result of interaction between I and II there is formed predominantly the complex $[Fe(C_2O_4)_3]^{3-}$.

Average value of its total K at 25 ± 0.1° is $2.3 \cdot 10^{-20}$

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

USSR/Inorganic Chemistry - Complex Compounds C.

Abs Jour : Referat Zhur - Khimiya, No 2, 1957, 4101

The previously proposed dissociation scheme (Schaeffer H., Z. S. anorg. Chem., 1905, 45, 283), is thus rejected. However the new dissociation scheme is based on a not entirely correct assumption of an equality in concentration of the ions $[FeC_2O_4]^+$ and $[Fe(C_2O_4)_2]^-$. Therefore the above-stated values of K_2 and K_3 are approximate.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.; PETROV, V.F.

The oxidation potential of dichromates. Zhur.ob.khim. 26 no.12:3233-3239 D '56. (MLRA 10:7)
(Dichromates) (Oxidizing agents)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

GRINBERG, A.A.; NIKOL'SKAYA, L.Ye.; PETRZHAK, G.I.; PTITSYN, B.V.; FILINOV, F.M. [deceased].

Use of rongalite for the production of difficultly soluble derivatives of tetravalent uranium. Zhur. anal. khim. 12 no.1:92-94 Ja-F '57. (MLRA 10:4)

1. Radiyevyy institut AN SSSR, Leningrad.
(Uranyl compounds) (Sodium formaldehydesulfoxylate)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

GRINBERG, A.A.; PTITSYN, B.V.; FILINOV, F.M.; LAVRENT'YEV, V.N.

Preparation of uranium hexacarbonyl. Trudy Radiev.inst.AN SSSR 7:14-16 '56. (MLRA 10:5)

(Uranium organic compounds)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.

GRINBERG, A.A.; PTITSYN, B.V.; TEKSTER, Ye.N.

Physical and chemical properties of the aqueous solutions of complex uranyl oxalates. Trudy Radiev.inst.AN SSSR 7:74-86 '56. (MLRA 10:5)
(Uranyl oxalates)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B. V.

"Preparation of Weakly Soluble Compounds of Tetravalent Uranium With the Aid of Rongalite," by A. A. Grinberg, L. Ye. Nikol'skaya, G. I. Petrzhak, B. V. Ptitsyn, and F. M. Filinor (deceased), Radium Institute of the Academy of Sciences USSR, Leningrad, Zhurnal Analiticheskoy Khimii, Vol 12, No 1, Jan/Feb 57, pp 92-94

A method for the reduction of uranyl salts with rongalite is described. As compared with hydrosulfite, rongalite has the advantage that no sulfur is precipitated under the conditions investigated. Tetravalent uranium could be quantitatively isolated when the reduction was carried out in oxalic acid or hydrofluoric acid solutions. The optimum concentration conditions for the isolation of uranium as its oxalate have been established. (U)

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APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

VINOGRADOVA, L.I.; PTITSYN, B.V.

Determination of instability constants for potassium trioxalatoferrate by means of displaced equilibrium. Zhur.neorg. khim. 1 no.3:427-431 Mr '56. (MLRA 9:10)

(Potassium oxalatoferrate (III))

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

VINOGRADOVA, L.I.; PTITSYN, B.V.

Determination of instability constants of potassium trioxalatoferrate from the interaction of an iron salt with silver oxalate. Zhur.neorg.khim. 1 no.3:432-437 Mr '56. (MLRA 9:10)

(Potassium oxalatoferrate (III))

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B.V.

GRINBERG, A.A. (Leningrad); BARAYEVA, A.V. (Moscow); YATSIMIRSKIY, K.B. (Ivanovo); GOREMYKIN, V.I. (Moscow); BOLIY, G.B. (Moscow); FIALKOV, Ya.A. (Kiyev); YAKSHIN, M.M. (Moscow); KEDROV, B.M. (Moscow); GEL'MAN, A.D. (Moscow); FEDOROV, I.A. (Moscow); MAKSIMYUK, Ye.A. (Leningrad); VOL'KENSHTEYN, M.V. (Leningrad); ZHDANOV, G.S. (Moscow); PTITSYN, B.V. (Leningrad); ABLOV, A.V. (Kishinev); VOLSHTEYN, L.M. (Dnepropetrovsk); TROITSKAYA, A.D. (Kazan'); KLOCHKO, M.A. (Moscow); BABAYEVA, A.V.; TRONEV, V.G. (Moscow); RUBINSHTEYN, A.M. (Moscow) CHERNYAYEV, I.I.; GRINBERG, A.A.; TANANAYEV, I.V.

Explanation of the transeffect. Izv.Sekt.plat.i blag.met. no.28: 56-126 '54. (MLRA 7:9)

(Compounds, Complex) (Platinum)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"

"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN B.V.

NIKOL'SKIY, B.P., redaktor; DOLGOV, B.N., redaktor; ZAL'KIND, Yu.S. [deceased] redaktor; MORACHEVSKIY, Yu.V., redaktor; POZIN, M.Ye., redaktor; PTITSYN, B.V., redaktor; SMIRNOV, N.I., redaktor.

[The chemist's handbook]Spravochnik khimika. Vol. 3. [Chemical equilibrium and kinetics. Solutions. Electrochemistry. Analytical and technical chemistry] Khimicheskoe ravnovesie i kinetika. Rastvory. Elektrokhimiia. Analiticheskaia i tekhnicheskaia khimiia. Leningrad, Gos.nauchno-tekhn. izd-vo khim. lit-ry. 1952. 1190 p. [Microfilm]
(Chemistry--Handbooks, manuals, etc.) (MLRA 7:10)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B. V., GORYUKHINA, V. A., KHODUNOV, P. A.

USSR (600)

Sulphur

Influence of the oxidation and deoxidation potential on the process of oxidation of coordination bound sulphur. Izv. Sekt. plat. i blag. met. no. 25, 1950.

Monthly List of Russian Accessions, Library of Congress, April 1952. UNCLASSIFIED.

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN, B. V.

"A Potentiometric Method of Dosage of Pt and De L'Ir Totaux,"

SO: Dok. AN, 51 No. 9, 1946. c1946-.

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PTITSYN, B.V.

"A Potentiometric Method of Dosage of Pt and De L'Ir Totaux," Dok. AN, 51, No. 9, 1946. o-1946-.

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PTITSYN, B. V.

"The Mechanism of the Reaction of Sodium Thiosulphate with Halogene-Iridates and on the Oxidation-Reduction Potentials of Halogene-Iridates," Zhur. Obshch. Khim., 15, No. 4-5 1945; Mbr., Lab. Inorganic Chemistry, Leningrad Chemico-Technological Inst.,-1941-.

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PTITSYN, G. A.

Tekhnologiya proizvodstva turbinnykh reduktorov
(Technology of Turbine Reduction Gear Production, by)
S. G. Printsental' i G. A. Ptitsyn. Moskva, Mashgiz, 1953.

303 p. illus., diagrs., tables.

"Literatura": p. (290)

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"APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2

PTITSYN. G. A.

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Zuboreznyye stank; spravochnoye posobie
(Gear-cutting machines; reference handbook, by)
V. N. Kokichev i G. A. Ptitsyn.
Moskva, Mashgiz, 1954.
355 p. illus., diagrs., tables.
"Literatura i Istochnike": p. (351)

APPROVED FOR RELEASE: 06/15/2000 CIA-RDP86-00513R001343510020-2"