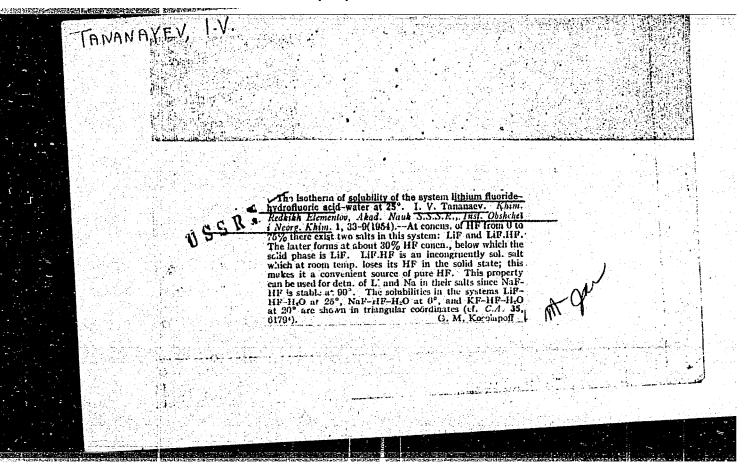
Jead, Theorium Systems	"Physicochemical Analysis of Systems of Important to Analytical Chemistry. XX. The Solubility of Precipitates in Complex (Really Existing) Analytical Systems," I. V. Tananayev, I. B. Misetskaya, A. D. Vinogradova, Inst of Gen and Inorg Chem, Acad Sci USSR	Studied soly in the system PbSO4 - Th(NO3 - Li2SO4-Studied soly in the system PbSO4 - Th(NO3 - Li2SO4-Hookel formula for calcg HoO at 25°C. The Debye-Hueckel formula for calcg the soly of PbSO4 is not suitable for this system, the soly of PbSO4 is not suitable for this system, the soly of PbSO4 is not suitable for this system, because of the marked chem interaction accompanied	USSR/Chemistry - Lead, Theorium Systems (Contd) Systems (Contd) 2. Considers by formation of ions of the type ThSO2 the importance of physicochem enalysis for theory the importance of pphn reactions, and a diagram shows and practice of pphn reactions, and a diagram shows the following types of ternary systems: ppt - the following types of ternary systems: ppt - possible systems with ppts in dependence on the ion type of the components and the character of the type of the components and the system.	CR 47 no.19: 9849 13)
USSR/Chemistry	"Physicochemics to Analytical C Precipitates in cal Systems," A. D. Vinogradisci USSR	Studied soly in Exp at 25°C. The soly of Presente of the	ussn/chemistr. by formation the importance the following electrolyte w possible syst type of the process of in	TANANAYEV, I. V.



TANANAYEV, I.V.; GIUSHKOVA, M.A.; SEYFER, G.B. Chemistry of lanthanum ferrocyanides and their application in ana-V.

lytic chemistry. Khim.redk.elem. no.1:58-86 154.

1. Institut obshchey i neorganicheskoy khimii im. H.S.Kurnakova AN SSSR. (Lanthanum ferrocyanide)

TANANAYN, I.V.; DETCHMAN, E.N.

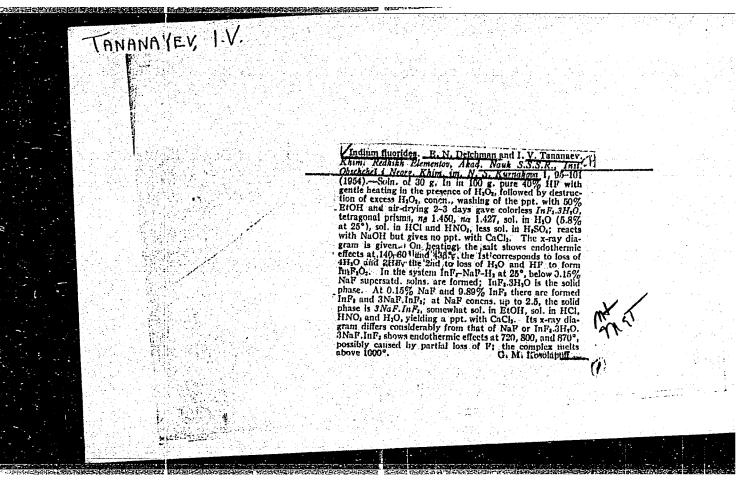
Study of the solutions of indium fluoride and ozalate. Ihim.

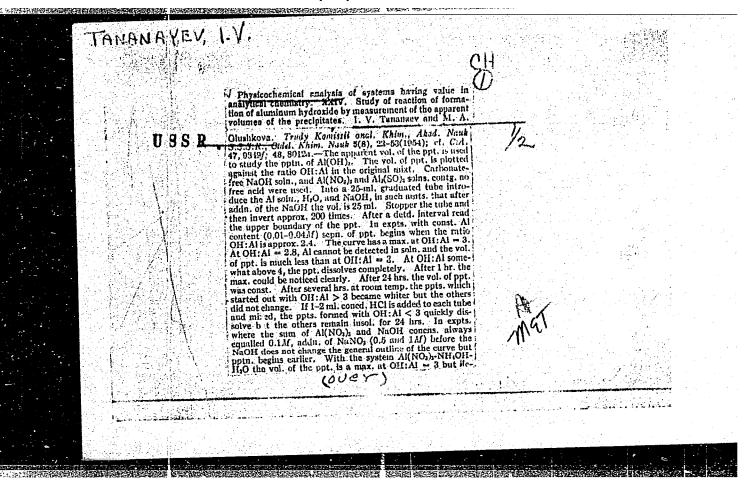
(MIMA 8:3)

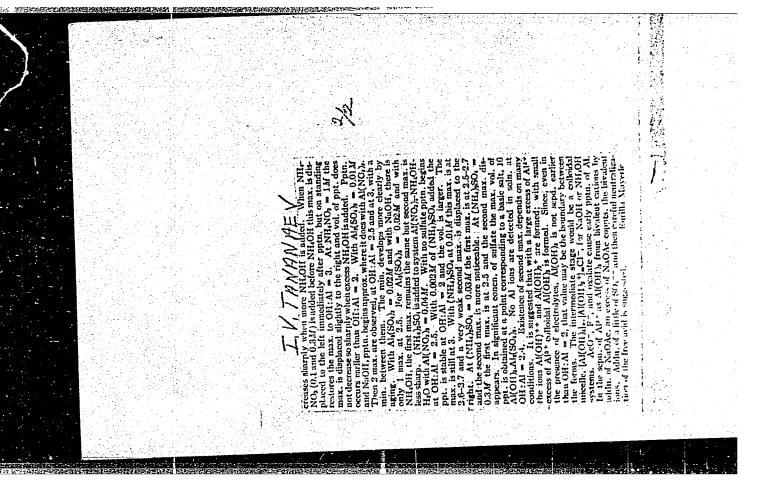
redk.elem. no.1:87-94 '54.

1. Institut obshchey i neorganicheskoy khimii im. N.S.Kurnakova
AN SSSR.

(Indium salts)







GRINBERG, A.A. (Leningrad); BARAYEVA, A.V. (Moscow); YATSIMIRSKIY. K.B. (Ivanovo); GOREMYKIN, V.I. (Moscow); BOLIY, G.B. (Moscow); FIAL— KOV, YA.A. (Kiyev); YAKSHIN, M.M. (Moscow); KEBOV, B.M. (Moscow); GEL'MAN, A.D. (Moscow); FEDOROV, I.A. (Moscow); MAKSIMYUK, Ye.A. (Leningrad); VOL'KENSHTEIN, M.V. (Leningrad); ZHDANOV, G.S. (Moscow); PTITSYN, B.V. (Leningrad); ABLOV, A.V. (Kishinev); VOLSHTEIN, L.M. (Dnepropetrovsk); TROITSKAYA, A.D. (Kazan'); KLECHKO, M.A. (Moscow); BARAYEVA, A.V.; TRONEV, V.G. (Moscow); RUBINTEYN, A.M. (Moscow) CHERNYATEV, I.I.; GRINBERG, A.A.; TANANAYEV, I.V. Explanation of the transeffect. Isv. Sekt. plat. i blag.met. no. 28: 56-126 '54. (MIRA 7:9) (Compounds, Complex) (Flatinum)

TANANAYEV, I. V. and BAUGOVA, N. V.

"On the Formation reaction of gallium hydroxide", Khimiya Redkikh Elementov, No. 2, p.12, 1955.

The mechanism of the formation of gallium hydroxide was investigated by a study of the system $GaCl_3$ -NaOH-H₂O. The measurements of solubility, light absorption and volumes of precipitates were used for the investigation. The reaction takes place in five stages, depending on the molar ratios of NaOH to $GaCl_3$ in the initial mixture, with the successive formation of $Ga(OH)Cl_2$, $Ga(OH)_2Cl_2$, $Ga(OH)_2Cl$ (soluble basic salts), $Ga(OH)_2.8Cl_{0.2}$ (insoluble basic salt), $Ga(OH)_3$ and then soluble gallate.

SO: D-413171

TANANAYEV, I. V. and BAUJOVA, N. V.

"A study of the chemistry of gallium fluorides and their utilization for the separation of gallium from other metals", Khimiya Redkikh Elementov, No. 2, p 21, 1955.

A method of preparation of GaF₃3H₂O by the action of hydrofluoric acid on metallic gallium. The reaction between GaCl₃4 and HF in aqueous solutions was investigated, the formation of a stable GaF₂ ion was established. Solubility in the system: GaF₃-NaF-H₂O at 25°C was investigated. The formation of a double salt of the composition 13NaF. 5GaF₃, practically insoluble in sodium fluoride solutions was established. On the basis of the latter a method of quantitative precipitation of gallium from Zn, Co, Ni, W, Mo, Cd and Cu was developed.

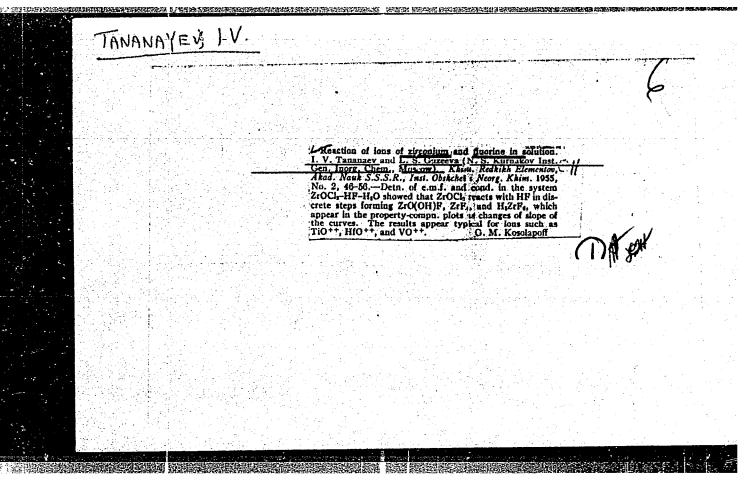
SO: D-413171

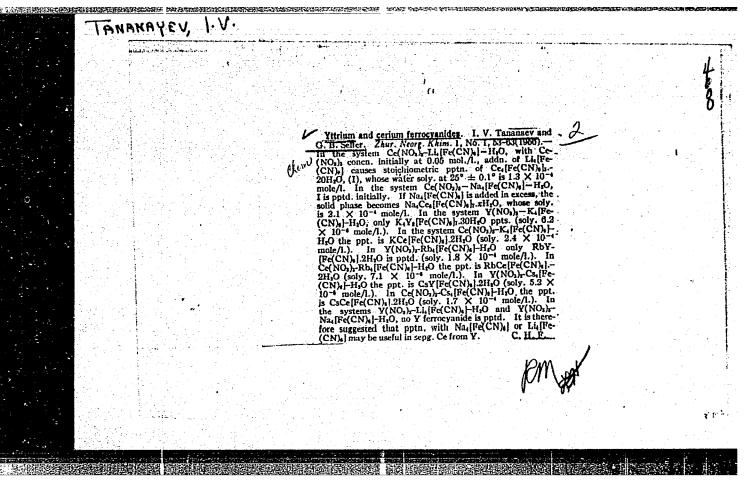
TANANAYEV, I. V. and DEYCHMAN, E. N.

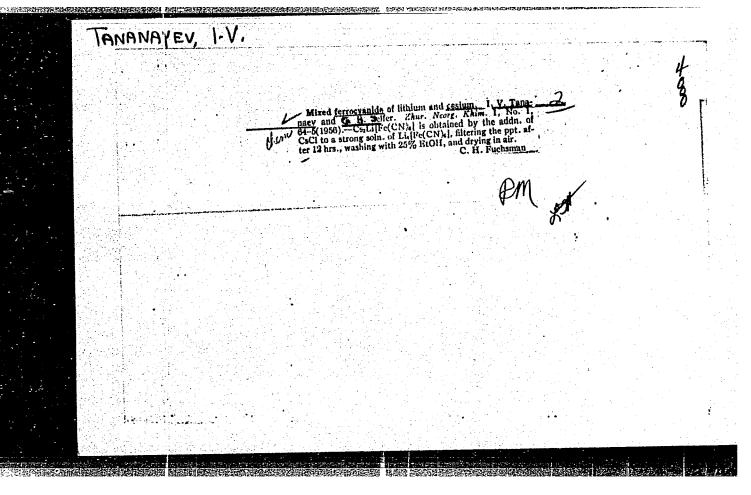
"On indium ferrocyanides", Khimiya Redkikh Elementov, No. 2, p 37, 1955.

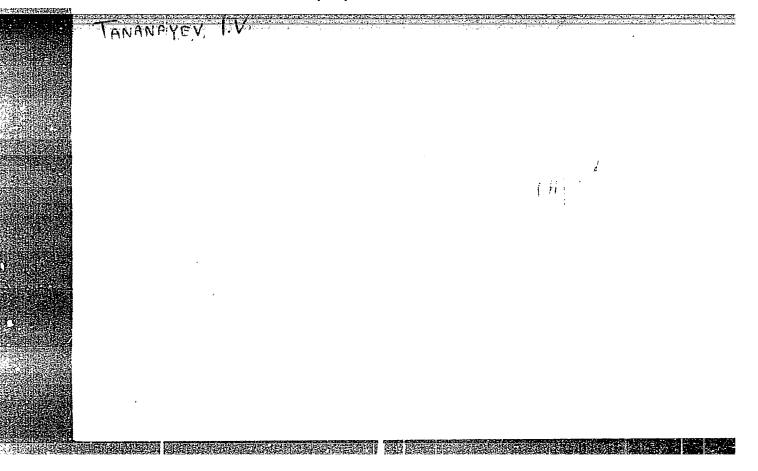
The systems: InCl₃ -Li₄Fe(CN)₈-H₂O; INCl₃-Na₄Fe(CN)₆H₂O and INCl₃-K₄Fe (CN)₆-H₂O were investigated using solubility electroconductivity, potentiometric and turbidometric methods. Concentrations of components were those used uncer normal analytical conditions. It was found that in the first two systems indium ions react forming salts of the normal composition In₄ $\sqrt{\text{Fe}(\text{CN})_6/3}$ and in the third system, in addition to the normal, a double salt is formed uncer certain conditions.

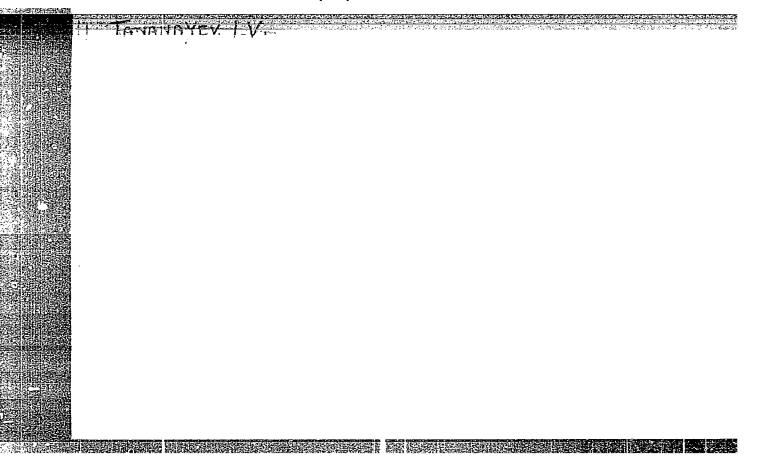
SO: D-413171

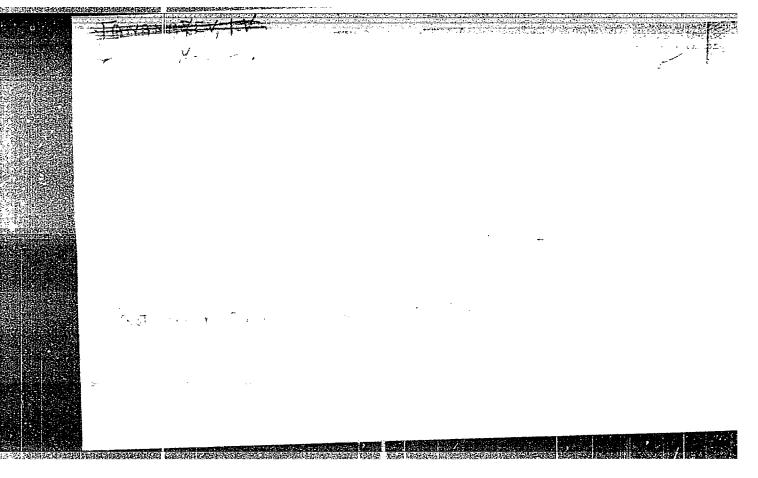


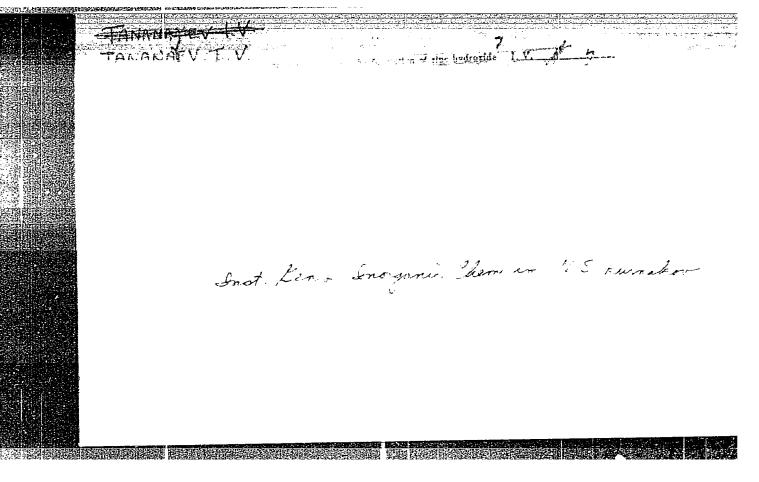


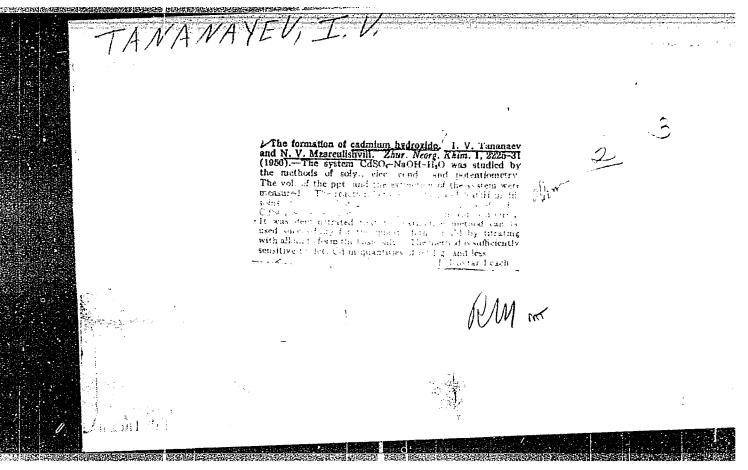












TANANAYEVY V

USSR/Physical Chemistry - Thermodynamics. Thermochemistry. Equilibrium. Physico-chemical Analysis. Phase Transitions, B-8

Abst Journal: Referat Zhur - Khimiya, No 1, 1957, 380

Author: Tananayev, I. V. Nikolayev, N. S., Buslayev, Yu. A.

Institution: None

Title: Investigation of the System HF-ZrFh-H₂O by the Isothermal Solubility Method (Isotherm 0.5°)

Original

Periodical: Zh. neorgan. khimii, 1956, Vol 1, No 2, 274-281

Abstract: The solubility at 0.5° in the system HF-ZrF_h-H₂O has been investigated for the range 0-100 percent HF. The following solid phases were found in the system: ZrOF₂·2H₂O(I), ZrF_h·3H₂O(II), H₂ZrF₆·H₂O(III), and ZrF_h. From thermographic data thermal decomposition reactions for I, II, III, and ZrOF₂ have been established. A method has been developed for the determination of both F and Zr when present

together.

Card 1/1

TANANAYEV, I.V.; MZAREULISHVILI, N.V. Study of the formation of silver hydroxide. Zhur.neorg.khim. 1 no.8:1826-1831 Ag 156. (Silver hydroxide)

CIA-RDP86-00513R001754820007-7 "APPROVED FOR RELEASE: 07/13/2001

TANANAEV, I.V.

USSR/Inorganic Chemistry - Complex Compounds.

C.

: Ref Zhur - Khimiya, No 9, 1957, 30291 Abs Jour

: Tananayev, I.V., Seyfer, G.B. Author

Inst

: Mixed Ferrocyanides of Magnesium, Rubidium and Cesium Title

: Zh. neorgan. khimii, 1956, 1, No 9, 2017-2023 Orig Pub

: On the basis of the results of a study of the system Abst

MCl - Mg / Fe(CN) /-H,0, wherein M is Rb or Cs, by the solubility method, it was found that the composition of the solid phase that separates, corresponds to the formula 3M4/Fe(CN), J. 4Mg2/Fe(CN), J.12H, O. The composition is not altered on use of an excess of both compo-

nents, in the case of Rb, while in the case of Cs, an excess of Mg./Fe(CN)6/causes the formation of Cs./Fe(CN)6/.2Mg./Fe(CN)6/.10H.0.

Card 1/1

IANANAYEN, I.U.,

USSR/Inorganic Chemistry. Complex Compounds.

C

Abs Jour : Ref Zhur - Khimiya, No. 8, 1957, 26490.

Tananayev, I.V., Mzareulishvili, N.V. Author Inst

Title : Study of Reaction of Zinc Hydroxide Formation.

Orig Pub : Zh. neorgan. khimii, 1956, 1, No. 10,

2216 - 2224.

Abstract

The system ZnSO₄ - NaOH - H₂O was studied by the solubility, the electrical conduc-tivity, the potentiometric methods and the methods of light extinction and of precipitate volume measurement. It was established that the reaction between ZnSO4 and NaOH in an aqueous solution proceeds with the formation first of 4Zn(OH)2.ZnSO4 (I) and of Zn(OH)2 (II) after that. Considering

Card 1/2

USSR/Inorganic Chemistry. Complex Compounds.

C

Abs Jour : Ref Zhur - Khimiya, No. 8, 1957, 26490.

the obtained data, the conclusion was arrived at that the quantitative determination of Zn²⁺ by titration with an alkali solution is possible by the following methods:

a) geometrical (by the formation of I and II)

b) conductometric (by the formation of I),

c) potentiometric (by the formation of II).

Card 2/2

THE CONTROL OF THE PROPERTY OF

TANANAYEU, I.V.

USSR/Inorganic Chemistry. Complex Compounds.

C

Abs Jour

: Ref Zhur - Khimiya, No. 8, 1957, 26493.

Author

Tananayev, I.V., Mzareulishvili, N.V.

Inst Title

Study of Reaction of Cadmium Hydroxide

Formation.

Orig Pub

Zh. neorgan. khimii, 1956, 1, No. 10, 2225 - 2231.

Abstract

The system CdSO₁ - NaOH - H₂O was studied by the solubility, the electrical conduc-tivity, the potentiometric methods and the methods of light extinction and of precipitate volume measurement. It was established that the reaction between CdSO₄ and NaOH in a diluted aqueous solution proceeds with the formation first of 4Cd(OH)2.CdSO4 and of

Card 1/2

BANGARAN PARA GRADALISA DER SANTAR DE SA

USSR/Inorganic Chemistry. Complex Compounds.

C

Abs Jour : Ref Zhur - Khimiya, No. 8, 1957, 26493.

Cd(OH)₂ after that. Considering the obtained data, the conclusion was arrived at that the method of light extinction is applicable to the quantitative determination of Cd²⁺ by titration with an alkali solution.

Card 2/2

TANANAYEV, I.V.; UL'YANOV, A.I.

的,我们就是我们的一个人,我们就是我们的一个人,我们会是一个人,我们会会们的一个人,我们就是我们的一个人,我们就是我们的一个人,我们就是我们的一个人,我们就会没

Physicochemical analysis of systems important in analytical chemistry. Part 25. Study of the coprecipitation of alkali metal sulfates with BaSO4. Trudy Kom.anal.khim. 7:3-20 '56. (MIRA 9:9)

1. Institut ebshchey i neorganicheskey khimii imeni H.S. Kurnakeva AN SSSR. (Sulfates)

TANANAYEV, I. V.

"The possibility of separating francium from cesium."

report presented at The Use of Radioactive Isotopes in Analytical Chemistry, Conference in Moscow, 2_4 Dec 1957
Vestnik Ak Nauk SSSR, 1958, No. 2, (author Rodin, S. S.)

Heodysium ferrocyanides. Khim.redk.elem. no.3:28-ho '57.

(MERA 10:8)

1.Institut ebehchey i neorganicheskey khimii im. M.S. Kurnakeva
Akademii nauk SSSR.

(Neodymium ferrecyanides)

CIA-RDP86-00513R001754820007-7 "APPROVED FOR RELEASE: 07/13/2001

TANANAYEV. I. V.

137-58-2-4398

Translation from: Referativnyy zhurnal, Metallurgiya, 1958, Nr 2, p 300 (USSR)

Tananayev, I.V., Bausova, N.V.

Gallium Ferrocyanides and Their Analytical Significance (Fer-AUTHORS: rotsianidy galliya i ikh analiticheskoye znacheniye) TITLE:

Khimiya redkikh elementov, Nr 3, 1957, pp 41-56

Tests of solubility, light absorption, electrical conductivity, PERIODICAL: and e.m.f. were used to study the reaction of Ga ions with the ABSTRACT:

ferrocyanides of Li, K, and Na. In all the systems studied it was found that Ga4 [Fe(CN)6] 3 forms. Whenever surplus K4 [Fe(CN)6] and Na4 [Fe(CN)6] were present, formation was observed of mixed salts of NaGa [Fe(CN)6] (the H2O dissociating) and KGa [Fe(CN)6] (soluble congruently). On the basis of the data obtained, new methods are proposed for determining Ga. Potentiometric titration with an Na4 [Fe(CN)6] solution made possible the determination of Ga in the presence of an Al

content 100 times greater. With Na4 [Fe(CN)6] in a 100-cc solution it was possible to determine 0.2-30 mg of Ga by means of the heterodyne method. Amperometric titration with a K4

[Fe(CN)6] solution could be carried out in the presence of Card 1/2

137-58-2-4398

Gallium Ferrocyanides and Their Analytical Significance

large quantities of Al, since the presence of the Al was reflected only in the slope of the titration curve.

N.G.

1. Gallium ferrocyanides—Analysis

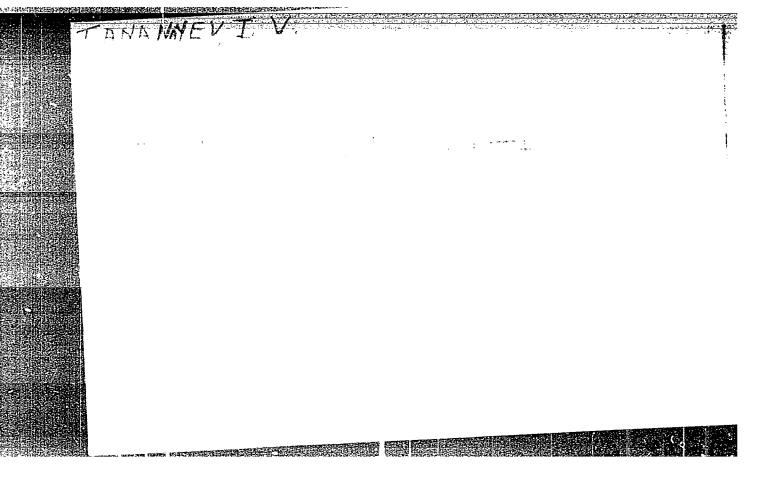
Card 2/2

TAMANAYEV, I.V.

Study of the reaction of indium hydroxide fermation. Khim; redk.
(MLEA 10:8)
elem. no.3:73-86 '57.

1.Institut obshehey i meorganicheskoy khimii im. H.S. Kurnakeva
Akademii nauk SSSR.

(Indium hydroxides)



62-12-1/20

New Data Concerning the Chemistry of Some Rare Elements (Novyye Tananayev, I.V. AUTHOR:

dannyye o khimii nekotorykh redkikh elementov) Lecture Delivered at the Meeting of the Department of Chemical Sciences of the AN USSR on October 31, 1957 (Doklad na sessii TITLE: otdeleniya khimicheskikh nauk Akademii nauk SSSR 31 oktyabrya

Izvestiya AN SSSR Otdelenije Khimicheskikh Nauk. 1957, Nr 12, PERIODICAL:

pp. 1421-1428 (USSR)

Among the questions of greatest interest at present in physical chemistry there is also that of the elaboration of new reactions of rare alkaline metals. Nearly all of them form practically unsoluble mixed ABSTRACT:

ferrocyanides with alkaline metals (see table 1). When solving the problem concerning easier methods for the production of a particularly pure liquid fluorine hydrogen and fluorine, the fluorine compounds of lithium, rubidium and cesium will attain particular im-

portance. The lecture contains new data concerning the composition and the physical properties of various compounds of some rare elements. Among other things, also a survey is given of the insoluble

mixed ferrocyanides, composed of various metals with lithium,

Card 1/2

New Data Concerning the Chemistry of Some Rare Elements. Lecture Delivered at the Meeting of the Department of Chemical Sciences of the AN USSR on October 31, 1957

62-12-1/20

rubidium and cesium. In this connection the lecturer recommends a number of new reagent precipitations. A rule governing the modification of the composition of mixed ferrocyanides containing rare alkaline earth metals was determined. Further, new details are given concerning the composition and durability of fluorides, oxalates, tartrates, and other compounds of gallium, indium, zirconium and germanium. There are 2 tables.

Institute for General and Inorganic Chemistry imeni N.S.Kurnakov ASSOCIATION:

AN USSR (Institut obshchey i neorganicheskoy khimii im.

N.S.Kurnakova Akademii nauk SSSR).

October 9, 1957 SUBMITTED:

Library of Congress AVAILABLE:

1. Chemical engineering-Conference 2. Flourine-Liquid 3. Flourine hydrogen-Liquid 4. Lithium 5. Rubium Card 2/2

CIA-RDP86-00513R001754820007-7" APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754820007-7 "APPROVED FOR RELEASE: 07/13/2001

TANANAYEV I.V.

560 Tananaev, I.V., Seifer, G.B., and Glushkova, M.A.

Ferrocyanides of Trivalent Iron (O Ferrotsianidakh Trekhvalent-AUTHOR: TI'ILE:

novo Zheleza.)

PERIODICAL: "Zhurnal Neorganicheskoy Khimii" (Journal of Inorganic Chemistry,

Vol.II, No.2, pp.268-280. (U.S.S.R.) 1957

In this investigation the behaviour of the cyanides of trivalent ABSTRACT:

iron in systems with lithium, sodium, potassium, rubidium and caesium has been studied by two methods: that of solubility and that of e.m.f. measurement. The equilibrium solutions obtained in the solubility experiments were used for the parallel study o the systems by the e.m.f. method using a platinum electrode with a saturated calomel electrode. From the results obtained conclusions are drawn on the general behaviour of trivalent iron ions in and ferrocyanides of different alkali metals when present together in systems. Differences due to the nature of the individual alkali metals were also noted. The tendency, to all the alk li metals was the formation of mixed ferrocyanides of iron although the nature of these compounds was very different for the different alkali metals. The stoichiometrically normal iron ferrocyanide Fe4 [Fe(Cn)6]3 is not formed with any of the systems studied. On incomplete precipitation of iron it either becomes adsorbed (systems with Li4 [Fe(Cn)6], and Na4 [Fe(Cn)6]), or remai

Card 1/2

Ferrocyanides of Trivalent Iron (Cont.)

K - Rb - Cs.

AND AND THE PARTY OF THE PARTY

560

in excess (systems with K_4 (Rb_4 , Cs_4) [Fe(Cn)6]), an effect evidently due to coprecipitation of the alkali metal.

ly due to coprecipitation of the alkali metal.

Ma Fe(Cn) enters the precipitate as soon as the precipitation of iron is complete even with lithium and sodium. This is especially pronounced with rubidium and caesium. The contents of alkali metal in the precipitate increased according to the series

It appears that with potassium, rubidium and caesium mixed salts of the simplest type M Fe [Fe (Cn)6] are formed at first. It is very likely that the salts RbFe [Fe (Cn)6] and CsFe [Fe (Cn)6] do exist, but they react easily with M4 [Fe (Cn)6] to form a second mixed salt of the composition M6Fe2 [Fe (Cn)6]3.

Although the continued adsorption of M₄ [Fe(Cn)₆] by the precipitate makes the interpretation of the experimental curves difficult the possibility of the formation of new incongruently soluble mixed salts of definite composition is not excluded.

As regards the strength of the bond in mixed ferrocyanides of iron and alkali metals, the latter can be arranged in the usual series with caesium having the greater strength and lithium the least. Apparently in the mixed salts the alkali metals can mutually replace each other according to the above order. From this point o view mixed ferrocyanides of trivalent iron can be considered as inorganic prototypes of ion-exchange resins.

14 Figures and 5 Tables.

Card 2/2

Mixed Ferrocyanides of Thallium, and Strontium (Osmesham) Mixed Ferrocyanides of Thallium. I. Mixed Ferrocyanides of Thallium. I. Mixed Ferrocyanides of Thallium. (Osmeshannykh Thallium with Magnesium, Calcium and Strontium. (Osmeshannykh Thallium with Magnesium, I. Smeshannye Ferrotsianidy talliya Ferrotsianidakh Talliya Strontsiem). TANANAYEV, I.V. and Glushkova, M.A.
R: "Zhurnal Neorganicheskov Khimii" (Journal of Inorganic Chemistry Vol. II, No.2, pp.281-286. (U.S.S.R.) AUTHOR: With the object of embracing as large a number of possible types of compounds formed by ferrocyanides of investigation the ferrocyanides of other metals. systems for investigation types of compounds formed by ferrocyanides of investigation the ferrocyanides of other metals, systems of the metals reptate ferrocyanides of other metals, systems of the metals reptate ferrocyanides of other metals, systems of the metals reptate ferrocyanides of the such a way that the valency the present of the should be different. In the systems were chosen them should be different. In the systems is sented in the study of solubility in by the physicon resented in the study of solubility in that mixed ferrocyanic results of the study of the study of the mixed formed, except with magnesium chemical analysis method. It was found, except with magnesium of the type of the solubility of the mixed from the type of the solubility of the mixed from the solubility of the solubility TITLE: PERIODICAL: ABSTRACT: cyaniaes of thallium and the alkali metals increased from agreesium to strontium being 2 x 104, 3.4 x 10-4 and 3.6 x 10-3 mol/litre at 2500. Comparison of data from the liter magnesium to strontium being 2 x 10, 3 data from the liter 10-3 mol/litre at 2500. Comparison of data from the alkali-parth or the composition of mixed farmovanides of alkali-parth On the composition of mixed ferrocyanides of alkali-earth alkali metals and those obtained in the present investigat on the composition of mixed ferrocyanides of alkali-earth investigat alkali metals and those obtained in the present investigate alkali that the analogous thalling compounds we obtained of Shows that the analogous thallium compounds we obtained c.

card 1/2

APPROVED HORREUGASEC 97/12/2001 That I That I was I Mixed Ferrocyanides of Thallium with Magnesium, Calcium and Strontium (Cont.)

resemble them: this is indicated by the fact that they all belong to the single type M2E [Fe(CN)6]. As regards solubility in water, the compounds of thallium considered most closely resemble the corresponding rubidium and caesium

There are 7 references, one of them Russian. There are 3 figures and 3 tables.

4_

Received on 22nd October, 1956.

Card 2/2

TANANAYEV, I.V.

78-3-12/35

AUTHORS: Tananayev, I. V. and Levina, M. I.

TITLE: Some Data on the Structure of Mixed Ferrocyanides.

(Nekotoryye dannyye o strogenii smeshannykh ferrotsianidov).

PERIODICAL: Zhurnal Neorganicheskoy Khimii, 1957, Vol.II, Nr.3, pp. 576-585. (USSR)

ABSTRACT: The addition of silver nitrate to an aqueous suspension of any mixed ferrocyanide leads to the displacement by the silver ions of both cations of the mixed salt from the precipitate into the solution. If the silver intrate is added gradually and the concentration of the nitrate is added gradually and the solution is determined cations of the mixed salt in the solution is determined on each addition, the order and degree of replacement of on each addition, the order and degree of replacement of the heavy and alkali metals of the mixed salt by silver can be found. This is the basis of the method used in the present investigation, except that the process was the present investigation, except that the process was supernatent liquid by measuring its radioactivity.

The following compounds again was used as the tracer.

Card 1/2 were studied: K4N14 [Fe(CN)6] 3 and

78-3-12/35

Some Data on the Structure of Mixed Ferrocyanides.

5Co₂[Fe(CN)₆].K₄[Fe(CN)₆] and also Ni₂[Fe(CN)₆]. The method used enabled the relative strength of the bonds of the metals in the outer sphere of mixed ferrocyanides with ferrocyanide ions to be found, and it is suggested that this should influence the way in which the chemical formulae of metal mixed ferrocyanides should be written. Based on a study of the conditions for the formation of KAg[Fe(CN)₆] analytical methods are recommended for the determination of small quantities of silver and ferrocyanide in solutions radiometrically with the use of Agllo and also potentiometrically. There are 7 figures, 13 tables, and 11 references, 3 of which are Slavic.

SUBMITTED: October 27, 1956.

AVAILABLE: Library of Congress.

Card 2/2

14 manayer J.V

78-3-13/35

AUTHORS: Tananayev, I. V. and Glushkova, M. A.

TITLE: Mixed Ferrocyanides of Thallium. (O smeshannykh ferrotsianidakh talliya.) II. Mixed Ferrocyanides of Thallium with Copper and Nickel. (II. Smeshannyye ferrotsianidy talliya s med'yu i nikelem.)

PERIODICAL: Zhurnal Neorganicheskoy Khimii, 1957, Vol.II, Nr.3, pp. 586-593. (USSR)

ABSTRACT: This investigation had as its aim to fill the gap in the literature on the nature of the reaction of copper and nickel ions with ferrocyanide ions in the presence of thallium. The potential importance of such an investigation is that the information it provided will be used to decide whether thallium in very low concentration could be isolated from solutions containing certain non-ferrous metals. Ordinary methods of chemical analysis were used in this investigation to study systems consisting of TI*, E2* (E = Cu, Ni) and [Fe(CN)6] 4. It is concluded that the slightly card 1/2 soluble mixed ferrocyanide of thallium and copper,

78-3-13/35

Mixed Ferrocyanides of Thallium. II.

Tl₂Cu₂ [Fe(CN)₆]₂, is formed from the appropriate ions for the whole range of concentrations dealt with. The solubility of the compound formed as a result of the displacement of Cu² ions by Tl ions from a precipitate of Cu₂ [Fe(CN)₆] was found to be so small that a suspension of this precipitate can be used to separate small quantities of thallium from solution. The mixed salt Tl₄Ni₄ [Fe(CN)₆]₃ was formed by the reaction of Tl with Ni² and [Fe(CN)₆]⁴ ions. The solubility of the precipitate of this salt, formed by the reaction of Tl ions with a precipitate of nickel ferrocyanide is so low that it, too, can be used for separating small quantities of thallium from solution. There are 4 figures, 7 tables, and 6 references 2 of which are Slavic.

SUBMITTED: October 26, 1956.

AVAILABLE: Library of Congress.

Card 2/2

TANANAY EV, I.V.

78-3-14/35

AUTHORS: Tananayev, I. V. and Glushkova, M. A.

Ferrocyanides of Thallium. (O Ferrotsianidakh talliya.) III. Mixed ferrocyanides of thallium with TITLE:

uranyl. (III. Smesharryye ferrotsianidy talliya s

uranilom.)

PERIODICAL: Zhurnal Neorganicheskoy Khimii, 1957, Vol.II, Nr.3, pp. 594-599. (USSR)

ABSTRACT: An account is given of the study of the solubility of the system TlNO₃UO₂(NO₃)₂_Li₄[Fe(CN)₆]-H₂O at 25°C. Solutions of these compounds were used, the molarity of the thallium nitrate solution being determined by precipitating thallium as chromate, that of lithium cyanide by titration of a known volume of the original solution by potassium manganate in sulphuric acid using crystal vielet as an indicator. The concentration of the uranium salt was determined by precipitation of (UO2)2+ in the presence of carbonate-free ammonium nitrate, the precipitate after washing and calcining

Card 1/2 being weighed as U308. No formation was observed of

78-3-14/35

Ferrocyanides of Thallium. III.

the simple ferrocyanide of uranyl, the reaction of the ions leading to the formation of two mixed ferrocyanides of thallium and uranium with the compositions: $Tl_2(UO_2)_3[Fe(CN)_6]_2$ and $Tl_4(UO_2)_4[Fe(CN)_6]_3$. All the alkaline metals and thallium can be arranged in the following series with respect to their ability to displace $[UO_2]_2^2$ ions from the precipitate of $(UO_2)_2[Fe(CN)_6]$: T1 is $\ge Cs > Rb > K > Na > Li$, from which it follows that suspensions of this precipitate can be used for separating small quantities of thallium from solution. There are 2 figures, 6 tables, and 5 references 2 of which are Slavic.

SUBMITTED: November 19, 1956.

AVAILABLE: Library of Congress.

Card 2/2

CIA-RDP86-00513R001754820007-7 "APPROVED FOR RELEASE: 07/13/2001 O SEE SECTION OF THE PROPERTY OF THE PROPERTY

TANANAYEV, L.

78-3-15/35

and Seyfer, G. B. AUTHORS: Tananayev, I. V.

Mixed Ferrocyanides of Calcium with Rubidium and Caesium. (Osmeshannykh ferrotsianidakh kal'tsiya TITLE: s rubidiyem i tseziyem.)

PERIODICAL: Zhurnal Neorganicheskoy Khimii, 1957, Vol.II, Nr.3, pp. 600-603. (USSR)

ABSTRACT: All methods previously proposed for utilizing the reaction of the formation of mixed calcium ferrocyanides have lacked a proper experimental foundation, and thus lead only to qualitative results. The provision of experimental data sufficiently accurate and full to enable quantitative results to be obtained is the object of the present investigation. Solubility methods were used to study the systems in the MC1- Ca2 [Fe(CN)6] -H20, where M = Rb+ or Cs+, which enabled the influence of excess of the components on the composition of the compound formed to be followed. method used has been previously described7. For the Card 1/3 rubidium system it was found that the ratio of

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78-3-15/35

Mixed Ferrocyanides of Calcium with Rubidium and Caesium.

 $[Fe(CN)_6]^{4-}$: Rb+ and of $[Fe(CN)_6]^{4-}$: Ca²⁺ do not depend on the ratio of either component in the mixture, that is the precipitate obtained is constant in composition. The mixed ferrocyanide obtained is represented by the formula Rb2Ca[Fe(CN)6]. Similar results were obtained for the caesium system, the precipitate here being Cs2Ca [Fe(CN)6] . Although results obtained suggested that the formation of the mixed ferrocyanides of calcium could be used for the gravimetric determination of rubidium and caesium, it was found that this could only be done with perfectly pure solutions, since the solubility of the precipitates formed is very strongly affected by the ionic strength of the solution; further, the small difference in the solubilities of the rubidium and caesium salts makes it unlikely that the reaction could be used for separating rubidium and caesium. There are 2 figures, 2 tables, Card 2/3 and 8 references 2 of which are Slavic.

Mixed Ferrocyanides of Calcium with Rubidium and Caesium.

ASSOCIATION: Institute of General and Inorganic Chemistry imeni
N. S. Kurnakov: of the Academy of Sciences of the USSR.
(Institut obshchey i neorganicheskoy khimii im. N.S.
Kurnakova AN SSSR.)

SUBMITTED: October 27, 1956.

AVAILABLE: Library of Congress.

Card 3/3

CIA-RDP86-00513R001754820007-7 "APPROVED FOR RELEASE: 07/13/2001

USSR/Physical Chemistry - Solutions, Theory of Acids and Bases. TANANAYEVIIIV Abs Jour: Referet. Zhurnal Khimiya, No 2, 1958, 3945. : Interaction of Aqueous Be(NO3)2 Solutions with Some Organic Acids. Author . I.V. Tananayev, Ye. P. Shcheglova. Orig Pub: Zh. neorgan. khimii, 1957, 2, No 6, 1365-1373. Abstract: pH of Be(NO₃)2 (I) solution mixtures with organic acids were mea-sured at a constant summary and variable ratios II: acid.

The character of H concentration divergence from additivity in the character of H concentration divergence. sured at a constant summary molarity and variable ratios I: acid.

The character of Hr concentration of Bed ions in an acid medium vith dicates that the interaction of Bed ions and does not proceed for the acids under study is insignificant and does not proceed the acids under study is insignificant. Inst Title dicates that the interaction of Beer ions in an acid medium with fur-the acids under study is insignificant and does not proceed fur-the than the formation of neutral or acid salts. the acids under study is insignificant and does not proceed furinterestion of neutral or acid salts. The interestion of neutral or scid salts. Series:

ther than the formation of neutral the following acid series:

therefore weaker in accordance maleic-citric > tertaric > succinic > tertaric > tertaric > succinic > tertaric > tertaric > succinic > tertaric > lic = melonic > salicylic > meleic-citric > terteric > succinic > cac204 the solubility of cac204 the solubility of cac204 > lactic. : 1/2 Card

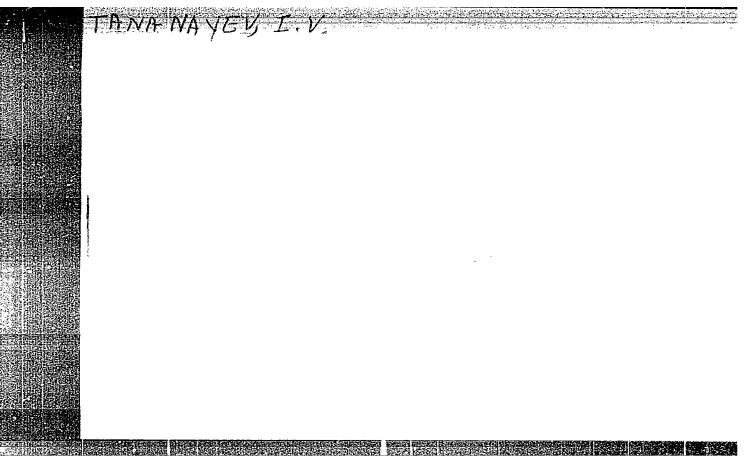
APPROVED FOR RELEASE: 07/13/2001

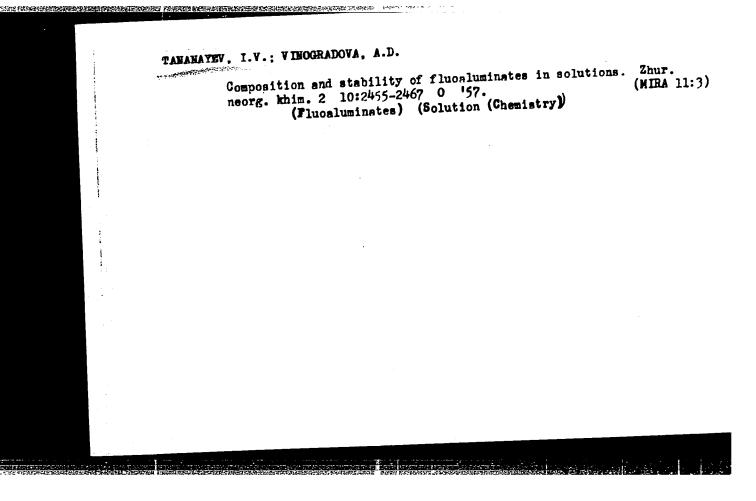
USSR/Physical Chemistry - Solutions, Theory of Acids and Bases. Abs Jour: Referat. Zhurnal Khimiya, No 2, 1958, 3945.

B-11

in I solutions and Al(NO₃)₂ solutions can be used for determination of Al in presence of Be.

Card : 2/2





TANABAYEV, I.V.; IOHOVA, Ye.A.

Zirconium ferrocyanides. Part 1: Interaction of ZrOCl₂ with

Neu[Fe(CH)6] and Nu[Fe(CH)6] in squeous solutions. Zhur. neorg.

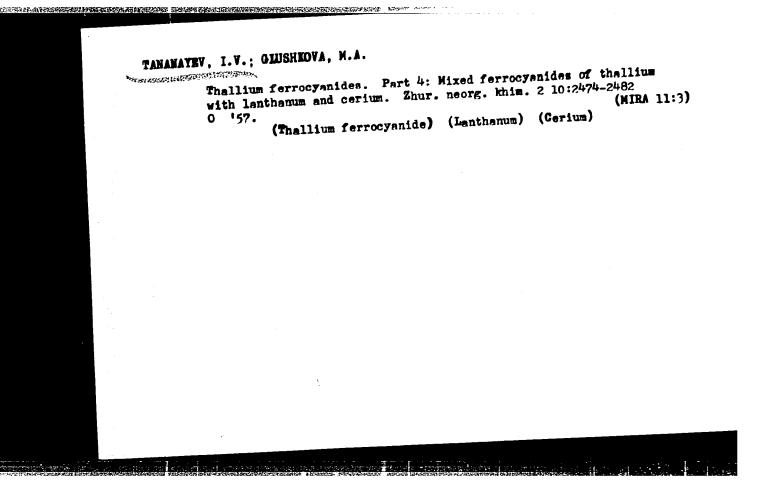
(MIRA 11:3)

khim. 2 10:2468-2474 0 57.

1. Institut obshchey i neorganicheskoy khimii im. H.S.Kurnakova
AN SSSR.

(Zirconium compounds) (Sodium ferrocyanide)

(Potassium ferrocyanide)



TANANAYHY I.V. BOKMEL'DER, M.Ya.

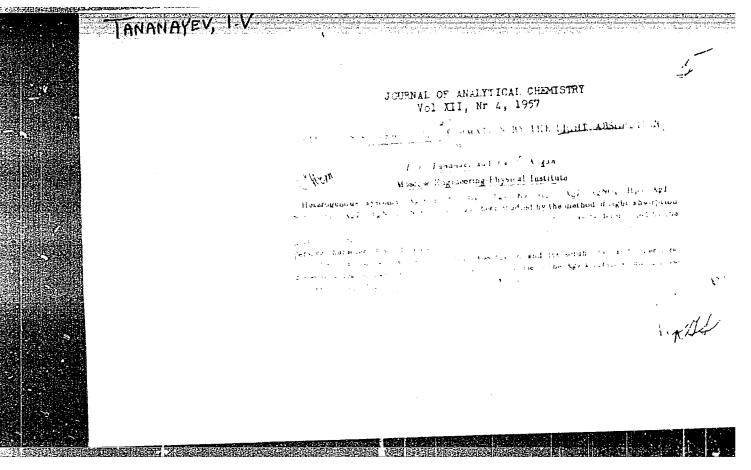
之中,但是一个人,但是一个人,也可以是一个人,也可以是一个人,也可以是一个人,也可以是一个人,也可以是一个人,也可以是一个人,也可以是一个人,也可以是一个人,也

Investigating the formation of nickel hydroxide in aqueous solutions.

Zhur. neorg. khim. 2 no.12:2700-2708 D 157. (MIRA 11:2)

1. Moskovskiy inzhenerno-fizicheskiy institut, Kafedra khimii.
(Nickel hydroxide)

"APPROVED FOR RELEASE: 07/13/2001 CIA-RDP86-00513R001754820007-7



CIA-RDP86-00513R001754820007-7 "APPROVED FOR RELEASE: 07/13/2001

ALTER CONTRACTOR OF THE PROPERTY OF THE PROPER

75-6-1/23 Tananayev, I. V., Shcheglova, Ye. P. AUTFORS: Investigations of the System BeSO 1-NaOH-H2O and its Analytical TITLE: Application (Issledovaniye sistemy BeSO₁₁-NaOH-H₂O i yeye analiti= cheskoye primeneniya). Zhurnal Analiticheskoy Khimii, 1957, Vol. 12, Nr 6, pp. 671-676 PERIODICAL: (USSR). The solubility in the system BeSO_L-NaOH-H₂O at 25°C was investiga= ABSTRACT: ted. It was stated hereby that with the formation of Be(OH)2 first, soluble basic salts are fermed and then follows the formation of insoluble basic salts of variable composition. Beryllium hydroxide dissolves in alkaline lye-solution when the content of alkaline lye attains 0,025 mol per liter. A combined alkali-oxy-quinolate method is preferred for the determination of beryllium in the presence of larger quantities of aluminum. The complete separation of beryllium and aluminum on the strength of the different solubility of their hydroxides in alkaline lyes does not lead to a quantitative separation of aluminum and beryllium. There are 1 figure, 3 tables, and 8 references, 4 of which are Sla Card 1/2

"APPROVED FOR RELEASE: 07/13/2001 CIA-RDP86-00513R001754820007-7

Investigations of the System BeSO $_{L}$ -NaOH-H $_{2}$ O and its Analytical 75-6-1/23 Application.

The second secon

ASSOCIATION: Moseow Institute of Engineering Physics (Moskovskiy inshenerno-fizicheskiy institut).

SUBMITTED: January 4, 1957.

AVAILABLE: Library of Congress.

1. Chemistry-USSR 2. Beryllium hydroxide-Solubility

3. Alkaline lye solution-Applications 4. Aluminum-Beryllium content-Determination

Card 2/2

"APPROVED FOR RELEASE: 07/13/2001 CIA-RDP86-00513R001754820007-7

5(4)

AUTHORS: Vasil'yev, V. P., Korableva, V. D., SOV/153-58-3-30/30 Yatsimirskiy, K. B.

TITLE: Conference Discussion on the Methods of Investigating the Complex Formation in Solutions (Soveshchaniye-diskussiya

po metodam izucheniya komplekscobrazovaniya v rastvorakh)

A STATE OF THE PROPERTY OF THE

PERIODICAL: Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i

khimicheekaya tekhnologiya, 1958, Nr 3, pp 173 - 174 (USSR) ABSTRACT:

From February 18 to 21, 1958 a conference discussion took place at the town of Ivanovo; it dealt with the subjects mentioned in the title. It was called on a decision of the VIIth All-Union Conference on the Chemistry of Complex Formations. More than 200 persons attended the conference, among them 103 delegates from various towns of the USSR. At the conference methods of determining the composition of the complexes in solutions were discussed, as well as the methods of calculating the instability constants according to experimental data and problems concerning the influence of the solvent upon the processes of complex formation.

I. I. Chernyayay Mamban Addison.

I. I. Chernyayev, Member, Academy of Sciences, USSR, stressed Card 1/16 in his inaugural lecture the great importance and actuality

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of the problems to be dealt with, and wished the conference all the best in its work. I. V. Tananayev, on behalf of the Orgkomitet (Organization Committee) held a lecture on: "The Method of Determining the Composition of Compounds Formed in Solutions". In his lecture, V. N. Tolmachev dealt with the problem of the graphical interpretation of the method by Ostromyslenskiy-Zhob. It was proved that this method can also be used in such cases where the equilibrium of complex formation was turned complex by the hydrolysis or dimerization of the central ion. In the lecture by A. K. Babko and M. M. Tananayko, "Physical and Chemical Analysis of the Systems With 3 Colored Complexes in the Solution", the results of a systematic investigation in copper-quinoline-salicylate, as well as in copper-pyridine-salicylate systems by means of the optical method were dealt with. In the lecture by Ya. A. Fialkov the idea of a further investigation of the complex formation processes in solutions was developed. Besides the determination of the composition and stability of the complexes also the physical and chemical properties, the chemical nature and the structure of the complex compounds must be investi-

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gated. The lecture by K. B. Yatsimirskiy dealt with the conditions of checking the usefulness of the method of isomolar series in the determination of the complex composition. be able to obtain objective results the position of the maximum at various concentrations of the components must be checked. A. K. Babko made several critical remarks concerning the lecture by I. V. Tananayev. He pointed out that such a method of investigation must be chosen that is connected with the characteristic properties of the system investigated. A. P. Komar' mentioned in his lecture that for the time being the method by Ostromyslenskiy-Zhob is the best for determining the complex composition, and should be employed as often as possible. This demands, however, that all instructions concerning this method are strictly obeyed. I. S. Mustafin, L. P. Adamovich and V. I. Kuznetsov took part in the discussion. K. B. Yatsimirskiy proved in his lecture "Hydrolytic Equilibria and the Polymerization in Solutions" that, if the hydrolysis products are polymerized, the "inclusion into the complex" and the "formation function" at a constant pH value are varied with the modification of the total concentration of

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the metal. Therefore all those methods may be employed for investigating the polymerization of this type which make the determination of at least one of the two functions mentioned possible. It was proved that the hypotheses on the existence of complexes of the type "nucleus + chain members" can also be founded from the viewpoint of structural concepts: particles the charge of which does not exceed unity can occur as "chain members". The usefulness of the characterization of areas of existence of polymers by means of surface diagrams: "total concentration of the metal - pH" was proved as well. I. I. Alekseyeva and K. B. Yatsimirskiy in their lecture "Investigation of the Polymerization of Iso-Poly Acids in Solutions" mentioned experimental results of the investigation of the polymerization in solutions of molybdic acid. The authors proved that especially the molybdic acid within a certain range of the pH values and the concentrations exists as a number of compounds that can be expressed by an overall formul MoO4(HMoO4) n-2. In the lecture by N. V. Aksel'rud and V. B. Spivakovskiy investigation results on basic salts taking into

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account the complex formation in solutions by means of the potentiometric method were mentioned for systems with zinc, potentiometric method were mentioned for systems with zinc, cadmium and indium. In the evaluation of their results the authors employed the method of the table difference. The authors employed the method of the table difference. The calculation of the consecutive constants was carried out calculation of the interpolation formula by Newton. M. A. according to the interpolation formula by Newton. M. A. Chepelevtskiy held a lecture on "pH Measurement Method of the Solutions in Cambination With the System Analysis of the Solutions in Cambination With the System Analysis of the Solutions in Cambination With the System Analysis. It

ting Complex Copper Compounds in Saturated Solutions". It was found that the substance at the bottom of the liquid is more basic than the solutions furthermore, the increased acidity of the solution from the viewpoint of the formation of hydroxy-chloro complexes in the solution was explained. It is a substant to the discussion with his lecture; he v. I. Kuznetsov opened the discussion with his lecture; he pointed out the necessity of utilizing the concepts worked pointed out the investigations of the polymerization in organic out in the investigations of polynuclear complexes. A. A. Chemistry in the chemistry of polynuclear complexes. A. A. Chemistry thinks that the new approach of the hydrolysis

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investigation as developed by the Scandinavian school is of high value. He also pointed to the necessity of studying the kinetics of the polymerization process and a quantitative determination of the strength of the polymers. A. K. Babko pointed out that the study of the polymer structure was necessary. N. P. Komar' mentioned in his lecture that the rather widely spread polymerization type according to the scheme "nucleus + chain members" is not obtained in all cases The following scientists took part in the discussion: V. N. Tolmachev, A. V. Ablov, I. S. Mustafin, I. V. Tananayev and K. B. Yatsimirskiy. A. K. Babko then discussed in his lecture "Methods of Determining the Dissociation Constant of the Complex Groups in Solutions" the main principles of determini the instability constants. N. P. Komar' discussed in his lecture "Calculation Methods of the Instability Constants of the Complex Compounds According to Experimental Data" the possibilities of using the known calculation methods of the instability constants for various cases of the complex forma. tion in solution. If several mononuclear complexes are form the displacement method by Abegg and Bodlender (completed by

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A. K. Babko) cannot be recommended for the calculation of the instability constant. The lecturer discussed the dissolution methods of the polynomials proposed by B'yerrum, Leden, Rossoti, Sketchard, Edsolloy and other authors. The constants calculated in this way are not very accurate. It was proved that the method of successive approximations can lead to wrong conclusions as to the chemical processes taking place in the system investigated. The most probable value of the physical constants can be obtained by the method of the least squares. B. V. Ptitsyn, Ye. N. Tekster and L. I. Vinogradova described the determination methods of the instability constants of the oxalate complexes of niobium, uranium and iron which are based on the investigation of the equilibrium displacement of the complex formation by silver ions. N. K. Bol shakova, I. V. Tananayev and G. S. Savchenko held a lecture on "The Role of the Time Factor in the Investig tion of the Complex Formation". In the discussion on the lectures A. A. Grinberg mentioned that due to the slow adjustment of the equilibria the methods discussed of determining the instability constants (palladium and cobalt

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complexes) can often not be employed. A. V. Ablov pointed out the necessity of devising direct methods of proving the existence of intermediate forms in a step-wise complex formation. K. B. Yatsimirskiy mentioned that the instability constants of slowly dissociating complexes can be calculated from thermochemical data. L. P. Adamovich, A. M. Golub among others took part in the discussion on the lectures. A. K. Babko requested inclusion in the next conference on the chemistry of complex compounds a lecture in which various calculation methods of the instability constants should be discussed by the example of actual cases. This should clarify to which divergencies of the values of the constants different methods of evaluating the experimental data can lead. N. P. Komar' stressed that in the determination of the instability constants all chemical equilibria should be taken into account that render complex the complex formation process in the solution, especially the hydrolysis processes of the central ion and the addendum. In the lecture delivered by V. M. Peshkova and A. P. Zozulya "Application of the Distribution Method to the Investigation of the Stability Constants

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of Some Thorium Complex Compounds" results obtained from the experimental investigation of the distribution of thorium compounds in the systems: acetylacetone - benzene - water, and 2-oxy-1,4,-naphthoquinone - chloroform - water were given. From these data the instability constants of the thorium complexes with acetyl-acetone and 2-oxy-1,4-naphthoquinone were calculated. I. V. Tananayev, G. S. Savchenko and Ye. V. Goncharov held a lecture on the application of the solubility method in the determination of the stability of complex compounds in solutions. In this lecture also other methods of investigating complex formation processes in the solution were discussed (pH measurement, measurement of the optical density, as well as of the heat of mixing). B. D. Berezin held a lecture on the "Application of the Solubility Method in Studying the Phthalocyanine Complexes of Metals". He used the determined quantitative characteristics of the reaction of the transition of the phthalocyanides of cobalt, nickel, copper and zinc, as well as of the free phthalocyanine into the sulfuric acid solution for the theoretical reasoning, and as an experimental proof of the existence of

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M-bonds in the complexes investigated. These characteristics also served him as a proof of new electronic formulae of phthalocyanine and its complex derivatives. In the lecture delivered by I. L. Krupatkin on "The Method of the Two Solvents as a Method of Investigating the Formation and Properties of Organic Complexes" it was proved that this method makes it possible to determine the number of complexes formed in the system, their composition and relative stability. V. I. Kuznetsov, A. K. Babko, N. P. Komar', I. S. Mustafin and Ya. I. Tur'yan took part in this discussion. In the lecture delivered by A. A. Grinberg and S. P. Kiseleva on the complex palladium compounds (II) with a coordination number above four it was proved that in the case of a Targe chlorine and bromine ion excess complexes with the coordination number 5 are formed. The instability constants of these complexes were estimated. L. P. Adamovich mentioned a new manipulation in the spectrophotometric investigation of the complex compounds that can be used in systems with the formation (or predomination) of one single complex. This method makes it possible to determine the composition and instability constant

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of the complex. In the lecture delivered by K. B. Yatsimirskiy and V. D. Korableva the application of the theory of crystalline fields for the determination of the composition and structure of the chloride complexes of cobalt, nickel and copper according to the absorption spectra of these complexes was discussed. It was proved that in a hydrochloric acid concentration above 5 mole/liter in the solution there exists an equilibrium between the tetrahedric and octahedric form of the cobalt chloro complexes. Yu. P. Nazarenko proved in his lecture "The Application of Radioactive Isotopes in the Investigation of the Solvation Equilibrium in Solutions of Complex Compounds" the possibility of using data on the isotope exchange to clarify the structure of the complex and mechanism of the hydration processes. V. Klimov mentioned in his lecture the use of radioactive isotopes in the study of tin and antimony complexes in non-aqueous solutions. A. V. Ablov, V. N. Tolmachev, V. I. Kuznetsov and A. M. Golub took part in the discussion of the lectures. The usefulness of employing the theory of the crystalline fields in explaining the results obtained from the absorption spectra of the com-

Card 11/16

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plex compounds was stressed. In the lecture delivered by I. A. Shek on "The Investigation of the Complex Formation by the Method of the Dielectric Permeability and the Polarization" the principles of the methods mentioned were presented. This method was employed for investigating the compounds of the type of the "affiliation" products. The lecture delivered by I. A. Shek and Ye. Ye. Kriss "Employing the Method of the Dielectric Constant for Investigating Complex Compounds of the Type of Crystal Solvates in Solutions" dealt with the investigation of the solvates of lanthanum and cerium chlorides with ketones, as well as with the study of the compounds formed in heterogeneous systems with tributyl phosphate and nitric acid. V. F. Toropova gave in her lecture "The Polarographic Method of Investigating the Complex Formation in Solutions" a survey of the applications of the polarographic method in the study of the complex compounds, and illustrated several fine characteristic features of this method. In the lecture delivered by T. N. Sumarokova "The Cryoscopic Method of Investigating the Complex Formation Reactions" a survey of the possibilities of the cryoscopic method was given, and its

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applicability in the study of several complex compounds of stannic chloride with organic substances was proved. A. M. Golub described the results of his investigations of thiocyanate complexes of several metals. A vivid discussion took place on the lectures held. Ya. A. Fialkov and Yu. Ya. Fialkow considered the cryoscopic method of investigating complex compounds to be of considerable value. K. B. Yatsimirskiy pointed out that the publication of the surveys on individual methods of investigating the complex formation reactions would be desired; this concerns especially the polarographic method. The cryoscopic method should be brought to a level that makes the calculation of the equilibrium constants of the processes to be investigated possible. The problem of the method of evaluating the experimental results becomes more and more important. Many scientists use the instability constants without taking into account the way in which they had been obtained. The calculation methods employed by A. M. Golub are one step back, as compared to those employed at present. In his lecture N. P. Komar pointed out the extremely great importance of the mathematical

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evaluation of the results obtained, as well as of the plotting of curves. A. K. Babko suggested selecting one or two systems that are experimentally well investigated, and to evaluate the results obtained according to different methods so that it is possible to check and evaluate them. Ya. I. Tur'yan took part in the discussion. Ya. A. Fialkov discussed in his lecture "The Effect of the Solvent on the Complex Formation Process as Well as on the State of Equilibrium in the Solutions of Complex Compounds" the influence exerted by the solvents upon the molecular state, upon the solvation of the system components, upon the stabilization of the complexes formed in the system, upon the step-wise dissociation of the complexes and upon a number of other processes. The influence exercised by the dielectric constant upon the complex formation process was discussed. It was concluded that a direct relation does not exist, and that the chemical nature of the solvent must be taken into account. A. V. Ablov and L. V. Nazarova held a lecture on "The Spectroscopic Investigation of Nickel Cobalt 'Pyridinates' in Various Solvents". The instability constants of the complexes were determined and it was proved that the

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Conference Discussion on the Methods of Investigating the Complex Formation in Solutions

stability of the 'pyridinates' is changed in dependence on the solvent. Ya. I. Tur yan in his lecture "The Influence of the Solvent Upon the Composition and Stability of Complex Ions" discussed the polarographic investigation method of the chloride and thiocyanate complexes of lead in aqueous ethanol solutions at different content of the non-aqueous solvent and at a constant ionic strength. A step-wise character of the complex formation was found as well as the instability constants of the complexes. The influence of the dielectric constant of the solution on the stability of the investigated complexes was proved. In the lecture by V. P. Vasil yew on the "Investigation of Aquo Complexes in Mixed Solvents" the main attention was devoted to the necessity of the qualitative recording of the solvation effects in the complex formation. The applicability of the polarographic method in the determination of the composition and stability of the aquo complexes in mixed solvents was proved and experimental material on the thermodynamics of the dissociation of the cadmium-aquo complexes in aqueous ethanol solutions was mentioned. V. N. Tolmachev, V. I. Kuznetsov

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and I. V. Tananayev stressed in their lectures the necessity of a more complete and general investigation of the solvation processes. A. K. Babko and A. M. Golub pointed out the processes. A. K. Babko and A. M. Golub pointed out the great importance of the investigations of the complex formation equilibria in non-aqueous solutions, and made several tion equilibria in non-aqueous solutions, and made several critical comments on the lecture by Ya. I. Tur'yan. The critical comments took part in this discussion: L. P. following scientists took part in this discussion: L. P. Adamovich, O. I. Khotsyanovskiy, A. P. Moskvin and A. G. Mustakhov. At the final meeting of the conference A. A. Mustakhov. At the final meeting of the conference A. A. Grinberg, Corresponding Member, AS USSR, said in his speech that such a conference was very urgent. A detailed discustant such a conference was very urgent. A detailed discussion of the determination methods of the composition of the sion of the determination methods of the composition of the complexes, as well as of the method used in the study of the complexes, as well as of the method used in the study of the complexes, as well as of the stepwise complex formation was extremely useful for all who attended this conference, tion was extremely useful for all who attended this conference.

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TANANAYEV, I,V

Tananayev, I. V., Petushkova, S. M.,

78-3-5-2/39

AUTHORS: Shpineva, G. V.

On the Preparation of Water-Free Lithium Iodide (0 TITLE:

poluchenii bezvodnogo yodistogo litiya)

Zhurnal Neorganicheskoy Khimii, 1958, Vol 3, Nr 5, PERIODICAL:

pp 1071-1074 (USSR)

Various methods for the preparation of lithium iodide were tested and are here described: ABSTRACT:

1. Preparation of lithium iodide by the application of

2. Immediate interaction between lithium and iodine, 3. Dehydration of lithium iodide in a HJ-current at 300°C,

4. Dehydration of lithium iodide melts in a vacuum. All the above-mentioned methods gave unsatisfactory results.

In a specially constructed vacuum destillation apparatus, water-free lithium iodide was produced by watery salts in a vacuum at 800-850°C and at a pressure of 0,01 Hg. The water-free lithium iodide is highly hygro-

scopic and decomposes under the influence of light.

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On the Preparation of Water-Free Lithium Iodide

78-3-5-2/39

There are 1 figure and 5 references, 5 of which are Soviet.

SUBMITTED:

May 21, 1957

AVAILABLE:

Library of Congress

1. Lithium iodide-Fraparation-Test results

card 2/2

78-3-6-1/30 Tananayev, I. V., Bokmel'der, M. Ya.

Investigations of the Reaction of the Production of Zirconium AUTHORS: Hydroxide (Issledovaniye reaktsii obrazovaniya gidrockisi TITLE:

tsirkoniya)

Zhurnal Neorganicheskoy Khimii, 1958, Vol. 3, Nr 6,

PERIODICAL: pp. 1273 - 1280 (USSR)

In the present paper the process of formation of zirconium ABSTRACT:

hydroxide in the interaction between diluted solutions of zirconium sulfate and zirconium oxichloride with soda lye is investigated. The precipitation of zirconium hydroxide was performed by means of the physico-chemical analyses, especially

by the methods of the solubility determination, the p_{H} -determination, the determination of the electric conductivity and of

the volume of the zirconium hydroxide precipitations. It was found in the investigations of the system $\mathrm{Zr}(\mathrm{SO_4})_2$ -NaOH-H₂O

that the reaction of zirconium sulfate with soda lye takes place

in three stages. First, soluble, basic salt 3Zr(SO₄)₂.Zr(OH)₄

forms, then the basic salt of the composition (Zr0.0H)SO4

Card · 1/2

78-3-6-1/30

Investigations of the Reaction of the Production of Zirconium Hydroxide

precipitates which at further addition of alkali hydroxide

passes into pure zirconium hydroxide. In the system ZrOCl2-NaOH-H2O the formation of zirconium hydroxide equally

takes place in three stages. First, soluble basic salt ZrO(OH)Cl forms, then the precipitation of the unstable basic salt

7ZrO(OH2)ZrOCl2 precipitates which furthermore passes into the

hydroxide form under the action of NaOH. The zirconium hydroxide precipitations adsorb considerable quantities of alkalies from the solution, which fact corresponds to the amphoteric character of zirconium hydroxide. There are 12 figures, 2 tables and 21

references, 5 of which are Soviet.

May 21, 1957 SUBMITTED:

Library of Congress AVAILABLE:

2. Chemical reactions -- Analysis 1. Zirconium hydroxide--Production card 2/2

Tananayev, I. V., Levina, M. I.

SOV/78-3-9-8/38

AUTHORS:

TITLE:

On Uranyl Ferrocyanides (O ferrotsianidakh uranila)

PERIODICAL:

Zhurnal neorganicheskoy khimii, 1958, Vol 3, Nr 9, pp 2045-2052

(USSR)

ABSTRACT:

The interaction in an aqueous medium in the system ${\rm UO_2(NO_3)_2}$ $M_4[Fe(CN)_6] - H_2O$ was investigated, where M denotes Li, Na, K, Rb and Cs. The investigations were carried out by determining the solubility, by potentiometric determination, and by recording the absorption spectrum. In the system UO2(NO3)2-Li4[Fe(CN)6] H₂O the normal uranyl ferrocyanide (UO₂)₂ Fe(CN)₆ was eliminate as solid phase. Also in the system $UO_2(NO_3)_2-Na_4[Fe(CN)_6]-H_2O$ only normal uranyl ferrocyanide is formed. When adding a surplus of LiR and NaR colloidal solutions are formed. In the systems with M4 [Fe(CN)6], where M denotes K, Rb, Cs, solid phases of the type $M_4(UO_2)_4[Fe(CN)_6]_3$, as well as the solid intermediate phase of the type $M_2(UO_2)_3$ [Fe(CN)6]3 are formed. In the system with

Card 1/2

sov/78-3-9-8/38

On Uranyl Ferrocyanides

 $K_4[Fe(CN)_6]$ in the presence of a surplus of the ion $[Fe(CN)_6]^{4-}$ and K^+ 6-8% K_2SO_4 the compound $K_{12}(UO_2)_8[Fe(CN)_6]_7$ is formed. There are 4 figures, 4 tables, and 12 references, 7 of which are Soviet.

SUBMITTED:

November 29, 1957

Card 2/2

SOV/78-3-9-26/38 Tananayev, I. Y., Avduyevskaya, K. A. AUTHORS:

The Interaction in the System GeO_2 -HF-H $_2O$ at a Temperature of TITLE:

25°C (O vzaimodeystvii v sisteme GeO2-HF-H2O pri 25°)

Zhurnal neorganicheskoy khimii, 1958, Vol 3, Nr 9, pp 2165-2171 PERIODICAL:

(USSR)

The forms and conditions of production of fluorine compounds ABSTRACT:

of germanium in the system GeO2-HF-H2O were analyzed with

physico-chemical methods. The methods used were the determinatio of solubility and electric conductivity. A diagram of solubility in the system GeO_2 -HF-H₂O was established. In the case of an

HF content of 0-35% there is a linear increase of the solubility

of germanium oxide until a molar ratio of HF: GeO2 = 4 is

reached. In the range of between 35 and 41% HF the solubility curve changes the direction. In this point a transformation of GeF₄•3H₂O into hexafluorine germanic acid - H₂GeF₆.2H₂O -

takes place. By determining the solubility and electric conductivity, as well as the pH-value of the solution it was

Card 1/2

SOV/78-3-9-26/38 The Interaction in the System ${\rm GeO_2-HF-H_2O}$ at a Temperature of 25°C

shown that the compounds of the solid phase also exist in the aqueous medium. In the system GeO2-HF-H2O only H2[GeOF4] and H₂ [GeF₆] are formed. Besides GeO₂ also H₂ [GeOF₄].H₂O and H₂ [GeF₆].2H₂O appear as solid phases in the system. The formation of these complex acids is confirmed by the determination of electric conductivity. There are 5 figures, 6 tables, and 13 references, 3 of which are Soviet.

SUBMITTED:

July 24, 1957

Card 2/2

Tananayev, I. V., Avduyevskaya, K. A. Analysis of the Interaction in the System $\mathrm{GeO_2^{-H_2C_2O_4^{-H_2O}}}$ AUTHORS: at 25°C (Issledovaniye vzaimodeystviya v sisteme GeO2-H2C2C4-TITLE: H₂0 pri 250) Zhurnal neorganicheskoy khimii, 1958, Vol 3, Nr 9, pp 2172-2177

PERIODICAL: In the paper under review the interaction of $\mathrm{GeO_2^{-H}_2C_2O_4}$ in (USSR) ABSTRACT:

aqueous solutions was analyzed by the potentiometric method as well as by determining the electric conductivity and solubility at 25°C. In the interaction of germanium oxide and oxalic acid the concentration of hydrogen ions increases. There is a continuous increase of the concentration of hydrogen ions and conductivity until a ratio of $H_2C_2O_4$: $GeO_2 = 3$: 1 is reached,

sov/78-3-9-27/38

after which it remains constant. The resulting conclusion is arter which it lemains compared with the formula $H_2[Ge(C_2O_4)_3]$ that a complex of germanic acid with the formula $H_2[Ge(C_2O_4)_3]$

exists in solution. The analyses of solubility at 25°C showed that 0,045 mol per liter dissolve in oxalic acid. In a saturate

solution of germanic acid with excess oxalic acid added

Card 1/3

SOV/78-3-9-27/38 Analysis of the Interaction in the System $GeO_2^{+H_2}C_2O_4^{-H_2}O$ at 25°C

germanium oxalic acid crystallizes in colorless needle-shaped crystals. This compound has the following composition: $H_2[Ge(C_2O_4)_3] \cdot 6H_2O$. The result of the analysis was as follows: $H_2[Ge(C_2O_4)_3] \cdot 6H_2O$. The result of the analysis was as follows: germanium oxalic acid exists as an ion only $[Ge(C_2O_4)_3]^{2-}$. A germanium oxalic acid exists as an ion only $[Ge(C_2O_4)_3]^{2-}$. A lessening of the acidity of the solution gives rise to ions with a lower ratio of $(C_2O_4)^{2-}$: GeO_2 . Thermograms of a complex of germanium oxalic acid were recorded. At 43°C an endothermic of germanium oxalic acid were recorded. At 43°C an endothermic effect occurs, which corresponds to the melting point of this compound. Ammonium and potassium salts of germanium oxalic acid were produced: $(NH_4)_2[Ge(OH)_2(C_2O_4)_2] \cdot 3H_2O$ and $K_2[Ge(OH)_2(C_2O_4)_2] \cdot 3H_2O$. There are 5 figures, 3 tables, and 10 references, 1 of which is Soviet.

SUBMITTED:

April 7, 1958

Card 2/3

Tananayev, I. V., Kogan, Ya. L.

The Investigation of the Formation Reaction of Silver AUTHORS: TITLE:

Chloride and Silver Bromide by Means of the Method of Light Absorption (Issledovaniye reaktsiy obrazovaniya khlorida

75-1-2/26

i bromida serebra metodom svetopogasheniya)

Zhurnal Analiticheskoy Khimii, 1958, Vol. 13, Hr 1,

PERIODICAL: _pp. 11-17 (USSR)

In an earlier paper (ref. 1) different systems with silver iodide were investigated by measuring light absorption. ABSTRACT:

It was shown that a connection exists between the solubility of a suspension of AgJ and its light absorption. In the present article the results of the analogous investigations

of systems with silver bromide and silver chloride are

described. The composition of the precipitates in the systems

 ${\rm AgNO_3}$ - ${\rm KBr}$ - ${\rm H_2O}$ and ${\rm AgNO_3}$ - ${\rm KCl}$ - ${\rm H_2O}$ was investigated by

measurement of the light absorption. The clearly marked maximum of the absorption of the developing suspension is

attained at a concentration ratio $AgNO_3$: KX = 1(X = Br, C1). Card 1/5

The Investigation of the Formation Reaction of Silver Chloride 75-1-2/26 and Silver Bromide by Means of the Method of Light Absorption

The curve and with it also the degree of dispersion of the precipitate is almost symmetrically modified toward both sides of the end point. Investigations of the systems sides of the end point. Investigations of the systems are increase in the concentration of the halogen ions causes an increase in absorption which is explained by the increase in increase in absorption which is explained by the increase in the solubility of the suspension by the addition of bromide the solubility of the suspension of the results obtained and chloride ions. In a comparison of the results obtained with those already known for the respective systems with with those already known for the curves is remarkable. The

and chloride ions. In a comparison of the results with with those already known for the respective systems with with those already known for the respective systems with with those already known for the respective systems with all, the uniform shape of the curves is remarkable. The absorption maximum in all 3 systems is near similar values absorption maximum in all 3 systems is near similar values absorption of the concentration of the concentration of a suspension of and 2.10-1 Mol/liter. The absorption of a suspension of and 2.10-1 Mol/liter, whereas at higher concentrations of 10-3 to 10-2 Mol/liter, whereas at higher concentrations of 10-3 to 10-2 Mol/liter, whereas at higher concentrations of 11 Sht absorption increases rapidly. The rapid increase in light absorption takes place within the same concentration in light absorption takes place within the same concentration range of MaCl where the solubility of silver chloride also rapidly increases. When the quantity of the solid phase rapidly increases. When the quantity of the solid phase rapidly increases, due to the increasing solubility, the becomes somewhat less, due to the increasing solubility, the value of the absorption maximum does not change. The loss value of the absorption maximum does not change.

Card 2/5

The Investigation of the Formation Reaction of Silver Chloride 75-1-2/26 and Silver Bromide by Means of the Method of Light Absorption

of solid phase is compensated by an increase in absorption based on an increase in particles. The investigation of the systems AgBr - NH3 - H2O showed that after the addition of systems AgBr - NH3 - H2O showed that after the addition of very little ammonia to a suspension of silver bromide a very rapid increase in light absorption takes place. A very rapid increase in light absorption takes place. A comparison with the system AgJ - NH3 - H2O showed great comparison with the system AgJ - NH3 different solubility differences which are due to the highly different solubility of AgJ and AgBr in NH3.

of AgJ and AgBr in Ang.

The system AgBr - AgNC3 - NH3 - NH3 - NH20 was investigated in order to determine the influence of excess silver ions in the presence of ammonia upon the absorption of a suspension of silver bromide. It became evident that the nature of the of silver bromide. It became evident that the system without curves of light absorption as against the system without excess AgNO3 does not change. In both cases the extinction excess AgNO3 does not change oncentration of ammonia, runs increases with increasing concentration of ammonia, runs through a maximum, and then decreases to 0, corresponding to through a maximum, and then decreases to 0, corresponding to a complete dissolution of AgBr. The higher the excess of a complete dissolution of the maximum widen. Thereagno, the more does the domain of the precipitate is only fore; the complete dissolution of the precipitate is only attained at considerably higher concentrations of ammonia.

card 3/5

The Investigation of the Formation Reaction of Silver Chloride 75-1-2/26 and Silver Bromide by Means of the Method of Light Absorption

AVAILABLE: Libra

Library of Congress

Silver chloride - Chemical reactions 2. Silver bromide - Chemical reactions 3. Light - Absorption - Measurement

Card 5/5

75-13-2-7/27 Deychman, E. N., Tananayev, I. V. AUTHORS:

Determination of Small Quantitites of Indium by Titrime-TITLE:

tric and Photometric Methods (Opredeleniye malykh kolichestv indiya titrimetricheskim i fotometricheskim meto-

dom)

Zhurnal Analiticheskoy Khimii, 1958, Vol. 13, Nr 2, PERIODICAL:

pp. 196-200 (USSR)

One of the best known titrimetric methods for the determi= nation of indium is based upon the potentiometric titration ABSTRACT:

by means of potassium ferrocyanide (Ref. 1). The composition KIn₅[Fe(CN)₆] is ascribed to the compound formed here. Kol'=

tgof (Ref. 2) says, however, that the composition of the precipitating deposit was not yet investigated. In the investigation of the systems of indium chloride and of the ferrocyanides of the alkaline metals (Refs. 3,4) it was found that the following order applies for the tendency to form mixed ferrocyanides with indium difficult to be solved: Li (Na (K (Rb (Cs. On the strength of the investiga=

tion of the authors it is very probable that the inaccu-

Card 1/4

Determination of Small Quantities of Indium by Titrimetric and Photometric Methods

75-13-2-7/27

racy of the mentioned potentiometric method is due to the variable composition of the formed precipitation, since in the case of a ratio of $K_{1}[Fe(CN)_{6}]$: InCl₃=0.75 in the ini= tial mixture a normal ferrocyanide of indium precipitates which then continues to react with K4[Fe(CN)6] under for= mation of the mixed salt KIn[Fe(CN)6], In consequence of this reaction the modification of the potential is not ob= vious enough. A uniform compound of a certain composition is formed only in the reaction of indium ions with the ferrocyanides of lithium and sodium. The point of equivalence in the potentiometric titration lies in this case at a ratio of the components which corresponds to the formation of $In_4[Fe(CN)_6]$, Hence follows that the analysis tical determination of indium the ferrocyanides of li= thium and sodium are to be preferred to the ferrocyani= des of potassium, rubidium, and cesium. A method for the determination of indium by means of potentiometric titra= tion with sodium ferrocyanide was worked out. Diphenylamine

Card 2/4

Determination of Small Landities of Indian by Titrimetric and Photometric Methods

75-13-2-7/27

was used here as indicator, Diphenylamine is oxydized in acid solutions under the influence of oxydizing agents to colorless phenyl beamidine and then to violet diphenyl= benzieine (Ref. 5). In to a great extent acid solutions a part of the diphenylbenzidine can be oxidized, before all diphenylamine is transformed into diphenylbenzidine. It was found that sulphuric acid is best suited for the titration of indium. In a solution of 5% of 112504 a stable and sufficiently intensive coloration is formed after 2 - 3 minutes. The best results were obtained in a concentration of 0.02-0.005 g indium in a solution of 50 ml. Small quantities of chlorides and sulfates do not disturb the determination, the coloration of the indicator is, however, formed much more slowly. Oxalic acid reduces to a small extent the coloration, phosphoric acid disturbs. Disturbing cations are chromium, tungsten,

copper and great quantities of iron. In presence of diphenylamine also zinc, cobalt, aluminum and tin show colorations, the disturbing influence of these elements

Card 3/4

Determination of Small quantities of Indium by Titrimetric and Photometric Methods

的时间,我们就是我们的一个时间,我们就是我们的,我们就是我们的一个时间,我们就是这个时间的,我们就会没有一个一个一个一个一个一个一个一个一个一个一个一个一个一个

75-13-2-7/27

can, however, be eliminated by addition of citric acid. Furthermore the authors worked out also a photometric method of determining small quantities of indium. In the case of adding a solution of potassium ferricyanide and diphenylamine to a diluted solution of an indium salt a coloration is produced the intensity of which is proportional to the concentration of indium. A solution of sulfuric acid of 5% is best suited as medium. The smallest quantity of indium which can be determined this way amounts to 4.10-5 g in a solution of 25 nl. Be, Ca, Hg, La, Cd, Ga, Ti, U, Th, Al, Co. Ni. In and In do not disturb the de= termination. In presence of zinc the coloration is produced sooner and to a greater degree. it corresponds, however, after 10 minutes to the normal coloration of the solution without foreign ions. Cr. Ni. Cu. Fe and Tl disturb. There are 7 tables and 5 references, 3 of which are Soviet. Institut obshchey i neorganicheskoy khimii im. N. S. Kur=

ASSOCIATION:

nakova AN SSSR, Moskva (Moscow Institute of General and Inorganic Chemistry imeni N.S. Kurnahov, AS USSR)

SUPPLEMENT:

2. Indium--Volumetric December 29, 1956 1. Indium-Determination

analysis 3. Photometry 4. Diphenylamine-Applications

Card 4/4

"APPROVED FOR RELEASE: 07/13/2001 CIA-RDP86-00513R001754820007-7 THE REPORT OF THE PROPERTY OF

SOV/75-13-4-6/29 Ponomarev, V. D., Tananayev, I. V. AUTHORS:

Investigation of the Formation Reaction of Mixed Ferrocyanides of Copper and Alkali Metals by a Potentiometric Method TITLE: (Izucheniye reaktsii obrazovaniya smeshannykh ferrotsianidov

medi i shchelochnykh metallov potentsiometricheskim metodom)

Zhurnal analiticheskoy khimii, 1958, Vol. 13, Nr 4, pp. 417-PERIODICAL:

422 (USSR)

The potentiometric titration of ferrocyanides, which nowadays is applied with varying success in the determination of a whole ABSTRACT:

series of metals (Ref 1), has one shortcoming. The position of the end point frequently is dependent on the kind of the alkal: metal which forms the cation of the ferrocyanide and also on the concentration of alkali metals in the solution. Knowledge of the mechanism of the formation reaction of an inscluble mixed ferrocyanide would disclose new possibilities of the analytical application of ferrocycnides. For this purpose it is indispensable to investigate one relation between the com-

position of the mixed ferrocyanides and the presence of one or the other alkali metal in the solution, and its concentration.

The present paper deals with potentiometric titration of corpe Card 1/3

sov/75-13-4-5/29

Investigation of the Formation Reaction of Mixed Ferrocyanides of Copper and Alkali Metals by a Potentiometric Method

ions by $[Fe(CN)_6]^{4-}$. Despite a series of published articles on this problem (Refs 3, 4), no satisfactory potentiometric method of titrating copper by ferrocyanides has hitherto been known. In all of these investigations it is titrated by $K_4[Fe(CN)_6]$, however, the possible influence of other alkali metals being present on the titration was not investigated. In some of the papers reference is made to the lowered solubility of the mixed ferrocyanides of copper in dependence on the ion radius of the alkali metal which is contained in the corresponding sediment (Refs 5-7). The authors of the present paper investigated the system consisting of the ions Cu^{2+} and $Me_4[Fe(CN)_6]$ (Me = Li, Na, K, Rb, Cs). The redox-potentials were determined on a potentiometer of the type JII - 6. It was found that common $Cu_2[Fe(CN)_6]$ is produced in the titration by ferrocyanides of lithium and sodium. In presence of ions of the remaining alkali metals, however, mixed ferrocyanides develop: $K_4Cu_{10}[Fe(CN)_6]_6$, $Rb_6Cu_4[Fe(CN)_6]_3$, and $Cs_2Cu[Fe(CN)_6]_6$. Among

card 2/3

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SOV/75-13-4-6/29

·Investigation of the Formation Reaction of Mixed Ferrocyanides of Copper and Alkali Metals by a Potentiometric Method

the potentiometric methods of titrating copper the titration by rubidium ferrocyanide or by the ferrocyanides of the other alkali metals in presence of rubidium salts is the most suitable one. The way of carrying out the investigations based on redox-potentials is described in detail. There are 3 figures, 3 tables, and 8 references, 6 of which are Soviet.

ASSOCIATION:

Moskovskiy inzhenerno-fizicheskiy institut (Moscow Engineering Physics Institute)

SUBMITTED: May 2

Мау 29, 1957

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1. Titration—Materials 2. Iron cyanide—Chemical reactions 3. Copper—Chemical reactions 4. Alkali metals—Chemical reactions 5. Metals—Determination

Card 3/3

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TANANNYEV, I.V., akademik, prof.; SHAPIRO, L.M., kand.khim, nauk

Phototurbidimetric titration of cobalt ions with rubeanic acid.

Shor.mauch.rab.Bel.politekh.inst. no.63:159-163

(Cobalt-Analysis) (Oxamide)

5(0)

Tananayev, I. V., Academician

sov/30-59-1-30/57

AUTHOR:

News in Brief (Kratkiye soobshcheniya) Congress of Austrian

TITLE:

Chemists (S"yezd avstriyskikh khimikov)

PERIODICAL:

Vestnik Akademii nauk SSSR, 1959, Nr 1, pp 117 - 118 (USSR)

ABSTRACT:

- n October 15 until 18, 1958 in memory of the 100th Andversary of Auer von Welsbach (Auer fon Vel'sbakh) who is well-known for his research work in the field of the chemistry of rare elements. The symposium which took place in the course of the Congress of Austrian Chemists was devoted to the survey of achievements in the field of division of rare earth elements as well as to the research of their properties and the viscous elements and compounds. Apart from the reports delivered by western scientists only L. Wolf and I. Massonn (both from the German Democratic Republic) gave a report. Wolf dealt with problems of complex formation and their utilization in the emanation of rare earths and Masson reported on the separation of large quantities of rare earths by cationites with trinitryl acetic acid serving as extractor.

Card 1/1

"APPROVED FOR RELEASE: 07/13/2001 CIA-RDP86-00513R001754820007-7

SOV/78-4-1-20/48 5(2) Tananayev, I. V., Lyutaya, M. D. . AUTHORS: I. On the Mixed Hexanitritonickelates of Lanthanum and Potassium (I. O smeshannykh geksanitronikeleatakh lantana i TITLE: kaliya) Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 1, pp 97-102 PERIODICAL: (USSR) Solubility in the system $La(NO_3)_3-K_4[Ni(NO_2)_6]^{-H_2O}$ was ABSTRACT: investigated at 25°. The solid phases separated out were analyzed and the thermograms of these compounds drawn. The solubility curves indicate the gradual formation of three solid phases with a rise of K4 Ni(NO2)6 content. The following solid phases are formed: $K_6La_2[Ni(NO_2)_6]_3$; $K_{21}La_5[Ni(NO_2)_6]_9 \cdot H_2O$; $K_5La[Ni(NO_2)_6]_2 \cdot H_2O$. The individuality of these compounds was proved by their thermograms. The thermograms of $K_6La_2[Ni(NO_2)_6]_3$ show an endothermic effect within the temperature response. within the temperature range 220-265°. Thereby the color of the salts changes from brown to black. The thermograms of Card 1/3

CIA-RDP86-00513R001754820007-7"

APPROVED FOR RELEASE: 07/13/2001

SOV/78-4-1-20/48

I. On the Mixed Hexanitritonickelates of Lanthanum and Potassium

 $K_{21}La_5 Ni(NO_2)_6 9 \cdot H_2O$ show two endothermic effects. The first effect at 130° indicates the dehydration of the salt. The second one at 230-270° indicates the decomposition of the salt. The thermogram of $K_5La Ni(NO_2)_6 2 \cdot H_2O$ shows an endothermic effect at 130° indicating the dehydration of the salt and an endothermic effect at 230-250° indicating the decomposition of the salt. The solubility of $K_5La Ni(NO_2)_6 2 \cdot H_2O$

in KNO $_2$ solutions (1-7 mol/l) was investigated. It was found that at the same time salting out takes place whereby the solid initial phase is changed to K_6La_2 Ni(NO $_2$)6 3. Rare earths can be separated by KNO $_2$ solutions by fractional crystallization of their mixed hexanitritonickelates. There are 6 figures, 4 tables, and 2 references.

Card 2/3

5(2) AUTHORS:

Tananayev, I. V., Lyutaya, M. D.

507/78-4-1-21/48

TITLE:

II. On Mixed Hexanitritonickelates of Praseodymium and Neodymium (II. O smeshannykh geksanitronikeleatakh prazeodima

i neodima)

PERIODICAL:

Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 1, pp 103-109

(USSR)

ABSTRACT:

Solubility in the systems $Pr(NO_3)_3 - K_4[Ni(NO_2)_6] - H_2O$ and $Nd(NO_3)_3 - K_4[Ni(NO_2)_6] - H_2O$ was investigated at 25°C. In the first system the phases $K_2 + Pr_5[Ni(NO_2)_6] 9 \cdot H_2O$ and

K5Pr[Ni(NO2)6] 2.H2O are gradually formed. The second system

also shows the gradual formation of two solid phases: $K_{21}^{Nd} = \frac{Ni(NO_2)}{6} = \frac{9 \cdot H_2^{O}}{9 \cdot H_2^{O}}$ and $K_5^{Nd} = \frac{Ni(NO_2)}{6} = \frac{1}{2} \cdot \frac{H_2^{O}}{6}$. The thermograms were drawn and are shown in the figures 3, 4, and 8, 9. The solubility of K_5 Pr $Ni(NO_2)_6$ 2. H_2O and K_5 Nd $Ni(NO_2)_6$ 2. H_2O in

KNO, solutions (1-7 mol/1) was investigated. It was found that

Card 1/2

SOV/78-4-1-21/48

II. On Mixed Hexanitritonickelates of Praseodymium and, Neodymium

the solid phases thereby change to $K_{21}^{Pr} \int_{0}^{Ni(NO_2)} \frac{1}{6} g^{H_2O}$ and $K_{21}^{Nd} \int_{0}^{Ni(NO_2)} \frac{1}{6} g^{H_2O}$. There are 10 figures, 8 tables, and 1 Soviet reference.

SUBMITTED:

August 2, 1958

Card 2/2