TEŽAK, B.

A comparison of the coagulation phenomena of dialyzed and freshly prepared silver iodide sols. F. P. Schulz and B. Težak (Univ. Zagreb, Yugoslavia). *Arhiv Kem.* 26, 187–90(1954)(in English).—The coagulation values of K, Ba, and La nitrates for dialyzed and freshly prepd. AgI sols in the presence of various amts. of stabilizing ion (I⁻) were compared. The results of Kruyt and Klompé (*C.A.* 38, 6158²) and Težak, *et al.* (*C.A.* 44, 6324c; 46, 1842f) were used and represented as plots of the coagulation values of neutral electrolytes *vs.* the neg. log values of the excess of KI (pI). The pI values in Kruyt and Klompé's measurements were 3.07, 4.00, and 5.08, and in Težak's results they varied from ~2 to ~4; this shows that in both cases these results were practically equal. The observed differences in coagulation values are due to different criteria used in their detns. (18 hrs. for dialyzed sols and 10 min. in the case of freshly prepd. sols). These differences are small because after 10 min. the coagulation rate of AgI sols is rather low. The agreement between the results obtained with sols of different origin was considered as a further proof that, in the case of sols in the nascent state, pure coagulation phenomena were studied, which were not essentially affected by the presence of smaller or greater amts. of foreign ions.

J. Kratohvil

CH

①

TEZAK, B.

✓ The complex solubility and the composition of aqueous complex solutions of silver halides and silver thiocyanate. J. Kratohvil, B. Težak, and V. B. Vouk (Univ. Zagreb, Yugoslavia). ~~Arhiv Kem.~~ 26, 191-200(1954)(in English).—A systematization and comparison of the numerous data on complex soly. of Ag halides and AgSCN obtained by various techniques was made. The results are represented as plots of the log of the total concn. of Ag ions *vs.* the log of the total concn. of halide or SCN⁻ ions at the soly. boundaries (so-called soly. curves). Satisfactory agreement between the data of various authors was noted. Except in the ionic soly. regions, the effect of the ionic strength was small or completely absent. Complex soly. of AgCl in solns. of various sol. chlorides increases in the order $NH_4Cl > KCl = LiCl > NaCl > HCl$ and $BaCl_2 > SrCl_2 > CaCl_2 > MgCl_2$. The effect of the accompanying cation from chloride component decreases with the increasing diln. of the solns. This was tentatively connected with the steric factors. The complex soly. in the solns. of the corresponding sol. halide or SCN⁻ salts (for amts. of solid phase greater than $10^{-4}M$) changes in the order $AgSCN > AgI > AgBr > AgCl$. Increase of the complex soly. in the solns. of $AgNO_3$ is in the order $AgI > AgSCN > AgBr > AgCl$. A higher concn. of Ag ions is necessary for dissolving the same amt. of solid phase than for halide or SCN ions. In order to evaluate the compn. of the complex solns. and to det. the constitution of the sol. complex species and their stability consts. in the various concn. ranges, a direct graphic method was elaborated.

③

1/2

(OVER)

The soly. curves could be approximated by a series of secants, whose slopes were in all cases only slightly different from small whole nos. (0 to 4). The values of the slopes of the secants represent the no. of ligands, minus one, bound to the central ion in a mononuclear complex. The segments of the secants on the ordinate represent the equil. consts. of soln. of solid phase and the formation consts. of the corresponding complexes, resp. The stability const. of a complex is equal to the ratio of the corresponding equil. const. and the soly. product. At high concns. of the component in excess the formation of polynuclear complexes is assumed. By the procedure described the complex soly. results were interpreted and the stability consts. of the complex ions compared with the conclusions of other authors. Extensive tables including the authors, the references, the concn. ranges of the component in excess, the compns. of the complex ions present, and the corresponding stability consts. are given. Essentially the same results were reached by various authors. The reliability of the interpretation of the soly. data is discussed.

J. Kratohvil

ZAK, B.,

YUGOSLAVIA/Thermodynamics. Thermochemistry. Equilibria. Physico-Chemical Analysis. Phase Transitions B-8

Abs Jour : Ref Zhur - Khimiya, No 8, 1957, 26150

Author : J. Kratohvil, B. Tezak

Title : Methodics of Precipitation Processes. XI. Complex Solubility of Silver Complex Halides and Silver Thiocyanate in Mixed Solvents

Orig Pub : Arhiv kemije, 1954, 26, No 4, 243-256

Abstract : The solubility (L) of AgCl, AgBr, AgI and AgCNS in solutions of halides and thiocyanates of alkali metals in isodielectric mixtures (water - methyl alcohol, water - ethyl alcohol, water - acetone) at 20 ± 0.1° was studied. This process is accompanied by the formation of complexes in the solution. L increases with the increase of the concentration of the organic component, and this increase means that less halide ions containing in the complex are necessary to the solution of the same amount of the solid phase. In mixtures of substances of equal dielectric constants (water - methyl alcohol,

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YUGOSLAVIA/Thermodynamics. Thermochemistry. Equilibria. Physico-Chemical Analysis. Phase Transitions. B-8

Abs Jour : Ref Zhur - Khimiya, No 8, 1957, 26150

water - ethyl alcohol), L increases approximately regularly (compared with L in pure water); water - acetone mixtures, in which L is noticeably greater, is an exception. The difference between L of AgBr in solutions of KBr and NaBr is conserved also at the dissolution in mixed solvents. L of silver halides and thiocyanate decreases in all solvents (including water) in the following order: AgCNS, AgI, AgBr, AgCl. L in 82% acetone is an exception, it decreases in the following order: AgI, AgCNS, AgBr, AgCl. The anomaly of water-acetone mixtures is connected with the change in the ion-dipole interraction (formation of solvates). The less the dielectric constant of the solvent is, the greater is the magnitude of the stability constant of the complex. The dielectric constants and the compositions of mixtures were computed by interpolation basing on data published earlier (Akerlof G., J. Amer. Chem. Soc., 1932, 54, 4125). The values of ion solubility at various magnitudes of the dielectric constants

Card : 2/3

YUGOSLAVIA/Thermodynamics. Thermochemistry. Equilibria. Physico-Chemical Analysis. Phase Transitions. B-8

Abs Jour : Ref Zhur - Khimiya, No 8, 1957, 26150

necessary to the computation of the constants of stability were obtained by means of the relation published earlier (Ricci J.E. and others, J. Amer. Chem. Soc., 1939, 61, 3274; 1940, 62, 407; J. Phys. Chem., 1941, 45, 1096; J. Amer. Chem. Soc., 1942, 64, 2305). The stability constants were computed by the method described earlier (RZhKhim, 1955, 26012). The methodics of the work was published earlier (Schulz K. and others, Arkhiv kem., 1951, 23, 200). See RZh Khim, 1955, 26012 for part X.

Card : 3/3

TEZAK, B.

Adsorption of constituent ions [illegible]

Ag[illegible] that function of [illegible] ions [illegible] the adsorption [illegible] coagulation, with [illegible] adsorption [illegible] were [illegible] effect of [illegible] The lowering of the adsorption capacity. In the stable region (concn. of $NaNO_3$ <0.1N), the adsorption capacities are relatively high (2-3 mols. I^- per thousand mols. AgI), but in the unstable region (concn. of $NaNO_3$ >0.1N) they have low values of about 0.3 to 0.4 mol. I^- per thousand mols. AgI. The possibility is shown that the surface potential is an important factor [illegible]

TEŽAK, B.

Complex precipitation reactions of silver phosphate sols in statu nascendi. E. Matijević and B. Težak. *Kolloid-Z.* 136, 74-80(1954).—By means of the Tyndall beam, the pptn. systems $AgNO_3$-alkali phosphate-KBr were investigated. In the system $AgNO_3$-Na_3PO_4, two tyndallometric max. existed, namely one for Ag oxide and one for Ag phosphate. On adding KBr, a 3rd AgBr max. existed. The influence of added HNO_3 was systematically studied and the corresponding pH values were measured with a glass electrode. The max. for Ag phosphate was always obtained at approx. the same pH. The characteristics of the max. were discussed. By using small amts. of gelatin it was found that the Ag phosphate system is represented by a crystn. max., whereas AgBr is represented by a coagulation max.

E. E. Gaiser

①

Methodics of the coagulation processes. XI. Coagulation effect of potassium succinate, potassium fumarate and potassium maleate on [illegible]

[illegible] is [illegible]. Coagulation values of K succinate and maleate are dependent not only on the concn. of the stabilizing ion but also on [illegible]

TEZAK, B

5

CH

Effect of dilution on cation-exchange equilibrium. E. Reiner, K. F. Schulz, and B. Težak (Univ. Zagreb, Yugoslavia). *Arhiv Kem.* 27, 93-8(1955)(in English).—The following ion exchanges were studied in batch systems on Amberlite IR-120 to see how the distribution of ions between the ion exchanger and the soln. depends on the diln. of the soln.: K^+-H^+, Mg^{++}-Ca^{++}, Ca^{++}-H^+, Ba^{++}-H^+, Ce^{+++}-H^+, and La^{+++}-H^+. One g. of air-dried resin was equilibrated during 24 hrs. with a soln. of HNO_3 for exchanges with H, and with a soln. of $Ca(NO_3)_2$ for Mg-Ca exchange. The concn. of the solns. varied from 0.0002 to 0.01*N*. When 2 batches differed only in the vol. of the aq. soln. (from 10 to 50 ml.) more ions of higher affinity passed into the resin phase in the more dild. soln., and the effect of diln. was nearly negligible for the exchange of ions of equal charge (K^+-H^+, Mg^{++}-Ca^{++}), whereas it was well marked for the exchange of ions of unequal charge. The results obtained are in qual. agreement with Donnan theory and the theory of Jenny and Davis (*C.A.* 40, 2370^4).

J. Krutohvil

AP

② MET

Težak, B.

A note on the preparation and optical properties of monodispersed lead iodate hydrosols. J. Herak, M. M. Herak, B. Težak, and J. Kratohvil (Univ. Zagreb, Yugoslavia). Arkiv Kem. 27, 117-18(1955)(in English).—Monodispersed $Pb(IO_3)_2$ hydrosols were prepd. by slowly pouring the soln. of KIO_3 into an equal vol. of $Pb(NO_3)_2$ soln. Concns. of about 0.001-0.003*N* were most convenient. The growth of $Pb(IO_3)_2$ particles was very fast and after a few min. the sols exhibited Tyndall spectra of the scattered light as brilliant as those of La Mer's sulfur sol (La M. and Barnes, *C.A.* 40, 3038^1) after 24 hrs. A plot of the ratio of the intensity of scattered light at 4360 and 3660 A. for various angles of observation shows characteristic orders of Tyndall spectra. The way of mixing of pptg. components had great influence on the optical phenomena, although Tyndall spectra always appeared. The sols were not stable; after about 15 min. sedimentation took place and the color bands began to disappear. The addn. of gelatin (0.1 to 1%) or filtration through ordinary filter paper immediately after mixing stabilized the sols. The radius of $Pb(IO_3)_2$ particles detd. by the sedimentation method of Johnson and La Mer (*C.A.* 41, 5707*c*) amounted in one case to 8000A. The appearance of Tyndall spectra was also observed on $AgIO_3$ and $La(IO_3)_3$ hydrosols, but to a lesser extent. J. K.

CH

(3)

Težak, B.

✓ Coagulation of hydrophobic sols in statu nascendi. III. The influence of the ionic size and valency of the counterion. B. Težak, E. Matijević, and K. F. Schulz (Univ. Zagreb, ~~Yugoslavia~~). *J. Phys. Chem.* 59, 769-73(1955); cf. *C.A.* 46, 1842e; 47, 6219i.—The coagulation values of different cations for neg. AgBr and AgI sols were detd. There exists a linear relation between the crystallographic radii of the counterions of the same valency and the logs of their coagulation values. By extrapolating these straight lines to the zero value of the radius a characteristic coagulation value was obtained that was connected in a simple manner with the valency of the counterion by means of the equation, $\log C_{coag.} = -(\delta/\Delta)z + \log C_{lim.}$, where $C_{coag.}$ is the coagulation value expressed in normalities, δ is crit. distance (3.58 A. for univalent point ions in water media), Δ is a distance which can be compared with the thickness of the ionic atm., and $C_{lim.}$ denotes the concn. of an uncharged species of particles that would exert the same interaction as the charged ions in the region of the crit. coagulation concn.

Henry Leidheiser, Jr.

②

TEZAK, B.

GERM

Precipitation phenomena of silver thiocyanate. V. Stubičan and B. Težak (Univ. Zagreb, Yugoslavia). Kolloid-Z. 140, 65-[illegible]; cf. C.A. 46, 1812e, 9012g; 47, 6210i.—The reaction of $AgNO_3$ and KCNS was followed turbidimetrically for 60 min. after mixing. A max. pptn. rate during the 1st min. occurs near the equivalence point; some excess of either reactant stabilizes the sol and slows pptn., but a larger excess produces a 2nd max. in pptn. rate, followed by a concn. range over which the product is sol. by complex formation. The stabilizing effect of excess CNS^- is exerted over a more-limited concn. range than that either of the Ag^+ in the present case, or of the halide ion with analogous Ag halide sols. This is a result, not of weak adsorption of CNS^-, but of the large soly. of AgCNS in CNS^-, which allows rapid recrystn. processes, by which the primary particles are linked into aggregates. Accordingly, the ppts. formed with excess Ag^+ are more easily peptized than those formed with excess CNS^-. X-ray diffraction patterns show a uniform cryst. structure for all ppts., with the primary particle size in the colloidal range except for concns. near the region of complex soly., where the size exceeds 10^{-4} mm. With increasing concn. of stabilizing ion, the coagulation value of K^+ for the sol increases, of Ba^{++} remains const., and of La^{+++} decreases (contrary to the Verwey-Overbeek theory), while Th^{++} causes charge reversal.

W. S. Brey, Jr.

TEZAK, B

YUGOSLAVIA/Inorganic Chemistry - Complex Compounds

C.

Abs Jour : Referat Zhur - Khimiya, No2, 1957, 4080

Author : Branica, M., Bona, E., Simunovic, N., Tezak, B.

Title : Extraction of Inorganic Ions with Organic Solvents. I. Continuous Extraction of Uranyl Nitrate with Tetrahydrosylvane and Tetrahydro Pyrane.

Orig Pub : Croat. chem. acta, 1956, 28, No 1, 9-12

Abstract : Tetrahydrosylvane (I) and tetrahydropyrane (II) are much more effective extraction agents for $UO_2(NO_3)_2$ (III) than ethyl ether (IV) and ethyl acetate (V). With a constant concentration of HNO_3 (0.25 N) and III (1 mg U in 12 ml) in the aqueous phase, % of extracted III increases with concentration of NH_4NO_3. The salting-out effect of NH_4NO_3 is least pronounced on using IV and V as extracting agents. The most effective extracting agent is I which removes 100% U from a solution that is 0.8 N in NH_4NO_3. II extracts III completely

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YUGOSLAVIA/Inorganic Chemistry - Complex Compounds

C.

Abs Jour : Referat Zhur - Khimiya, No 2, 1957, 4080

from a solution that is 1.2 N in NH_4NO_3. Extractability of III from 0.8 N NH_4NO_3, containing 3 1 mg U per 12 ml, increases with concentration of HNO_3. However, IV and V do not extract completely the III. 0.012-180 mg U per 12 ml of a solution that is 0.8 N in NH_4NO_3 and 0.1 N in HNC_3, are extracted completely with I in 70 minutes. II extracts 100% III from 1.2 N NH_4NO_3 + 0.25 N HNO_3; for a continuous extraction use is made of the previously described micro-extractor (RZhKhim, 1955, 21563). NH_4Cl, $(NH_4)_2SO_4$ and $(NH_4)_2HPO_4$ (especially the last named) decrease extractability of III on using I or II as extracting agents. In the presence of added $Fe(NO_3)_3$ (0.6 g-equivalent/liter) III can be extracted from solutions containing chloride, sulfate and even the phosphate of ammonium.

Card 2/2

TEZAK, B.

Methorics of coagulation processes. XIII. Systematic comparison of coagulation values for various cations on negative silver halogenide sols. In English.

p. 81 (Croatica Chemica Acta. Vol. 28, no. 2, 1956. Zagreb, Yugoslavia)

Monthly Index of East European Accessions (EEAI) LC. Vol. 7, no. 2, February 1958

[illegible]

Tezak, B.

TEZAK, B

[illegible]

TERPA, B.

[illegible]

TEZAK B.

YUGOSLAVIA/Physics of Solids - Morphology of Crystals. Crystallization E-8

Abs Jour : Ref Zhur - Fizika, No 2, 1958, No 3530

Author : Cernicki B., Tezak, B.
Inst : University of Zagreb, Yugoslavia
Title : Methodics of the Percipitation Processes. XIV. New Maxima on the Percipitation Curves of Mixed Silver Halides.

Orig Pub : Croat. chem. acta, 1957, 29, No 1, 7-13

Abstract : New Maxima were obtained on the tyndallometric curves (TC) of the percipitation of silver halides from solutions of $AgNO_3$ and KCl with various small additions of KI of KBr. The concentration of $AgNO_3$ was 2×10^{-4}N (equivalent concentration), that of KCl was 2×10^{-3}N and 4×10^{-4}N. The concentration of KI or KBr was varied over a wide range, but remained always less than the concentration of the more soluble halogenide (KCl). At a KCl concentration of 4×10^{-4}N, two maxima were observed on the curve, and at 2×10^{-3}N one of the maxima became flat. The course of the TC corresponded essentially to the course of the dispersion curves. Increasing the temperature shifted the position and height of the maxima of

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YUGOSLAVIA/Physics of Solids - Morphology of Crystals. Crystallization E-8

Abs Jour : Ref Zhur - Fizika, No 2, 1958, No 3530

the TC and stopped the sol particles from growing. The new maxima on the TC show that at certain concentrations there are percipitated crystals of mixed silver halogenides. Evidence in favor of this assumption is the above mentioned ratio of the halide ion concentrations, which give a more and less soluble silver halogenide and also the fact that upon contact between two colloidal silver halogenides a redistribution of the components takes place leading to the formation of mixed crystals. The influence of temperature on the speed of percipitation is attributed to its effect on both the speed of crystallization and on the ionic solubility.

Card : 2/2

YUGOSLAVIA/Physical Chemistry - Thermodynamics, Thermochemistry, Equilibria, Physical-Chemical Analysis, Phase Transitions. B.

Abs Jour : Ref Zhur - Khimiya, No 14, 1958, 45954

The methods of work were descirbed formerly (Arhiv, kem., 1951, 23, 200). All the measurements were carried out at 20°. The solubility rise with the complex formation is approximately the same in the cases of isodielectric water-alcohol mixtures independently of the used alcohol. Dioxane produces a somewhat different effect. The obtained data were compared with the data published formerly (report XIV, RZhKhim, 1958, 641).

XVI. The process of Ag, Pb and La iodate precipitation from aqueous solutions of electrolytes was studied and the concentration ranges, in which their separation proceeded in the solid phase, were investigated. The concentrations, at which the precipitation takes place, agree with the data of other authors (RZhKhim, 1955, 26012; 1957, 22335, 50711) only in the case of $AgIO_3$.

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YUGOSLAVIA / Physical Chemistry. Thermodynamics. Thermochemistry. Equilibria. Physicochemical Analysis. Phase Transitions. B-3

Abs Jour: Ref Zhur-Khimiya, No 2, 1959, 3875.

Author : Cernicki, B. and Tezak, B.
Inst : Not given.
Title : Methorics [sic] of the Precipitation Process. XVII. A Study of the Simultaneous Precipitation of Silver Halides with Silver Cyanide and Silver Thiocyanate.

Orig Pub: Croat Chem Acta, 20, No 1, 1-7 (1958) (in English with a Serbo-Croat summary)

Abstract: The simultaneous precipitation of the various silver halides with silver cyanide or with silver thiocyanate has led to the preparation of mixed 'double' systems of crystals: AgCl-AgCN, AgCl-AgCNS (I)

Card 1/2

YUGOSLAVIA / Physical Chemistry. Thermodynamics, Thermochemistry. Equilibria. Physicochemical Analysis. Phase Transitions. B-8

Abs Jour: Ref Zhur-Khimiya, No 2, 1959, 3875.

Abstract: AgCN-I, AgCN-AgBr, AgCN-AgI, I-AgBr, I-AgI; these systems were studied tyndallimetrically. The systems AgCl-I and AgCN-I gave reproducible precipitation curves indicating the formation of mixed crystals. For Communications XV and XVI see RZhKhim, 1958, 45954. -- From a summary by the authors.

Card 2/2

Country : Yugoslavia
Category : Colloid Chemistry. Disperse Systems B-14

Abs. Jour. : Ref Zhur - Khimiya, No 6, 1959 18731

Author : Bifl, M.; Tezak, B.
Institut. :
Title : Methorics of the Coagulation Processes. XIV. Simultaneous Coagulation of the Mixed Systems of Silver Halides, Silver Cyanide and Silver Thio-*
Orig Pub. : Croat. chem. acta, 1958, 30, No 1, 9-14

Abstract : Study of coagulation of mixed sols: AgCl - AgBr, AgCl - AgI, AgCl - AgCN, AgCl - AgCNS and AgBr - AgI, with K-nitrate. Initial $AgNO_3$ concentration was 2 . 10^{-4} N, concentration of ion forming more soluble compound with Ag^+ (Cl^-, Br^-), 2 . 10^{-3} N, concentration of ion forming less soluble compound, 0 - 2 . 10^{-3} N. Turbidity - log c_{KNO_3} curves of systems AgCl-AgBr and AgCl-AgI show that with low concentration of Br^- (respectively, I^-) coagulating concentration c_k almost does not differ from c_k for AgCl sol. When concentration of added Br^- (I^-) ions becomes equal to concentration of Ag^+, c_k changes sharply and approximates c_k for AgBr (AgI) sol. This is due to the fact that on equal

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* -cyanate by Monovalent Counter-Ions.

Country : Yugoslavia

Category : Colloid Chemistry. Disperse Systems. B-14

Abs. Jour. : Ref Zhur-Khimiya, No 6, 1959 18731

Author :

Institut. :

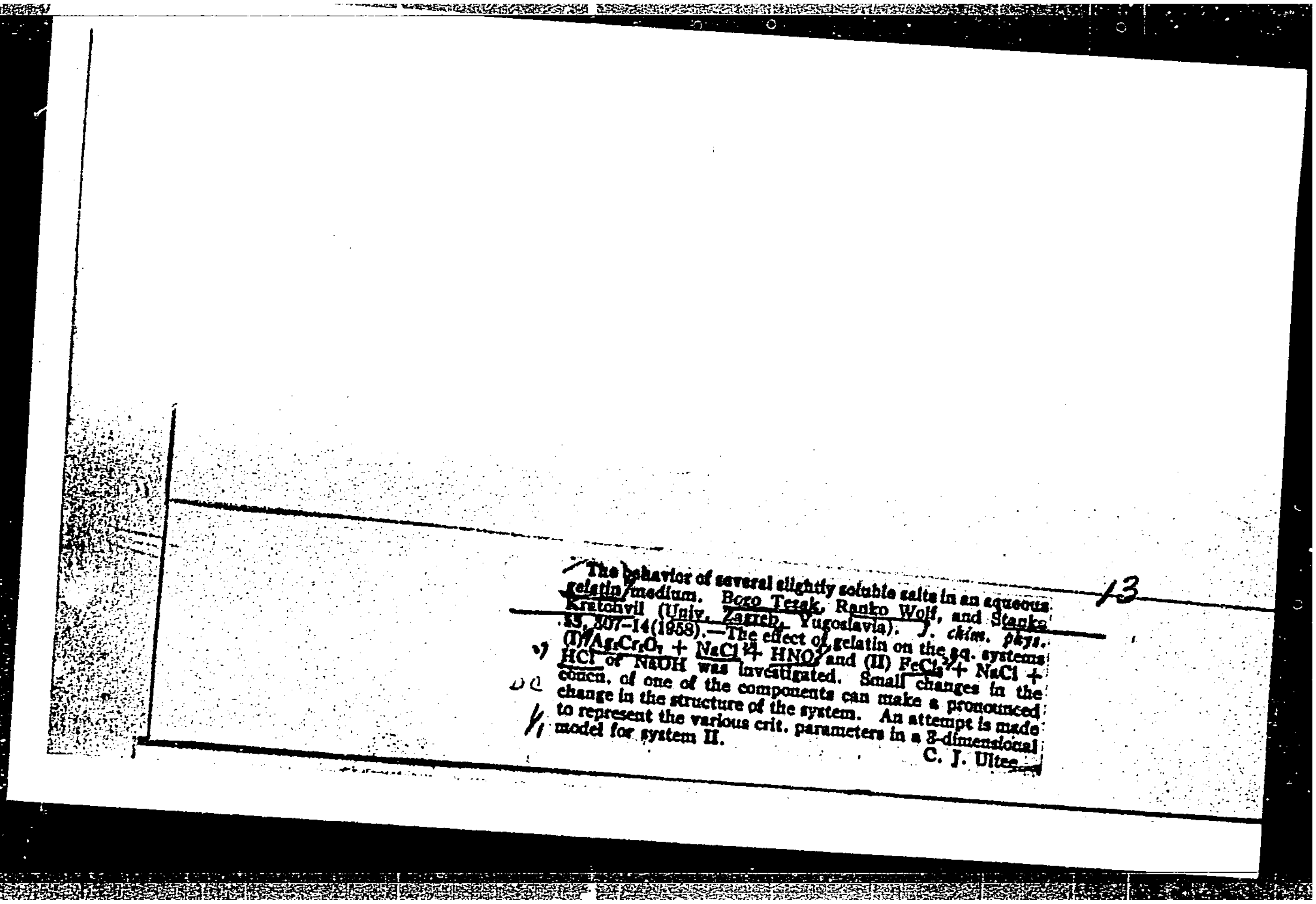
Title :

Orig. Pub. :

Abstract : concentration of Ag^+ and Br^- (I^-) there is primarily formed the less soluble AgBr (AgI) and particles are stabilized by Br- (I-) ions. On deficiency of Br- (I-) ions stabilization is determined by Cl^- ions, taking part in building of crystal lattice. AgCl-AgCN system behaves analogously. In AgBr-AgI system, on increase of I^- concentration c_k gradually approximates c_k for AgI sol, which is due to conjoint participation of Br- and I- in building of crystal lattice and in stabilization of sol. Preceding communication see RZhKhim, 1958, 641, 27879. -- G. Vasil'yev.

Card: 2/2

B-27



13

The behavior of several slightly soluble salts in an aqueous gelatin medium. Bozo Tezak, Ranko Wolf, and Stanko Kratohvil (Univ. Zagreb, Yugoslavia). *J. chim. phys.* 55, 307-14(1958).—The effect of gelatin on the aq. systems (I) $Ag_2Cr_2O_7$ + NaCl + HNO_3 and (II) $FeCl_3$ + NaCl + HCl or NaOH was investigated. Small changes in the concn. of one of the components can make a pronounced change in the structure of the system. An attempt is made to represent the various crit. parameters in a 3-dimensional model for system II.

C. J. Ultee

FUREDI, H.; TEZAK, B.

Precipitation and hydrolysis of uranium(VI) in aqueous solutions. Pt. 2. Croat chem acta 36 no.3:119-131 '64.

1. Department of Physical Chemistry of the Ruder Boskovic Institute, Zagreb, and Institute of Physical Chemistry of the Faculty of Mathematics and Natural Sciences, Zagreb. Submitted April 17, 1964.

TEZAK, B.

Prof. Maks Samec, 1881-1964; obituary. Croat chem acta 36 no.3: 177-178 '64.

Prof. Panta Tutundzic, 1900-1964; obituary. Ibid.:179-180

TEZAK, D.; TEZAK, B.

Methorics of the precipitation processes. Part 19. Croat chem acta 36 no.2:59-66 '64.

1. Institute of Physical Chemistry of the Faculty of Natural Sciences and Mathematics of the University of Zagreb, Zagreb.

BILINSKI, H.; FÜREDI, H.; TEZAK, B.

Precipitation and hydrolysis of thorium (IV) in aqueous solution. II. Influence of pH and neutral electrolytes upon the precipitation in the system thorium nitrate - potassium phthalate. Croat chem acta 35 no.1:31-42 '63.

1. Department of Physical Chemistry, Institute "Ruder Boskovic" and Laboratory of Physical Chemistry, Faculty of Science, University of Zagreb, Zagreb, Croatia, Yugoslavia.

BILINSKI, H.; FÜREDI, H.; BRANICA, M.; TEZAK, B.

Precipitation and hydrolysis of thorium (IV) in aqueous solution: throium nitrate - potassium hydroxide. I. Determination of solubility constants on Th (OH)4. Croat chem acta 35 no.1:19-30 '63.

1. Department of Physical Chemistry, Institute "Ruder Boskovic", and Laboratory of Physical Chemistry, Faculty of Science, University of Zagreb, Zagreb, Croatia, Yugoslavia. 2. Glavni urednik, "Croatia chemica acta" (for Tezak).

WOLF, R. H. H.; TEZAK, B.

The behaviour of the system: gelatin-ferric chloride-neutral electrolyte-hydrochloric acid or sodium hydroxide in an aqueous medium. II. The behaviour of the simpler system: ferric chloride-sodium hydroxide in an aqueous medium. III. The behaviour of the simpler system: ferric chloride-sodium chloride-sodium hydroxide in an aqueous medium. Croat chem acta 33 no.1:1-13 '61.

1. Laboratory of Physical Chemistry, Faculty of Science, University of Zagreb, Zagreb, Croatia, Yugoslavia. 2. Editorial Board, "Croatia chemica acta", editor (for Tezak).

(Gelatin) (Iron chlorides) (Salt)
(Hydrochloric acid) (Sodium hydroxide)
(Hydrogen-ion concentration)

TOMAZIC, B.; BRANICA, M.; TEZAK, B.

Precipitation and hydrolysis of uranium (VI) in aqueous solutions: uranyl nitrate-potassium hydroxide-neutral electrolyte. Croat chem acta 34 no.1:41-50 '62.

1. Department of Physical Chemistry, Institute "Ruder Boskovic" and Laboratory of Physical Chemistry, Faculty of Science, University of Zagreb, Zagreb, Croatia, Yugoslavia.

TEZAK, B.

Methorica of the coagulation processes. XIV. Simultaneous coagulation of the mixed systems of silver halides, silver cyanide, and silver thiocyanate by univalent counterions. M. Biffl and B. Težak (Univ. Zagreb, Yugoslavia). *Croat. Chem. Acta* 30, 9–14(1958)(in English); cf. *C.A.* 50, 16272*h*.—Coagulation of mixed slightly sol. Ag salts in presence of KNO_3 was followed by a tyndallometric technique (*C.A.* 46, 1842*e*). The sol concn., detd. by the concn. of $AgNO_3$ in reacting mixts., was const. at $2 \times 10^{-4} M$. The concn. of the component that made a more sol. ppt. with Ag (KCl or KBr, resp.) also was const. ($2 \times 10^{-3} M$) in order to avoid the effect of changing the concn. of stabilizing ion. Series were prepd. within each system with different concns. ($0-2 \times 10^{-3} M$) of the component that made a less sol. ppt. with Ag. The concn. of counterions (K^+) varied within such a particular series, taking into consideration the concn. of K^+ ions both in neutral electrolyte (KNO_3) and in the reacting components. The concn. gradient of counterions was extended on both sides of the coagulation value. The crit. time of coagulation was 10 min., except in the case of AgCl–AgSCN where it was 1 min. The temp. was 20°. In the system $AgNO_3$–KCl–KBr–KNO_3 small addn. of Br^- to the pure chloride system did not change the coagulation value of KNO_3 until the concn. of Br^- became equal to the concn. of Ag^+ when it suddenly shifted to the coagulation value of the pure AgBr. Similar results were obtained in the systems $AgNO_3$–KCl–KI–KNO_3 and $AgNO_3$–KCl–KCN–KNO_3. This indicates the original and entire pptn. of less sol. component (AgBr, AgI, and AgCN, resp.). When there are insufficient Br, I, or CN ions to react with all Ag^+ ions, chloride ions participate in building up the crystal lattice and stabilize the particles. After the equality of the concns. of Ag^+ and Br^- (I^- or CN^-, resp.) is reached there are still enough of these anions in soln. even for the stabilization of particles and the coagulation value changes to that of pure AgBr (AgI or AgCN, resp.). In the system $AgNO_3$–KCl–KSCN–KNO_3 small addn. of SCN^- to the pure chloride system greatly reduced its turbidity. Further addn. increased it again and the coagulation curve gradually acquired thiocyanate character, which was still not reached, however, at the concn. of SCN^- equal to that of Ag^+. In the system $AgNO_3$–KBr–KI–KNO_3 addn. of I^- to pure bromide system gradually shifted the coagulation value toward that of the pure AgI. The behavior of the last 2 systems indicates simultaneous pptn. of both components, presumably in the form of mixed crystals.

J. Kratohvil

WOLF, R.H.H.; TEZAK, Bozo, dr.,prof.

The behavior of the system; gelatin-ferric chloride-neutral electrolyte-hydrochloric acid or sodium hydroxide in an aqueous medium. II. The behavior of the simpler system: ferric chloride--sodium hydroxide in an aqueous medium. Croat chem acta 33 no.1: 1-8 '61..

1. Laboratory of Physical Chemistry, Faculty of Science, University of Zagreb, Croatia, Yugoslavia 2. Editor,"Croatiac chemica acta, Arhiv za kemiju" (for Tezak)-

TEZAK, Bozo

Ordinary annual session of the Croatian Chemical Society. Report of the Editor. Croat chem acta 34:no.2:Suppl.:34:A3-A4 '62.

1. Clan Redakcionog odbora i glavni urednik, "Croatica Chemica Acta".

TEZAK, Bozo

Ordinary annual session of the Croatian Chemical Society. Report of the Permanent Member at the Union of Chemical Societies of Yugoslavia. Croat chem acta 34 no.2:Suppl.: 34:A8 '62.

1. Clan Redakcionog odbora i glavni urednik, "Croatica Chemica Acata".

WOLF, R.H.H.; TEZAK, Bozo,dr.,prof.

The behavior of the system: gelatin-ferric chloride-neutral electrolytite-hydrochloric acid or sodium hydroxide in an aqueous medium. III. The behavior of the simpler system: ferric chloride--sodium chloride-sodium hydroxide in an aqueous medium. Croat chem acta 33 no.1:9-13 '61.

1. Laboratory of Physical Chemistry, Faculty of Science, University of Zagreb, Croatia, Yugoslavia. 2. Editor, "Croatica chemica acta, Arhiv za kemiju" (for Tezak).

TEZAK, D.; TEZAK, B.

Methorics of the precipitation processes. Part 19. Croat chem acta 36 no.2:59-66 '64.

1. Institute of Physical Chemistry of the Faculty of Natural Sciences and Mathematics of the University of Zagreb, Zagreb.

TEZAK, M.

SCIENCE

TEZAK, M. (Hrvatsko Kemijsko drustvo, Sveuciliste u Zagrebu i Hrvatsko prirodoslovno drustvo) Zagreb.

No.2, 1958. Methorics of the precipitation processes. XVIII. The influence of temperature and concentration on the precipitation of silver bromide. In English. p. 119.

Monthly Index of East European Accessions (EEAI) LC, Vol 8, No. 4, April 1959

MARICIC, S.; TEZAK, S. "Production of bayerite from aluminum."
Arhiv Za Kemiju, Zagreb, Vol 26, No 2, July 1954, p. 117

SO: Eastern European Accessions List, Vol 3, No 10, Oct 1954, Lib. of Congress

TEŽAK, S.

Oxidation of aluminum amalgam. Stanko Težak (Fac. Sci., Zagreb, Yugoslavia). Croat. [illegible] (English summary). — Al turnings of 99.99% purity were cleaned with 5% NaOH at 30-40° for 3-5 min., washed with water, and exposed to 5% $HgCl_2$ soln. for 1-2 min. at room temp. The Al-Hg was decompd. with H_2O and oxidized; the product obtained consisted of 59.5% Al_2O_3 and 40.5% Hg, which volatilized completely upon heating. Thermal analyses were made at 50, 80, 100, 110, 115, 135, 150, 170, 180, 200, 245, 300, 400, 500, 600, 700, and 800° and temps. were measured with a Pt-PtRh thermopile and a Lange Multiflex galvanometer. Debye-Scherrer x-ray diagrams were obtained at 20° and at 100-900° in 100° steps. Electron microscope diffraction patterns were made at 20°, 110°, 200°, and up to 900° in 100° steps. These methods all proved that the Al_2O_3 obtained was the cubic modification.

Werner Jacobson

TEZAK, S.

YUGO.

✓**Preparation of bayerite from aluminium.** S. Maričić and S. Težak (*Arhiv Kem.*, 1954, 26, 117—118).—The use of 99·99% pure Al and of conductivity water for preparation of bayerite, as recommended by Schmäh (*Z. Naturforsch.*, 1946, **1**, 323), is not necessary; equally good results are obtained with ordinary Al, containing Fe 0·33 and Si 0·36%, and with distilled water. The product always contains some boehmite, formed secondarily from bayerite.

R. TRUSCOE.

Gas in Aluminium Alloys. F. Nonveiller, A. Lahodny, and S. [illegible] (*Tehnički pregled (Zagreb)*, 1953, 5, 51-56; *C. Abs.*, 1955, 49, 8772).—In casting Al alloys contg. up to 1% Mg, the following mixture added to the extent of 0·25–0·50% of the melt was found suitable as a degassing agent: C_2Cl_6 25, NaCl 36·5, KCl 28·5, NaF 5, and Na_3AlF_6 5%. A laboratory apparatus for controlling the gas content of Al alloys is described.

MG

②

Tezak S.

HUNG.

Gas in aluminum alloys. [illegible] A. [illegible] S. Tezak. [illegible] 5, 51 [illegible] ~~casting Al~~ alloys contg. [illegible] Mg, the following [illegible] added in the amt. of [illegible] was found [illegible] gassing agent: [illegible] NaAlF [illegible] Al alloys is discussed. [illegible]

m

TEŽAK S.

YUGO.

Distribution of metals in bauxitic oölites. H. Iveković, S. Težak, and T. Marin. *Bull. intern. acad. yougoslave sci. et beaux-arts*, [N.S.] *Livre* 12, *Classe sci. math., phys. et tech.*, *Livre* 4, 89-92(1954)(in English).—See *C.A.* 48, 6925g. N. Plavšić

C P ②

TEZAK, S

YUGO.

✓**The preparation of bayerite from aluminum.** S. Maričić and S. Težak (Inst. lake metale, Zagreb, Yugoslavia). *Arhiv kem.* 26, 117-18(1954).—Bayerite was prepd. from Al by modifying the Schmäh procedure (*C.A.* 41, 5359*f*) in that distd. H_2O and Al contg. 0.28 Fe and 0.10% Si were used instead of cond. H_2O and refined 99.99% Al. With Al contg. 0.33 Fe and 0.30% Si both bayerite and böhmite resulted. X-ray diffraction patterns of the products prepd. are given.

N. Plavšić

TEZAK, S.

④ go

The distribution of metals in bauxitic oölites. H. Iveković, S. Težak, and T. Marin (Zagreb Univ., Yugoslavia). *Radovi Jugoslav. Akad. Znanosti i Umjetnosti* 296, 101-8 (1953).—Chem. analyses of 7 fractions of oölites obtained from a sample of bauxite from Drniš (Yugoslavia) indicate that linear relationships exist between their Fe_2O_3 and V_2O_5 or Cr_2O_3 contents. For calcg. the approx. V_2O_5 and Cr_2O_3 contents of a bauxite one can use the formulas: % V_2O_5 = % Fe_2O_3 × 4.8 × 10^{-3} and % Cr_2O_3 = % Fe_2O_3 × 5.1 × 10^{-3}. There are no relationships with regard to the MnO, TiO, Al_2O_3, and SiO_2 contents. N. Plavšić

11/15/54

TEZAK, Stanko, inz. kemijske tehnologije (Rudnik i zeljezara Vares, Vares)
NOVAK, Visnja, inz. kemijske tehnologije

Blast furnaces. Tehnika Jug 19 no.1:Suppl:Rudarstvo metalurg 15 no.1:67-72 Ja '64.

1. Sef OTK Rudnika i zeljezare Vares, Vares (for Tezak). 2. Sef metalurske operative Rudnika i zeljezare Vares, Vares (for Novak).

TEZAK, S.

"Substances obtained through oxidation of alumium amalgam", p. 45 (Arhiv Za Kemiju., Vol. 24, 1952, Zagreb)

East European Vol. 2, No 9
SO: Monthly List of [redacted] Accessions, Library of Congress, September 1953, Uncl.

USSR / General Biology. Individual Development. B
Regeneration.

Abs Jour	: Ref Zhur - Biologiya, No 4, 1959, No. 14391
Author	: Tezekbayev, S. D.
Inst	: Not given
Title	: The Regeneration of Skin After Preliminary Treatment with a Chemical Irritant (Ozokerite Salve)
Orig Pub	: Byul. eksperim. biol. i meditsiny, 1956, 41, No 5, 66-69
Abstract	: One ear of rabbits of the first series was smeared with ozokerite salve (OS) for 7 days every day, and then tunnel wounds were inflicted upon both ears. The wounds of ears healed much sooner if at the time of operation their tissues were irritated because of

Card 1/3

USSR / General Biology. Individual Development. B
Regeneration.

Abs Jour : Ref Zhur - Biologiya, No 4, 1959, No. 14391

the OS effect (the epidermis was 40 mu thick instead of 12-18 mu as in controls). The character of the healing process did not change. In rabbits of the second series one ear was treated with OS for 7-10 days before the operation and the second ear 7-10 days after the operation. The epithelization of the wound proceeded much faster in ears treated before the operation, while the healing of the wound terminated much sooner in ears treated with salve during the period of healing. When the ears of rabbits of the 3rd series were treated with OS every day for 7 days before the operation and every day (until the wounds were healed) after the operation, the wounds

Card 2/3

USSR / General Biology. Individual Development.
Regeneration.

B

Abs Jour : Ref Zhur - Biologiya, No 4, 1959, No. 14391

healed much faster than when the salve was only applied during the postoperative period. -- I. V. Markelova

Card 3/3

BRAUN, A.A.; TEZEKBAYEV, S.D.

Regeneration of preliminarily irritated tissues. Trudy KirgNOAGZ no.2:22-24 '65. (MIRA 18:11)

1. Iz kafedry gistologii (zav. - prof. A.A.Braun) Kirgizskogo gosudarstvennogo meditsinskogo instituta.

TEZEKBAYEV, S.D.

Regeneration of skin after preliminary treatment with a chemical irritant (azokerit ointment) Biul.eksp.biol.med. 41 no.5:66-69 May '56.
(MLRA 9:8)

1. Iz kafedry gistologii (zav. prof. A.A.Braun) Kirgizskogo meditsinskogo instituta, Frunze. Predstavlena deystvitel'nym chlenom AMN SSSR D.N.Nasonovym.

(PETROLEUM PRODUCTS, eff.
petrolem ointment, on regen. of pretreated skin after exper. wds. on rabbit ears)
(SKIN, wounds and inuries
exper., eff. of petroleum ointment pretreatment on regen. of rabbit ears)
(WOUNDS AND INJURIES, exper.
skin, eff. of petroleum ointment pretreatment on regen. of rabbit ears)

TEZEKBAYEV, S.D

[illegible]

TEZENKOVA, V.V., mladshiy nauchnyy sotrudnik

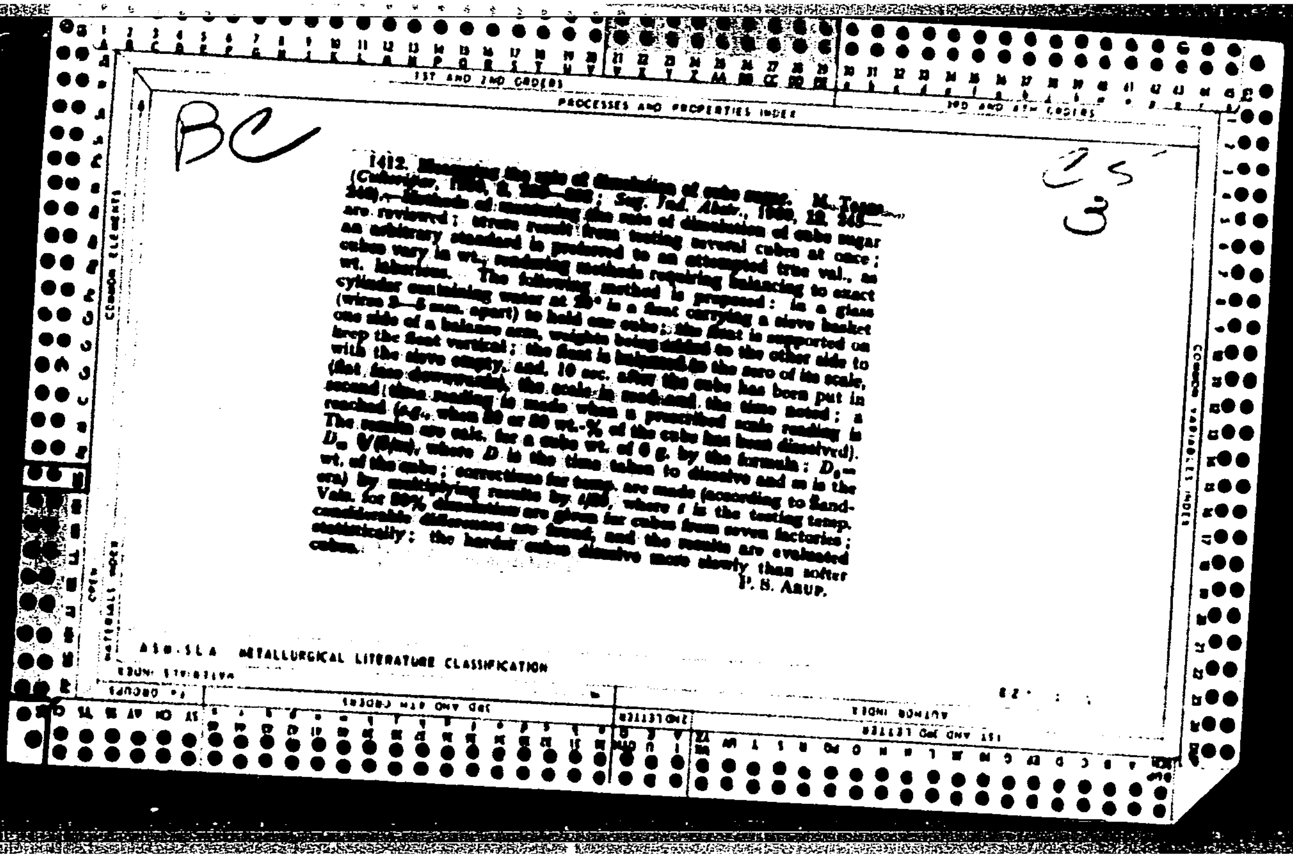
Use of the "KMTs" thickener in printing rayon fabrics with titanium dioxide and bronze powder. Tekst.prom. 22 no.6: 62-63 Je '62. (MIRA 16:5)

1. Kompleksnyy nauchno-issledovatel'skiy institut legkoy promyshlennosti Latviyskogo soveta narodnogo khozyaystva (KNIILP).

(Latvia--Textile printing)

TEZENKOVA, V.V., inzh.-khimik

Use of Czechoslovak auxiliary products in finishing processes. Tekst.prom. 21 no.2:68-69 Ja '61. (MIRA 14:3)
(Textile finishing)



BC

35
3

1412. Measuring the rate of dissolution of cube sugar. M. Tegze (Cukoripar, [illegible]; Sug. Ind. Abstr., [illegible], 18, 245—246).—Methods of measuring the rate of dissolution of cube sugar are reviewed; errors result from testing several cubes at once; an arbitrary standard is preferred to an attempted true val., as cubes vary in wt., rendering methods requiring balancing to exact wt. laborious. The following method is proposed: In a glass cylinder containing water at 20° is a float carrying a sieve basket (wires 5—6 mm. apart) to hold one cube; the float is supported on one side of a balance arm, weights being added to the other side to keep the float vertical; the float is balanced to the zero of its scale, with the sieve empty, and, 10 sec. after the cube has been put in (flat face downwards), the scale is read and the time noted; a second (time) reading is made when a prescribed scale reading is reached (e.g., when 50 or 80 wt.-% of the cube has been dissolved). The results are calc. for a cube wt. of 6 g. by the formula: $D_6 = V(6/m)$, where D is the time taken to dissolve and m is the wt. of the cube; corrections for temp. are made (according to Sandera) by multiplying results by $t/20$, where t is the testing temp. Vals. for 80% dissolution are given for cubes from seven factories; considerable differences are found, and the results are evaluated statistically; the harder cubes dissolve more slowly than softer cubes.

P. S. Arup.

EKONOMU, T.; FRUNZA, A.; GAVRILITA, N.; TEZHA, Zhorzhen

Pseudarthrosis of the graft following arthrodesis in Pott's disease in children. Khirurgiia 15 no.2/3:242-246 '62.

1. Iz Detska klinika po khirurgiia i ortopediia, Meditsinski institut - IAsh.
(TUBERCULOSIS SPINAL surg)
(PSEUDARTHROSIS etiol)

1ST AND 2ND ORDERS

PROCESSES AND PROPERTIES INDEX

CA

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Raw materials for the nonferrous metallurgical plants at Ordjonikidze (formerly Vladikavkaz). N. Tsaley. *Tsvetnue Metal.* 1935, No. 2, 24-8.—The known deposits of Pb and Zn have increased since 1930, when they were 50,100 and 101,800 tons, resp., to 390,500 and 677,800 tons, resp., in 1934. S. L. Madorsky

COMMON ELEMENTS

MATERIALS INDEX

ASM-SLA METALLURGICAL LITERATURE CLASSIFICATION

TEZHKY, A. [Tezky, A.]

Development of radioactive core sampling. Prace ust naft 18:40-41 '61.

14(1) SOV/66-59-5-21/35

AUTHOR: Tezikov, A., Candidate of Technical Sciences

TITLE: Prevention of Deformation of Walls in the Ice Producer of a Dry Ice Installation

PERIODICAL: Kholodil'naya tekhnika, 1959, Nr 5, p 63 (USSR)

ABSTRACT: The majority of the dry ice plants in the USSR use ice producers of the Moscow Plant "Kompressor" which operate on liquid carbon dioxide under 8 atmospheres pressure. The article describes design and functioning of the apparatus which consists of an inner and outer compartment connected with 2 diaphragms. These should be handled with great care following precisely the instruction of the makers. Premature opening permits part of the liquid carbon dioxide to penetrate into the jacket, where it forms a hard substance which is dangerous. Likewise a sudden closing of the diaphragms and of the shut-off valve brings about a rapid rise of pressure in the jacket, which causes deformation of the walls of the inner compartment which seriously damages the ice producer. The article closes

Card 1/2

SOV/66-59-5-21/35

Prevention of Deformation of Walls in the Ice Producer of a Dry Ice Installation

with recommendations as to how prevent such accidents. There is one diagram.

Card 2/2

SOV/66-59-1-14/32

AUTHORS: Popova, K., Tezikov, A.

TITLE: The Work of the Dry Ice Plants Using a Solution of Monoethanolamine With the Addition of Calcinated Soda (Rabota zavodov sukhogo l'da na rastvore monoetanolamina s dobavleniyem kal'tsinirovannoy sody)

PERIODICAL: Kholodil'naya tekhnika, 1959, Nr 1, pp 56-57 (USSR)

ABSTRACT: The Experimental Dry Ice Plant of VNIKhI has investigated the effectiveness of adding to a solution of monoethanolamine a certain amount of calcinated soda. To a polluted solution of monoethanolamine, which had been in operation for 126 days, portions of calcinated soda (74 g/l) were added, which raised the coefficient of absorption by 16.9% and increased the daily output by 15.8%. Moreover, the presence of calcinated soda in the solution reduces the mechanical losses of monoethanolamine by 35% and eliminates its consumption for polymerization and combination with sulfuric compounds. Various dry ice plants have followed the recommendation given by VNIKhI and were likewise successful in reducing the consumption of mono-

Card 1/2

SOV/66-59-1-14/32

The Work of the Dry Ice Plants Using a Solution of Monoethanolamine With the Addition of Calcinated Soda

ethanolamine and increasing production of dry ice. A comparative table gives the detailed results obtained by 4 different dry ice plants in Leningrad, Moscow, Kiyev and Kramatorsk. There is one table.

Card 2/2

14(1) SOV/66-59-5-23/35

AUTHOR: Tezikov, A.

TITLE: Equipping and Putting Into Operation of the First Dry Ice Plant in People's Republic of China

PERIODICAL: Kholodil'naya tekhnika, 1959, Nr 5, p 65 (USSR)

ABSTRACT: The author who participated in the installation and launching of the first dry ice plant belonging to the Girinskiy Chemical Combine, describes the machines and equipment supplied by the Moscow Plant "Kompressor" and the Odessa Press Plant. The capacity of the plant is 25 tons of dry ice per day. As basic raw material expander gas is used, which is obtained as by-product from the production of synthetic ammonia and methanol in the chemical combine. The production of the dry ice is rendered comparatively simple by the fact that the expander gas, before entering the dry ice plant, undergoes preliminary purification from hydrogen sulfide. After being compressed by a 3-stage compressor to a pressure of 65 atmospheres, the gas is directed to a condenser of CO_2-evaporator of NH_3 in which the carbon dioxide is liquified. The liquid gas enters into high pressure receivers of the dry ice installation, while the uncondensed gases (CO, H_2, N_2, O_2) pass from the condenser into the atmosphere. The dry ice press delivers blocks of

Card 1/2

SOV/66-59-5-23/35

Equipping and Putting Into Operation of the First Dry Ice Plant in People's Republic of China

dry ice measuring 500 x 500 x 250 mm and weighing 100 kg. The operation of this press is fully automatic; the removal of the ice blocks, their cutting up and packing is mechanized.

Card 2/2

AUTHOR: Tezikov, A. SOV/66-59-1-18/32

TITLE: Optical Level Indicator (Svetovoy ukazatel' urovnya)

PERIODICAL: Kholodil'naya tekhnika, 1959, Nr 1, p 60 (USSR)

ABSTRACT: So far there has not existed a reliable instrument for determining the level of liquid carbon dioxide in intermediate tanks of cold storage houses. In the Experimental Refrigeration Plant of VNIKhI an optical level indicator has been designed and fitted to a carbon dioxide reservoir. This level indicator has now been in operation for over a year. Beside the author of this article, I. Barulina and I. Golubev of VNIKhI have participated in the construction of the apparatus. This device consists of a float which moves up and down in a cylinder which communicates with the carbon dioxide reservoir. The float is connected with a pin which follows the movement of the float and actuates a switch, which on its

Card 1/2

Optical Level Indicator SOV/66-59-1-18/32

passage closes or interrupts circuits, which in turn light electric bulbs, thus indicating different levels attained in the reservoir.

There is 1 diagram.

Card 2/2

POPOVA, K.; TEZIKOV, A.

Dry ice plants operating with a solution of monoethanolamine containing soda ash. Khol.tekh. 36 no.1:56-57 Ja-F '59. (MIRA 12:3)
(Dry ice) (Ethanol)

TEZIKOV, A.

Level indicator using electric lamps. Khol.tekh.36 no.1:60 Ja-F '59.
(MIRA 12:3)
(Dry ice) (Liquid level indicators)

TEZIKOV A.

TEZIKOV, A., kand. tekhn. nauk; KRIVOV, A., inzh.

Containers for the transportation and storage of dry ice. Khol. tekh. 34 no.4:48-50 O-D '57. (MIRA 11:1)

(Dry ice)

[illegible]

USSR/Chemical Technology. Chemical Products and Their Application -- Preparation and separation of gases, I-10

Abst Journal: Referat Zhur - Khimiya, No 2, 1957, 5363

Author: Tezikov, A.

Institution: None

Title: Removal of Mechanical Admixtures from Flue Gases

Original Publication: Kholodil'n. tekhnika, 1956, No 2, 57-60

Abstract: In the manufacture of dry ice and liquid carbon dioxide from flue gases (FG), produced by combustion of solid fuels, an important step of the process is the preliminary purification of FG from solid (ashes, soot, etc) and gaseous (SO_2, H_2S, etc) impurities. With a change-over to the new absorbent for CO_2 (monoethanolamine) the conventional purification of FG was found to be inadequate and experiments were carried out on additional purification in a centrifugal scrubber VTI 350 mm in diameter. At an average ingress velocity of the gases of 20 m/second and an initial dust content of 40 mg/nm3

Card 1/2

USSR/Chemical Technology. Chemical Products and Their Application -- Preparation and separation of gases, I-10

Abst Journal: Referat Zhur - Khimiya, No 2, 1957, 5363

Abstract: the VTI scrubber has shown adequate dust-removal efficacy (~75%). Concurrently with recovery of solid particles a certain amount of SO_2 and H_2S was absorbed in water within the scrubber.

Card 2/2

TEZIKOV, A.; YELUFIMOV, N.

Rotating jet sprayers of monoethanolamine solution. Khol.tekh.
37 no.2:53 My-Ap '60. (MIRA 13:10)
(Ice industry--Equipment and supplies)

TEZIKOV, Aleksandr Dmitriyevich; CHICHKOV, N.V., red.; BABICHEVA, V.V., tekhn.red.

[Production and uses of dry ice] Proizvodstvo i primenenie sukhogo l'da. Izd.2., perer. i dop. Moskva, Gos.izd-vo torg. lit-ry, 1960. 126 p. (MIRA 13:9)
(Dry ice)

PHASE I BOOK EXPLOITATION SOV/5016

Tezikov, Aleksandr Dmitriyevich

Proizvodstvo i primeneniye sukhogo l'da (Manufacture and Utilization of Dry Ice) 2d ed., rev. and enl. Moscow, Gostorgizdat, 1960. 128 p. 3,000 copies printed.

Ed.: N. V. Chichkov; Tech. Ed.: V. V. Babicheva.

PURPOSE: This booklet is intended for technical personnel of dry ice manufacturing plants and for workers using dry ice in various branches of the national economy.

COVERAGE: This edition of the booklet describes the designing of apparatus and flow sheets of dry ice production. Raw materials discussed include flue gases, natural sources, alcohol fermentation, and ammonia synthesis. Some recommendations on the application of dry ice are given. No personalities are mentioned. There are 41 references, all Soviet.

~~Card 1/4~~

PIMENOVA, Tat'yana Fedorovna; NOVOZHILOV, Nikolay Mikhaylovich;
TEZIKOV, A.D., kand. tekhn. nauk, nauchnyy red.; KAPLUN,
M.S., red.; MAMONTOVA, N.N., tekhn. red.

[Use of dry ice in electric welding; scientific report]
Primenenie sukhogo l'da v elektrosvarochnoi tekhnike; nauchnoe
soobshchenie. Moskva, Gos. izd-vo torg. lit-ry, 1961. 36 p.
(MIRA 15:2)

(Electric welding) (Dry ice)

ALTUNDZHI, Sergey Vladimirovich; BUKHARIN, Viktor Vladimirovich; DOBKINA, Yevgeniya Abramovna; KUZNETSOV, Nikolay Mikhaylovich, inzh.; POPOVA, Kseniya Georgiyevna; TEZIKOV, Aleksandr Dmitriyevich; FRADIN, Leon Romanovich; BADYL'KES, I.TS., doktor tekhn.nauk, retsenzent; SKIRSTYMONSKIY, A.I., inzh., retsenzent; PRITYKINA, L.A., red.; SOKOLOVA, I.A., tekhn.red.

[Production and use of liquid carbonic acid] Proizvodstvo i primenenie zhidkoi uglekisloty. Moskva, Pishchepromizdat, 1959. 207 p. (MIRA 13:2)

(Carbonic acid)

TEZIKOV, A.,m kandidat tekhnicheskikh nauk.

Packaging dry ice during the manufacturing process. Khol.tekh.33
no.1:67-68 Ja Mr '56. (Dry ice) (MIRA 9:7)

TEZIKOV, A., kandidat tekhnicheskikh nauk.

Interaction of sulfur dioxide with monorthanolamine. Khol.tekh.33 no.1:67 Ja Mr '56. (Sulfur dioxide)(Ethanol) (MIRA 9:7)

TEZIKOV, A., kandidat tekhnicheskikh nauk.

The possibility of regulating monoethanolamine solution pumping over operations by means of a suction valve. Khol.tekh. 32 no.3:62 Jl - S '55. (MLRA 9:1)
(Ethanol) (Dry ice)

TEZIKOV, A., kandidat tekhnicheskikh nauk.

Determining carbon dioxide content in steam. Khol.tekh. 32 no.3: 56-57 Jl - S '55. (MLRA 9:1)

(Carbon dioxide)

TEZIKOV, A., kandidat tekhnicheskikh nauk.

Flue gas cleaning from mechanical impurities. Khol.tekh.33 no.2: 57-60 Ap-Je '56. (MIRA 9:9)
(Scrubber (Chemical technology))

TEZIKOV, A., kandidat tekhnicheskikh nauk

Isothermic container for transporting dry ice. Khol.tekh. 32 no.1:
36-39 Ja-Mr '55. (MLRA 8:7)
(Dry ice--Transportation) (Containers)

TEZIKOV, A., kandidat tekhnicheskikh nauk

Controlling the capacity of an auxiliary compressor. Khol.tekh. 32 no.1:69 Ja-Mr '55. (MLRA 8:7)
(Compressors)

TEZIKOV, A., kandidat tekhnicheskikh nauk.

Technical and economic indexes of dry-ice plant operations with monoethanolamine and potash. Khol.tekh.31 no.1:26-28 Ja-Mr '54. (MLRA 7:4)

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